



**SYNTHESIS AND CHARACTERIZATION OF
TRIORGANOPHOSPHINEGOLD(I)
THIOLATES**

by

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Amended version of Table 4.4.2.

Complex	H ²	H ⁹	H ⁸	Phenyl Protons	H _a	H _b
6-MPH	8.45(s)	13.62(br,s)	8.26(s)	-	-	-
[dppm(Au(6-MP)) ₂]	8.36(s)	13.12(br,s)	8.17(s)	7.95 - 7.37(br,m)	4.64(m)	-
[dppe(AuCl)(Au(6-MP))]	8.36(s)	13.25(br,s)	8.09(s)	7.87 - 7.47(br,m)	3.02(m)	-
[dppe(Au(6-MP)) ₂]	8.37(s)	13.13(br,s)	8.24(s)	7.89 - 7.49(br,m)	3.04(m)	-
[dppp(AuCl)(Au(6-MP))]	8.40(s)	13.22(br,s)	8.20(s)	7.74 - 7.50(br,m)	3.07(m)	1.73(m)
[dppp(Au(6-MP)) ₂]	8.38(s)	13.18(br,s)	8.19(s)	7.83 - 7.42(br,m)	3.16(m)	1.89(m)

Note: Coupling constants, in parentheses, are in units of Hertz: a: ³J_{H-H}, b: ³J_{P-H} and c: ²J_{P-H}.

DECLARATION

This work contains no material which has been accepted for the award of any other degree or diploma in any university or other institution and, to the best of my knowledge and belief, contains no material previously published or written by another person, except where due reference has been made in the text.

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ABSTRACT

The aim of this work was to synthesize a range of novel triorganophosphinegold(I) 6-mercaptapurinate complexes and to evaluate their potential anti-arthritic activity. The resultant complexes are based on the P–Au–S moiety, structurally related to the widely available pharmaceutical Auranofin, where the phosphorus atom is part of a triethylphosphine ligand and the sulphur atom derived from a tetraacetylated thioglucose anion.

Via alteration of the identity of the triorganophosphine group, a variety of complexes were synthesized from triorganophosphinegold(I) chloride precursors, falling into three main types: 1) triorganophosphinegold(I) 6-mercaptapurinate complexes with the general formula $[R_3PAu(6-MP)]$, where $R_3P = Et_3P, Cycl_3P, PhMe_2P, Ph_3P, (o-Tol)_3P, (m-Tol)_3P$ or $(p-Tol)_3P$; 2) $[\mu-1,n-bis(diphenylphosphino)alkane]gold(I) chloride gold(I) 6-mercaptapurinate$ complexes with the general formula $[(Ph_2P(CH_2)_nPPh_2)(AuCl)(Au(6-MP))]$ where $n = 2$ or 3 ; and 3) $[\mu-1,n-bis(diphenylphosphino)alkane]bis(gold(I) 6-mercaptapurinate)$ complexes with the general formula $[(Ph_2P(CH_2)_nPPh_2)(Au(6-MP))_2]$ where $n = 1, 2$ or 3 . These complexes and the triorganophosphinegold(I) chloride precursors were characterized using multinuclear magnetic resonance, infrared and Fast Atom Bombardment mass spectroscopic techniques. Unambiguous structure determinations of a selection of the complexes were achieved by single crystal X-ray crystallographic methods. Unit cell dimensions were: $[PhMe_2PAuCl]$, orthorhombic space group $P2_12_12_1$, $a = 12.639(4)$, $b = 16.931(6)$, $c = 9.458(3)$ Å, $V = 2024(1)$ Å³ and $Z = 4$; $[Ph_3PAu(6-MP)].C_2H_5OH$, triclinic space group $P\bar{1}$, $a = 11.066(3)$, $b = 13.552(3)$, $c = 8.705(2)$ Å, $\alpha = 91.51(2)$, $\beta = 113.06(2)$, $\gamma = 89.69(2)^\circ$, $V = 1200.8(5)$ Å³ and $Z = 2$; and $[(o-Tol)_3PAu(6-MP)].C_2H_5OH$, monoclinic space group $P2_1/n$, $a = 10.067(2)$, $b = 10.518(2)$, $c = 25.416(4)$ Å, $\beta = 98.42(2)^\circ$, $V = 2662.1(9)$ Å³ and $Z = 4$. The structures were refined to final R values of 0.035, 0.034 and 0.040, respectively, for reflections satisfying the $I \geq 3.0\sigma(I)$ criterion: 1608, 3978 and 4183, respectively. The results for $[PhMe_2PAuCl]$ were utilized, in part, for a cone-angle to bond length correlation

study on triorganophosphinegold(I) chloride complexes. The structures of $[\text{Ph}_3\text{PAu}(6\text{-MP})]$ and $[(o\text{-Tol})_3\text{PAu}(6\text{-MP})]$ revealed a near linear P–Au–S chromophore, with angles of $173.71(6)$ and $177.03(8)^\circ$, respectively. A crystal structure analysis of a closely related triorganophosphinegold(I) thiolate complex, $[\text{Cycl}_3\text{PAu}(6\text{p}2\text{-TU})]$, revealed a similar P–Au–S chromophore with an angle of $177.6(1)^\circ$. Unit cell dimensions were: monoclinic space group $P2_1/c$, $a = 9.539(2)$, $b = 16.452(4)$, $c = 16.880(2)$ Å, $\beta = 95.37(2)^\circ$, $V = 2637.4(8)$ Å³ and $Z = 4$. The final refinement value was $R = 0.043$, for 3695 reflections with $I \geq 3.0\sigma(I)$. The results for the three thiolate complexes were utilized in a correlation study of cone-angles to intramolecular parameters for triorganophosphinegold(I) thiolate complexes in general.

The combined microanalytical, spectroscopic and crystallographic studies verified the formation of all the thionucleobase complexes mentioned above and demonstrated that the gold centre is linearly bound to both the phosphorus and sulphur atoms.

A number of the triorganophosphinegold(I) 6-mercaptopurinate complexes were tested for their anti-arthritis activity in Dark Agouti rats with promising results.

ABBREVIATIONS

Å	Angström
br	broad
°C	degree Celsius
¹³ C NMR	carbon-13 nuclear magnetic resonance
Cycl	cyclohexyl
d	doublet
dd, dm, dt	doublet of doublets, multiplets, triplets
dec.	decomposition point
dppe	μ-1,2-bis(diphenylphosphino)ethane
dppm	μ-bis(diphenylphosphino)methane
dppp	μ-1,3-bis(diphenylphosphino)propane
Et	ethyl
FAB-MS	Fast Atom Bombardment - mass spectroscopy
g	gram
¹ H NMR	proton nuclear magnetic resonance
Hz	hertz
IR, ir	infrared spectroscopy
^x J _{AB}	coupling constant between nuclei A and B over x bonds
m	multiplet (nmr)
m	medium (ir)
M	molar (mol dm ⁻³)
[M] ⁺	molecular ion
2mbaH	2-mercaptobenzoic acid
6m2-TUH	6-methyl-2-thiouracil
Me	methyl
MHz	mega-hertz

ml	millilitre
mmol	millimole
m.p.	melting point
6-MPH	6-mercaptapurine
MW	molecular weight
m/z	mass to charge ratio
6p2-TUH	6-n-propyl-2-thiouracil
NMR, nmr	nuclear magnetic resonance (Fourier Transform)
N.O.	not observed
Obs.	obscured
^{31}P NMR	phosphorus-31 nuclear magnetic resonance
Ph	phenyl
PhO	phenoxy
ppm	parts per million
q	quartet
s	singlet (nmr)
s	strong (ir)
sh	shoulder
t	triplet
TMS	tetramethylsilane
<i>m</i> -Tol	<i>meta</i> -tolyl
<i>o</i> -Tol	<i>ortho</i> -tolyl
<i>p</i> -Tol	<i>para</i> -tolyl
2-TUH	2-thiouracil
vs	very strong
w	weak
%Y	percentage yield
$\delta(\text{A-B})$	bending frequency of A-B bond
λ	wavelength
$\nu(\text{A-B})$	stretching frequency of A-B bond

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CHAPTER 1

Introduction

1.1 Introduction

Interest in gold(I) compounds has increased in recent years due to the growing importance of certain gold(I) complexes in the treatment of a variety of common ailments. The most well known use of gold compounds in medicine is in the treatment of rheumatoid arthritis, where commercially available drugs are now utilized widely. The compound known as Myochrysine, shown in Figure 1.1, is the lead compound for the treatment of rheumatoid arthritis. Another gold(I) thiolate compound, Auranofin, has been found to be effective in the treatment of both rheumatoid arthritis and cancerous tumours^{1,2}, while other less well known compounds, such as Solganal, Allochrysine and Sanocrisin (Figure 1.1), are also in current use against rheumatoid arthritis. Whereas Auranofin displays some anti-tumour activity^{1,2}, the compound bis(μ -1,2-bis(diphenylphosphino)ethane)gold(I) chloride has shown more promising results in the treatment of cancer³. However, these are only a few examples of how gold plays a useful role in medicine. In fact, gold has a long history of medicinal applications. This chapter will discuss the history of gold compounds in medicine, and the use of modern day gold compounds, especially the class of compounds known as phosphinegold(I) thiolates, in the treatment of rheumatoid arthritis. The chemistry of gold and the triorganophosphinegold(I) thiolate complexes will be introduced, and how the spectroscopic and crystallographic characterizations of a selection of new triorganophosphinegold(I) thiolates is of interest to this field will be discussed.

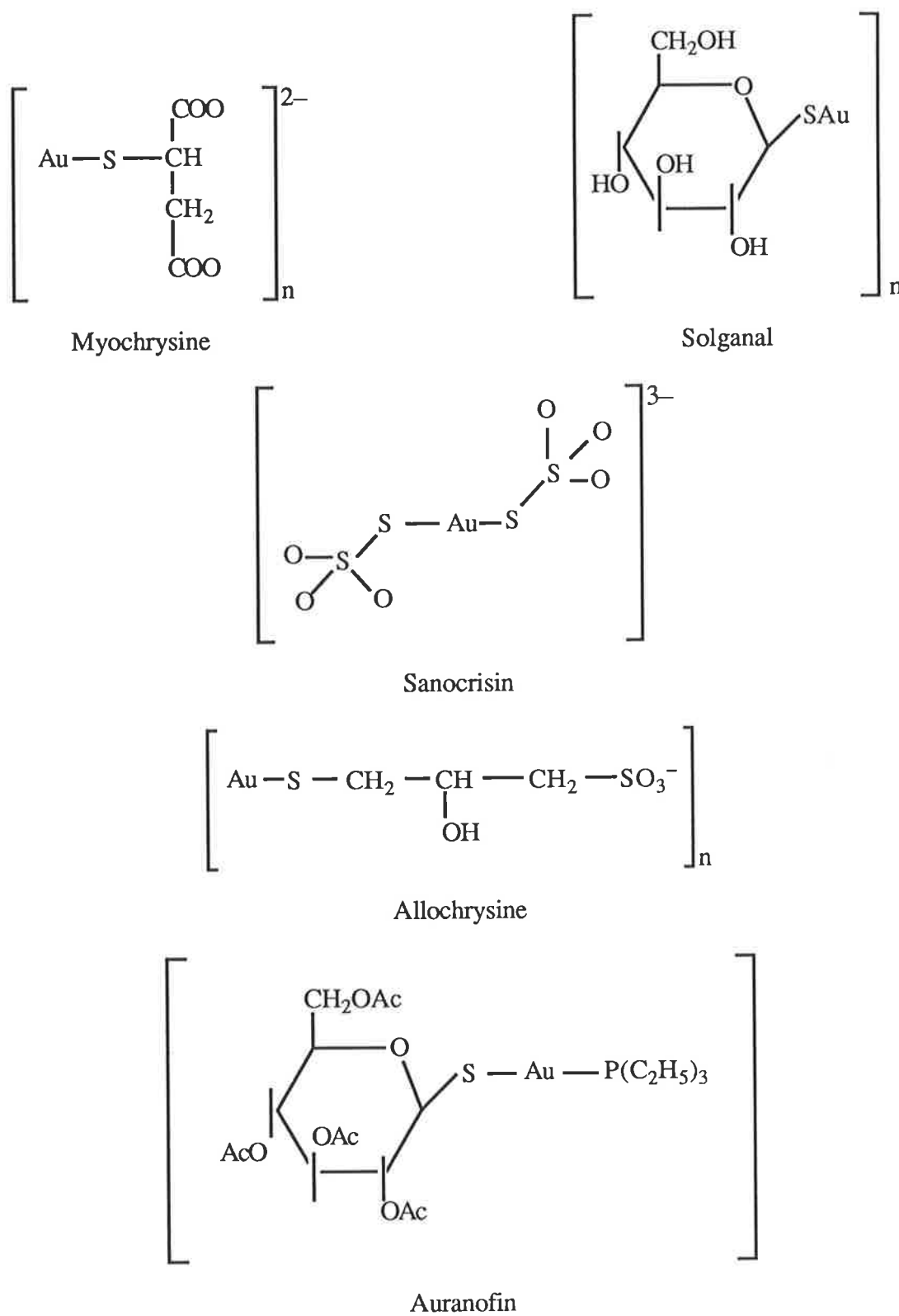


Figure 1.1: Structures Of Some Biologically Active Gold(I) Thiolates.

1.2 History of gold in medicine

The use of gold in medicine dates back to ancient times. As early as 2500 B.C., Chinese and Arabic physicians were reported to have used gold preparations, whilst Pliny in the 1st century recorded that gold could be used successfully in the cures of such ailments as haemorrhoids, warts and fistulas^{4,5}. These early remedies were based on concoctions using metallic gold, and usually associated with ceremonial incantations. Gold was considered a substance of the gods, a holy metal that possessed magical properties: meso-American Indians saw gold as the 'sweat of the sun'; and the Egyptian pharaohs and priests considered it a giver of life, a connection between this world and the next^{4,5}. The rarity of gold and its consequent availability to only the rich and the religious only promoted the superstitious aura surrounding the metal. In reality, the chemical inertness of metallic gold meant that it was probably quite inactive in the internal biochemistry of those who used it. Hence any cures that might have resulted could in a large part have been of a purely psychological nature.

In the Middle Ages, gold began to be utilized in a less superstitious manner. While metallic gold was still being used to gild medicinal tablets, and royalty drank wine from gold coated vessels as a tasteful way of consuming the metal, the 'science' of alchemy was coming to the fore. In the 13th century, Geber reported the preparation of the substance *aqua regia* from a combination of mineral acids, now known to be nitric and hydrochloric acids, having the ability to dissolve gold⁵. This was an important discovery, as actual compounds of gold could at last be prepared. However, physicians of the Renaissance expressed little interest in doing so, since the purity of metallic gold was still considered to be the most important factor in the curative properties of the metal⁵.

The first significant compound to be prepared after this time was the 'muriate of gold and soda', $\text{Na}[\text{AuCl}_4]$ ⁵. The French physician Chrestien in 1811 described this compound as being an effective treatment for syphilis and chronic alcoholism: it is still recognized today as having legitimate effect against at least the former⁵. The advances in medical and scientific knowledge

around this period mark the beginning of modern medicine, and serious investigations into the medicinal applications of gold compounds began.

1.3 Modern applications of gold in medicine

The beginning of the role of gold in modern medicine began with the discovery early this century of the effectiveness of aurothioglucose in treating rheumatic fever⁵. Tuberculosis and rheumatoid arthritis were also found to be combatted to some extent by gold compounds⁵. Despite the promising results that emerged, concerns about the toxicity and side-effects of these compounds, which were usually manifested as kidney and liver problems, to a large extent discouraged further investigations in this context⁵. It was the almost accidental discovery of the *in vitro* anti-cancer activity of the platinum based drug cisplatin⁶ (a square planar Pt(II) complex: *cis*-diaminodichloroplatinum), shown in Figure 1.3, that renewed interest in the possible medicinal usefulness of platinum group metals, and hence research into gold based drugs gained new momentum.

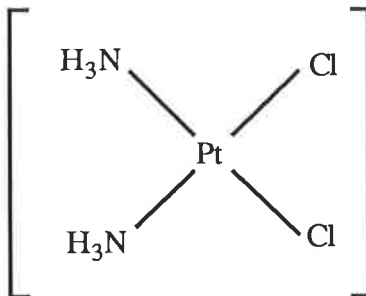


Figure 1.3: *Cisplatin*.

Perhaps the most significant gold compound found to date is Auranofin, discovered in 1972, which is useful in the treatment of rheumatoid arthritis². Recent unpublished results from research conducted in the Department of Chemistry at Adelaide, concentrating on the effect of triorganophosphinegold(I) thionucleobase complexes on rheumatoid arthritis, stems from the success of Auranofin. The project embodied in this thesis is concerned with the characterization of certain triorganophosphinegold(I) thiolate complexes which may also prove to be active against this ailment.

1.4 Arthritis and treatment with gold complexes

The debilitating disease of rheumatoid arthritis is characterized by the chronic inflammation of joint tissue about the skeletal connections in the body, and usually leads with time to the disintegration of bone structure in these regions, crippling the sufferer. The biochemical cause for the onset of the disease is still unclear, but the effects can be neutralized in part by the suppression of the inflammatory action. One physical answer is to apply pressure to the region, especially during periods of rest, when swelling of the synovial fluid due to inflammation is at its greatest. The biochemical approach is by application of certain drugs, although the precise biochemical reactions that make such drugs effective are not known.

The field known as chrysotherapy has hence developed, which is generally defined as the treatment of rheumatoid arthritis with gold-based drugs⁴. Of these drugs, the gold compounds depicted in Figure 1.1 have dominated the treatment of this disease. Two classes of compounds are represented: the polymeric type, such as Solganal (gold sodium thiosulphate), Myochrysine (sodium aurothiomalate), Allochrysine (gold sodium thiopropanol sulphonate) and Sanocrisin (gold sodium thiosulphate), and the monomeric type, such as Auranofin (*S*-2,3,4,5-tetraacetyl-1- β -D-thioglucose(triethylphosphine)gold(I)). The administration of these two classes of compounds is determined by their solubilities in bodily fluids: the polymeric compounds are hydrophilic and are thus injected intravenously, while the monomeric species Auranofin is lipophilic, and can be administered orally¹. In terms of ease and expense of application, the orally administered Auranofin has been found to be the popular compound for routine use.

The difference in solubility and application method might suggest different modes of action for the two classes of compounds. As has been mentioned above, the exact biochemical reactions involved to combat the symptoms of rheumatoid arthritis are not known. A few hypotheses exist, however, as to how the compounds are delivered to the site of inflammation. The polymeric complexes all feature Au-S bonds, and studies on the activity of Myochrysine have revealed that this bond is cleaved in the blood stream in order to create another Au-S bond between the gold and the thiol group of cysteine-34 of a plasma protein albumin^{1,7}. The

resultant auroalbumin complex may then undergo ligand exchange to form a dimeric species, which is then ingested by certain immunological macrophages to be transported to the site of disease. Reactive leukocytes at the sites of inflammation have the ability to produce cyanide from glycine, which reacts with gold(I) to form a $[\text{Au}(\text{CN})_2]^-$ species, thus concentrating gold in this region. The $[\text{Au}(\text{CN})_2]^-$ anion can permeate many types of cell membrane, possibly then interfering with DNA-based functions which lead to the production of inflammatory fluids.

A dose of Auranofin, however, consists of discrete molecules, and the interaction of Auranofin with the albumin protein is likely to be different. Auranofin can be classed as a triorganophosphinegold(I) thiolate: while clinical trials on gold(I) thiolates have proven to be ineffective via oral administration⁷, a phosphine group bound to the gold imparts lipid solubility to the complexes, allowing this mode of administration. It is possible for either the phosphine or thiolate group to be cleaved from Auranofin on reaction with albumin. However, the thiolate group seems to be the more labile, and initial binding to the albumin appears to occur via loss of thiolate¹. The phosphine group can be substituted for another thiol molecule accompanied by oxidation of the phosphine, and the gold centre then becomes doubly coordinated by sulphur atoms, and is thus carried to the site of inflammation. The final fate of the gold is hence the same for both polymeric and monomeric compounds, the main differences in activity being due to the initial mode of transportation in the body. This is possibly related to the solubilities of the molecules⁸.

In finding other compounds that might be useful in this field, it is necessary to alter the structure of the present compounds in order to obtain complexes of differing solubilities (if we wish to work on the same hypothesis concerning their initial metabolism). The polymeric molecules are restricted in their chemical composition by the ability to form a polymer. Monomeric compounds, however, are neutral and discrete, and the structures of the phosphine and thiolate moieties can potentially be altered to achieve a desired measure of solubility for an orally ingested drug.

1.5 Recent advances in the treatment of rheumatoid arthritis by triorganophosphinegold(I) thiolates

Previous work, published and unpublished, has been performed on other compounds analogous to those appearing in this thesis. The utilization of thionucleobases as the thiolate component is a choice that can be based on toxicity; the biological system contains many examples of nucleobases, the most well known being those associated with the macromolecules RNA and DNA. 2-thiouracil (2TUH), the thio analogue of uracil, bound to gold in the complex 2-thiouracilato(triphenylphosphine)gold(I), Figure 1.5.1, has been tested for anti-arthritis activity in a model rat strain and found to be comparable with Auranofin⁹. Other pyrimidine- and purine-based gold(I) phosphines have also been tested in a study¹⁰ which yielded an interesting general result. The best phosphine found in terms of the activity of the resultant gold(I) complex was triphenylphosphine; the triethylphosphine analogues, e.g. Figure 1.5.2, were found to be ineffective and even toxic in some cases. This contrasts with Auranofin, which is active and contains a Et₃P group. Such an anomalous result goes against any structure-activity correlation that might have been forthcoming from the study.

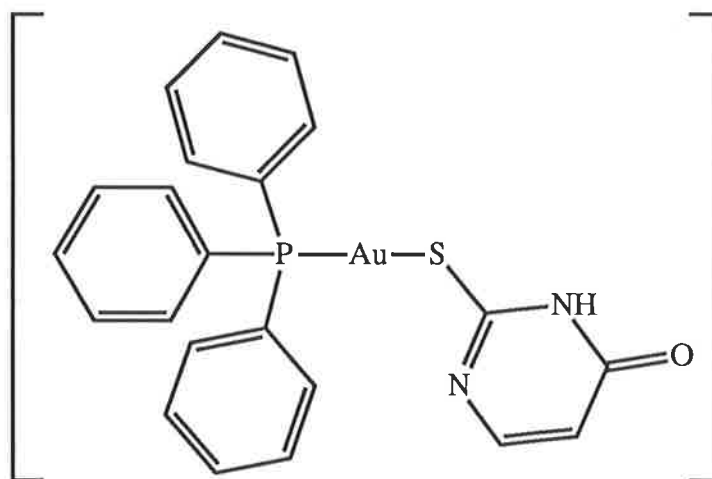


Figure 1.5.1: 2-thiouracilato(triphenylphosphine)gold(I).

The most effective thionucleobase of those tested was found to be 6-mercaptopurine, with the complex 6-mercaptopurinato(triphenylphosphine)gold(I), [Ph₃PAu(6-MP)], shown in Figure 1.5.3, being even more active against the disease than Auranofin. This compound has already been noted in the literature as having significant antineoplastic activity against leukemia in

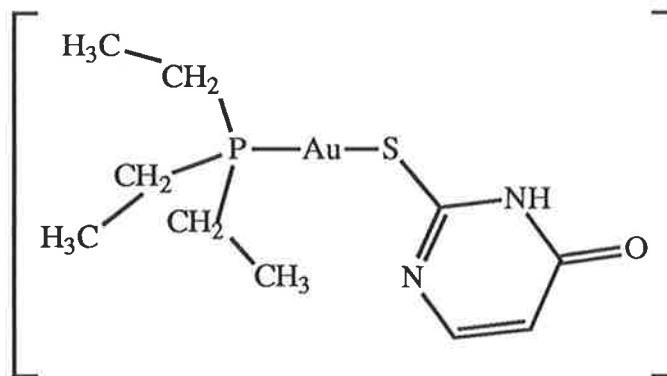


Figure 1.5.2: 2-thiouracilato(triethylphosphine)gold(I).

mice¹¹. 6-mercaptopurine (6-MPH) itself has been reported as possessing chemotherapeutic activity, notably in the treatment of leukemia¹², and its complexes have shown anti-tumor activity e.g. Cu(I)(6-MPH)Cl₂ and Cd(6-MPH)₄Cl₂¹³. Other uses found have been for the inhibitory action by ribonucleoside analogues on *de novo* purine biosynthesis¹⁴, and against such conditions as urate microcrystal arthritis in poultry¹⁵. The compound (8-thiotheophyllinato)(triphenylphosphine)gold(I), a structural variant of [Ph₃PAu(6-MP)], has also been effective in the treatment of arthritis, leukemia and cancerous tumours^{16,17}.

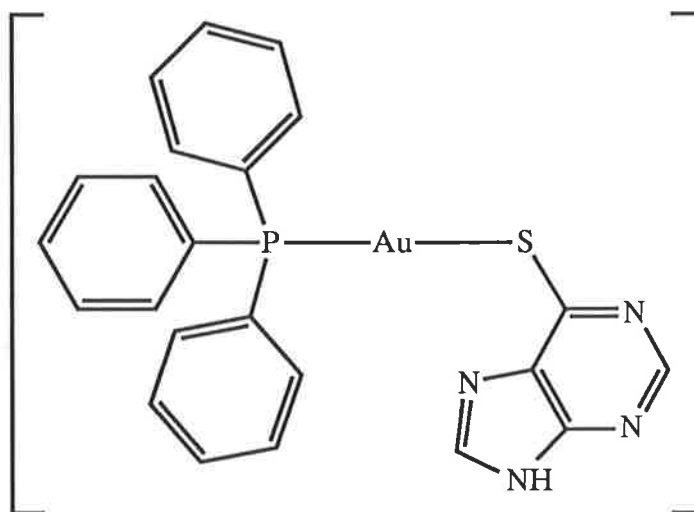


Figure 1.5.3: 6-mercaptopurinato(triphenylphosphine)gold(I).

The choice of thiolate for all the complexes prepared in this thesis was thus chosen to be 6-mercaptopurine, and the phosphines chosen were: Et₃P, Cycl₃P, PhMe₂P, Ph₃P, (*o*-Tol)₃P, (*m*-Tol)₃P and (*p*-Tol)₃P, to give complexes of the type [R₃PAu(6-MP)]; and [μ -bis(diphenylphosphino)methane], [μ -1,2-bis(diphenylphosphino)ethane] and [μ -1,3-bis(di-

phenylphosphino)propane], to give mono- and di-substituted compounds of the type shown schematically in Figure 1.5.4.

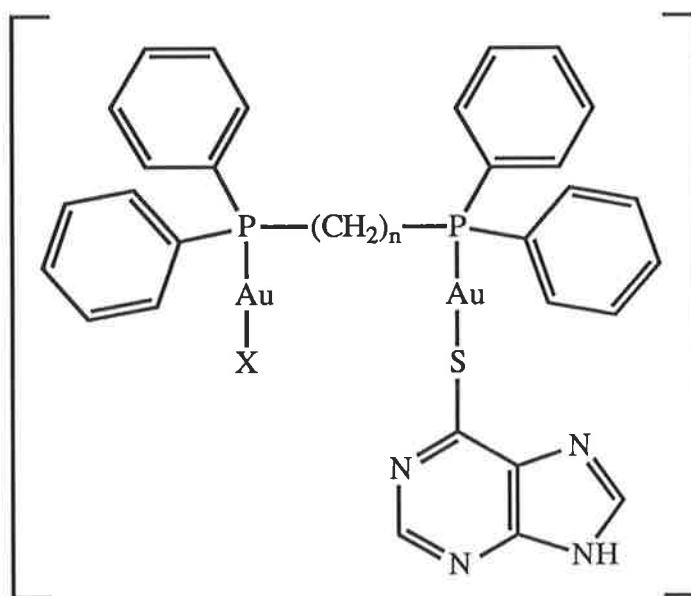


Figure 1.5.4: Diagram For Mono- And Di-substituted Complexes.
 $n = 1, 2$ or 3 ; $X = Cl$ For Mono-substituted, $X = 6\text{-MP}$ For Di-substituted.

1.6 The chemistry of gold and triorganophosphinegold(I) thiolates

Gold can exist in a variety of oxidation states, but it is dominated in its chemistry by the two most stable states, I and III¹⁸. Gold(III) has the electronic configuration $[\text{Xe}]4f^{14}5d^8$, and is thus isoelectronic with platinum(II). This would suggest that gold(III) complexes analogous to cisplatin might have comparable anti-tumour activity. Some complexes have proven active¹, but the reducing mammalian environment tends to reduce the gold(III) complexes, and so they have not been found to be generally effective. Gold(I), with the stable 'filled shell' electronic configuration of $[\text{Xe}]4f^{14}5d^{10}$, is thus better suited to the body's biochemistry, a fact which is illustrated by the variety of anti-arthritis gold(I) compounds already mentioned.

Both gold(III) and gold(I) are soft metal ions, with gold(I) more so, and these ions prefer to bind to soft donor atoms such as sulphur and phosphorus. Whereas a 'naked' gold(I) ion would be transmuted to gold(0) and gold(III) in the body, a gold(I) atom stabilized by sulphur will be less susceptible to oxidation and reduction. The phosphorus of the phosphine ligand is a σ -electron donor, and is also thought to be a π -electron acceptor from the $5d^{10}$ orbital of the gold(I) atom. Thus, both sulphur and phosphorus help to stabilize the gold(I) species.

Triorganophosphinegold(I) thiolates possess certain structural characteristics. Thiolate groups containing one exocyclic sulphur tend to coordinate in a monodentate mode via the sulphur atom to the gold centre. Other possible coordinating atoms or functional groups of the thiolate can associate with the gold but only through secondary interactions. One common feature is the tendency of nitrogen-containing thiolates to orient themselves in the crystal lattice such that the nitrogen is in close proximity to the gold centre. The intramolecular distance is invariably less than the sum of the van der Waals radii. The main feature of these complexes is the presence of a P–Au–S chromophore, found from crystallographic studies to be near linear in most examples regardless as to the identity of the phosphine or thiolate. These structural characteristics were looked for in the crystal structures reported in this thesis, and how they compare to other gold(I) thiolate complexes.

The compounds studied in this thesis were prepared via an established method, involving first the synthesis of a triorganophosphinegold(I) chloride precursor. This is detailed in Chapter 2; briefly, the procedure involved the reduction of an aqueous gold(III) chloride solution, by using thiodiglycol, to give an aqueous gold(I) species, which was then reacted *in situ* with the stoichiometric amount of the desired phosphine. The resultant air-stable triorganophosphinegold(I) chloride was then reacted in an equimolar, metathetical reaction with 6-mercaptopurine in the presence of a base.

1.7 Discussion of results

The focus of this thesis is on the preparation, spectroscopic characterization and, where appropriate, X-ray crystal structure determination, of series of complexes with the general formulae of $[R_3PAu(6-MP)]$ (where $R_3P = Et_3P, Cyc_3P, PhMe_2P, Ph_3P, (o-Tol)_3P, (m-Tol)_3P$ or $(p-Tol)_3P$), $[(Ph_2P(CH_2)_nPPh_2)(AuCl)(Au(6-MP))]$ (where $n = 2$ or 3), and $[(Ph_2P(CH_2)_nPPh_2)(Au(6-MP))_2]$ (where $n = 1, 2$ or 3).

The thesis is divided into three main sections. The first is the experimental section, in which the preparative methods employed will be detailed, along with the instrumentation used. The next two sections are devoted to the spectral characterization of the complexes prepared and of

the triorganophosphinegold(I) chloride precursors that were involved. As little information exists in the literature regarding the characterizations of triorganophosphinegold(I) chlorides, this data has been collected here. The X-ray crystal structure of $[\text{PhMe}_2\text{PAuCl}]$ is presented, followed by a study to find a possible correlation between the phosphine cone-angle and the Au-Cl bond length in triorganophosphinegold(I) chloride complexes. The effects on the 6-mercaptapurine moiety upon complexation to gold are deduced from spectroscopic and crystallographic analyses. The methods utilized were Fast Atom Bombardment mass spectrometry, ^1H , ^{13}C and ^{31}P nuclear magnetic resonance spectroscopy and Fourier Transform infrared spectroscopy. X-ray crystallographic techniques were performed on two complexes, $[\text{Ph}_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$ and $[(o\text{-Tol})_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$, in order to confirm the precise nature of the products and to determine what effect the Au-S coordination has on the electronic structure of the 6-mercaptapurinate moiety.

The X-ray crystal structure determination of the complex 6-n-propyl-2-thiouracilato(tricyclohexylphosphine)gold(I), $[\text{Cycl}_3\text{PAu}(6\text{p}2\text{-TU})]$, is given, which, along with data from a variety of other triorganophosphinegold(I) thiolates, is utilized in a cone-angle correlation to determine the effect of the phosphine ligand on the P-Au-S chromophore.

A selection of the complexes prepared were assessed for their anti-arthritis activity using the commercially available drugs as standards. The results ^{of} these tests are given, where available, with comments on the possible structure / activity correlations. The conclusion summarizes what the characterizations and crystallographic analyses have revealed about the new complexes.

CHAPTER 2

Experimental

2.1 Introduction

This chapter comprises three main sections: 1) a description of the experimental methods used in the preparations of the triorganophosphinegold(I) chloride and 6-mercaptopurinate complexes, including the melting points, yields and microanalytical data - examples for the preparation of $[\text{Ph}_3\text{PAuCl}]$ and $[\text{Ph}_3\text{PAu}(6\text{-MP})]$ are given; 2) brief notes on the reagents and the instrumentation used for the characterizations; and 3) a general discussion of the instrumentation used and the methods employed in solving the crystal structures detailed in later chapters via an example, namely $[(o\text{-Tol})_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$.

2.2 Preparation of the complexes

The preparation of all the complexes was based on a procedure adapted from ref. [19] for the preparation of the compound 2-thiouracilato(triphenylphosphine)gold(I). The procedure involved the equimolar metathetical reaction between triphenylphosphinegold(I) chloride and 2-thiouracil in the presence of a base in an ethanolic solution. This reaction has since been found to be generally applicable for the preparation of analogous complexes containing a variety of phosphines and thionucleobases¹⁰, and so was utilized to prepare the complexes reported in this thesis. General reaction schemes for the preparation of the complexes with the general formulae of $[\text{R}_3\text{PAu}(6\text{-MP})]$ (where $\text{R}_3\text{P} = \text{Et}_3\text{P}, \text{Cycl}_3\text{P}, \text{PhMe}_2\text{P}, \text{Ph}_3\text{P}, (o\text{-Tol})_3\text{P}$,

(*m*-Tol)₃P or (*p*-Tol)₃P), [(Ph₂P(CH₂)_nPPh₂)(AuCl)(Au(6-MP))] (where n = 2 or 3) and [(Ph₂P(CH₂)_nPPh₂)(Au(6-MP))₂] (where n = 1, 2 or 3) are shown in Figure 2.2.1.

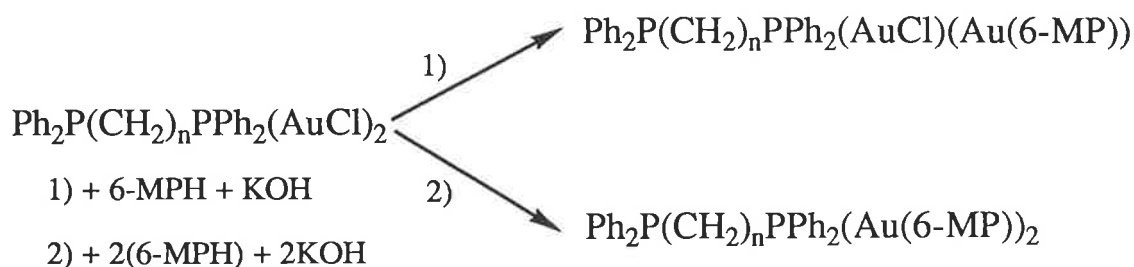


Figure 2.2.1: General Reaction Schemes For The Preparation Of The Complexes.

In this thesis, for reasons of brevity those compounds of the general formulae [(Ph₂P(CH₂)_nPPh₂)(AuCl)(Au(6-MP))] and [(Ph₂P(CH₂)_nPPh₂)(Au(6-MP))₂] will be given the abbreviations for the Ph₂P(CH₂)_nPPh₂ phosphine ligand of dppm, dppe and dppp for n = 1, 2 or 3, respectively. A detailed example of a preparation is given later.

The triorganophosphinegold(I) chloride complexes of the general formulae [R₃PAuCl] (where R₃P is as defined previously) and [Ph₂P(CH₂)_nPPh₂(AuCl)₂] were prepared by a procedure based on an example found in the literature^{20,21}. This method basically involves the reduction of tetrachloroauric acid ([HAuCl₄].3H₂O)²² by thiodiglycol followed by the addition of the desired phosphine species in molar quantities. What follows is an example of one such preparation.

Example: Preparation of [Ph₃PAuCl].

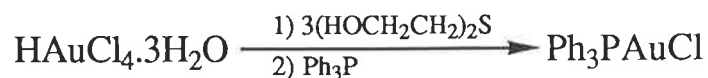


Figure 2.2.2: Reaction Scheme For The Preparation Of [Ph₃PAuCl].

To a stirred solution of acetone (1 cm³) and H₂O (3 cm³) held in an ice bath under a nitrogen atmosphere was added HAuCl₄.3H₂O (0.600 g, 1.52 mmol). Thiodiglycol (0.56 g, 4.57

mmol) was then added dropwise to the yellow-orange solution over a period of 2 h. Extra drops were added until a clear solution was obtained. The solution was then filtered under nitrogen to remove undissolved solids. Ph_3P (0.417 g, 1.52 mmol) dissolved in hot acetone (*ca* 5 cm^3) was added dropwise over a period of 10 min to the stirred solution. The resultant white solid that formed immediately was collected via vacuum filtration and washed with a small quantity of acetone. The solid product was air-dried for 5 min and recrystallized from hot ethanol. The resultant crystalline material was dried for 12 h over anhydrous phosphorus pentoxide under vacuum. Yield = 0.678 g; % yield = 90.0%; m.p. (dec.) = 233 - 234° C.

All the other compounds were prepared in an analogous manner using the appropriate molar quantities. The only variation to this procedure was for the preparation of $[\text{Et}_3\text{PAuCl}]$; Et_3P is a liquid at room temperature, and so was added as such via a syringe over a 5 min period.

The transparent crystalline products are all air stable at room temperature. For the most part, this procedure gave good yields: percentage yields based on 0.500 g $\text{HAuCl}_4 \cdot 3\text{H}_2\text{O}$ used are shown in Table 2.2.1, along with melting points and the corresponding literature values where available.

Example: *Preparation of $[\text{Ph}_3\text{PAu}(6\text{-MP})]$.*



Figure 2.2.3: *Reaction Scheme For The Preparation Of $[\text{Ph}_3\text{PAu}(6\text{-MP})]$.*

To a stirred ethanolic solution (*ca* 30 cm^3) of $[\text{Ph}_3\text{PAuCl}]$ (0.200 g, 0.405 mmol) and 6-mercaptopurine (0.062 g, 0.405 mmol) aqueous potassium hydroxide (0.200 mol dm^{-3} , 0.405 mmol) was added dropwise over a few minutes. After 15 min of stirring, a pale yellow solid started to form in the clear solution. The solution was left to stir for a further 1 h, then left in a fumehood until the solvent evaporated. The solid residue was dissolved with stirring into boiling acetone (*ca* 100 cm^3), then vacuum filtered to remove undissolved solids. The filtrate was left to stand until the acetone evaporated, and the off-white product was

Table 2.2.1: Melting Points And Yields For Triorganophosphinegold(I) Chloride Complexes.

Compound	m.p. (°C)	yield (g)	% yield	literature values (°C)
[Et ₃ PAuCl]	68 - 69	0.405	91.0	84 - 86 [23]
[Cycl ₃ PAuCl]	117 - 118	0.379	58.2	-
[PhMe ₂ PAuCl]	132 - 133	0.308	65.4	-
[Ph ₃ PAuCl]	242 - 243	0.590	94.0	242 [24]
[(<i>o</i> -Tol) ₃ PAuCl]	(274 - 275)	0.619	90.8	283 - 285 [25]
[(<i>m</i> -Tol) ₃ PAuCl]	152 - 153	0.590	86.6	-
[(<i>p</i> -Tol) ₃ PAuCl]	(186 - 187)	0.578	84.7	-
[dppm(AuCl) ₂]	270 - 271	0.407	75.5	273 [26]
[dppe(AuCl) ₂]	(288 - 289)	0.565	96.4	290 - 292 [27]
[dppp(AuCl) ₂]	(255 - 256)	0.358	64.3	256 - 257 [28]

Note: Brackets around the melting point value indicates decomposition point.

recrystallized twice from a small quantity of 1:1 ethanol / dichloromethane to give a pale yellow microcrystalline product. Yield = 0.231 g; % yield = 93.7%; m.p. (dec.) = 254 - 255° C.

This procedure was adapted to prepare all the complexes for this project. They are all air-stable solids, and vary in colour from very pale to bright yellow. The melting points, % yields (based on 0.200 g of $[R_3PAuCl]$ or $[Ph_2P(CH_2)_nPPh_2(AuCl)_2]$ used), colour and microanalytical results are shown in Tables 2.2.2 and 2.2.3.

2.3 Instrumentation

All melting points were determined using a Gallempkamp melting point apparatus calibrated with benzil.

Infrared spectra for all complexes were recorded on a Perkin-Elmer 1720X FT spectrometer calibrated with the polystyrene absorption at 1601 cm^{-1} , as KBr discs in the range of $400 - 4000\text{ cm}^{-1}$.

Proton and carbon-13 NMR spectra were recorded on a Bruker ACP-300 NMR spectrometer with d_6 -dimethylsulphoxide as the solvent. The recording frequencies used were 300.13 MHz for 1H NMR and 75.47 MHz for ^{13}C NMR. The internal reference used was $SiMe_4$ (TMS). Phosphorus-31 NMR spectra were recorded on a Bruker CXP-300 NMR spectrometer at 121.5 MHz, also as d_6 -dimethylsulphoxide solutions, with the ~~internal~~^{external} reference being 85% H_3PO_4 in D_2O .

FAB mass spectra were obtained using a VG ZAB-2HF spectrometer. The excitation gas was argon at a source pressure of typically 10^{-6} mbar. The FAB voltage was 7 kV with a current of 1 mA, the ion accelerating potential being 8 kV. A drop of a *ca* 0.5 mol dm^{-3} solution of the complex in dichloromethane was added to a drop of 3-nitrobenzyl alcohol matrix and applied to the probe tip. The spectra were recorded as a mass to charge ratio, m/z. Relative abundance was

Table 2.2.2: Melting Points, Yields And Colours For The Triorganophosphinegold(I) 6-mercaptapurinate Complexes.

Compound	m.p. (°C)	yield (g)	% yield	colour
[Et ₃ PAu(6-MP)]	102-103	0.264	99.2	pale yellow
[Cycl ₃ PAu(6-MP)]	139-140	0.228	93.2	light yellow
[PhMe ₂ PAu(6-MP)]	173-175	0.212	80.6	yellow-green
[Ph ₃ PAu(6-MP)]	254-255	0.231	93.7	off white
[(<i>o</i> -Tol) ₃ PAu(6-MP)]	(253-254)	0.235	96.7	pale yellow
[(<i>m</i> -Tol) ₃ PAu(6-MP)]	108-109	0.238	97.9	pale yellow
[(<i>p</i> -Tol) ₃ PAu(6-MP)]	105-106	0.238	98.1	pale yellow
[dppe(AuCl)(Au(6-MP))]	(183-185)	0.215	94.8	pale yellow
[dppp(AuCl)(Au(6-MP))]	(203-204)	0.212	93.6	pale yellow
[dppm(Au(6-MP)) ₂]	179-180	0.189	74.3	off white
[dppe(Au(6-MP)) ₂]	150-151	0.209	82.4	off white
[dppp(Au(6-MP)) ₂]	161-162	0.211	83.5	off white

Note: Brackets around the melting point value indicates decomposition point.

Table 2.2.3: Microanalytical Data For The Triorganophosphinegold(I) 6-mercaptopurinate Complexes.

Compound	%C _{calc}	%C _{found}	%H _{calc}	%H _{found}
[Et ₃ PAu(6-MP)]	28.33	28.19	3.89	4.06
[Cycl ₃ PAu(6-MP)]	43.95	44.02	5.77	6.03
[PhMe ₂ PAu(6-MP)].0.5EtOH	33.02	33.15	3.36	3.11
[Ph ₃ PAu(6-MP)]	45.26	45.19	2.97	2.98
[(<i>o</i> -Tol) ₃ PAu(6-MP)]	47.86	47.71	3.71	3.90
[(<i>m</i> -Tol) ₃ PAu(6-MP)]	47.86	47.82	3.71	3.94
[(<i>p</i> -Tol) ₃ PAu(6-MP)].H ₂ O	46.86	46.85	3.71	3.81
[dppe(AuCl)(Au(6-MP))]	38.03	38.12	2.78	2.75
[dppp(AuCl)(Au(6-MP))].H ₂ O	38.02	37.87	3.09	3.08
[dppm(Au(6-MP)) ₂].2H ₂ O	37.63	37.72	2.87	2.46
[dppe(Au(6-MP)) ₂].2H ₂ O	38.24	38.15	3.03	2.60
[dppp(Au(6-MP)) ₂].CH ₂ Cl ₂	38.24	38.07	2.87	2.63

Note: Microanalysis performed by Chemical And Microanalytical Services Pty. Ltd.

calculated by designating the most abundant peak as 100% and determining the abundance of the other peaks based on their relative heights in the spectra to this peak.

2.4 Chemicals

The chemicals utilized in the synthesis of the complexes and their sources were: triphenylphosphine (B.D.H.), triethylphosphine (Fluka), tricyclohexylphosphine (Strem), tri(*o*-tolyl)phosphine (Aldrich), tri(*m*-tolyl)phosphine (Aldrich), tri(*p*-tolyl)phosphine (Aldrich), [μ -bis-(diphenylphosphino)methane] (Strem), [μ -1,2-bis(diphenylphosphino)ethane] (Strem), [μ -1,3-bis(diphenylphosphino)propane] (Strem), thiodiglycol (Aldrich) and 6-mercaptopurine (Sigma). All the solvents employed were of analytical grade.

2.5 Crystallography: Instrumentation and methods

All the crystals were grown from the vapour diffusion of diethyl ether into an ethanolic solution of the compound. The colourless crystals were collected and their identities confirmed by spectral and melting point comparisons with the bulk material. The dimensions and crystallographic parameters are given in later chapters where the results are presented. The following description concerns the data collection and structure solution for [(*o*-Tol)₃PAu(6-MP)].C₂H₅OH; the methods utilized for the other crystal structure determinations were similar.

The crystal was mounted on a glass fibre using cyanoacrylate glue, and then placed on the goniometer of a Rigaku AFC6R diffractometer fitted with graphite-monochromatized MoK α radiation, $\lambda = 0.71073 \text{ \AA}$. The unit cell dimensions were determined from the least squares refinement of 25 well-centred reflections in the range of $7.7 < \theta < 12.8^\circ$, and, with the aid of Delauney reduction and a Laue symmetry check, was found to be monoclinic primitive with the Laue class of 2/m. Intensity data was then collected at 23° C in the ranges of $0 < h < 12$, $0 < k < 13$ and $0 < l < 33$, to a maximum Bragg angle of $2\theta = 55.8^\circ$. The intensities of 6280 reflections were measured using the $\omega:2\theta$ scan technique, of which 5913 were unique. The

value for R_{amal} was 0.025, where

$$R_{\text{amal}} = \frac{\sum_{i=1}^n \sum_{j=1}^m | \langle F^2_i \rangle - F^2_{ij} |}{\sum_{i=1}^n m \times \langle F^2_i \rangle}$$

n = the number of unique reflections that were observed more than once

m = the number of times a given reflection is observed

$\langle F^2_i \rangle$ = the average value of F^2 for the unique reflection i

No significant decomposition of the crystal occurred during the data acquisition. The data set was processed and corrected for Lorentz and polarization effects, and the space group determined to be $P2_1/n$, based on the analysis of systematic absences. A total of 4183 reflections satisfied the criterion of observability of $I \geq 3.0\sigma(I)$ and were used in the subsequent analysis.

The structure was solved by direct methods using the SHELXS86²⁹ program and refined by a full matrix least-squares procedure based on F^2 . The function minimized was

$$\sum_{i=1}^n w_i (|F_{\text{obs}}|_i - |F_{\text{calc}}|_i)^2$$

where n is the number of reflections, and

$$w = \frac{1}{\sigma^2(F_{\text{obs}})}$$

The non-hydrogen atoms were refined with anisotropic thermal parameters, and the hydrogen atoms were included in the model in calculated positions of C–H = 0.97 Å and N–H = 0.95 Å. The absorption correction used was DIFABS³¹. A weighting scheme was introduced based on sigma weights. At convergence, the values of R and R_w were 0.040 and 0.041 respectively,

where

$$R = \frac{\sum_{i=1}^n (|F_{obs}|_i - |F_{calc}|_i)}{\sum_{i=1}^n |F_{obs}|_i}$$

and

$$R_w = \left\{ \frac{\sum_{i=1}^n w_i (|F_{obs}|_i - |F_{calc}|_i)^2}{\sum_{i=1}^n w_i |F_{obs}|_i^2} \right\}^{1/2}$$

The maximum and minimum residual electron density peaks in the final difference map were 1.25 and -1.46 e Å⁻³. Scattering factors for all the atoms were those incorporated in the teXsan software package³⁰ which was installed on a Silicon Graphics Indigo computer system. Tables of bond distances and bond angles, fractional atomic coordinates, anisotropic thermal and hydrogen atom parameters and mean plane data are found in Chapter 3 for [PhMe₂PAuCl], Chapter 5 for [Ph₃PAu(6-MP)].C₂H₅OH and [(*o*-Tol)₃PAu(6-MP)].C₂H₅OH and Chapter 6 for [Cycl₃PAu(6p2-TU)]. Structure factors for all determinations are located in the Appendix. The expression for the anisotropic thermal parameter of the non-hydrogen atoms is

$$T_{\text{aniso}} = \exp[-2\pi^2(h^2a^2U_{11} + k^2b^2U_{22} + l^2c^2U_{33} + 2hka*b*U_{12} + 2hla*c*U_{13} + 2klb*c*U_{23})]$$

and the expression for B(eq) for the hydrogen atoms is

$$B(\text{eq}) = 8\pi^2(U_{11} + U_{22} + U_{33})/3.$$

CHAPTER 3

Characterization of Triorganophosphinegold(I) Chloride and [μ -1,n-bis(diphenylphosphino)alkane]bis(gold(I) chloride) Complexes

3.1 Introduction

This chapter discusses the characterization of the complexes of the general formula $[R_3PAuCl]$ (where $R_3P = Et_3P, Cycl_3P, PhMe_2P, Ph_3P, (o-Tol)_3P, (m-Tol)_3P$ or $(p-Tol)_3P$) and those of the general formula $[(Ph_2P(CH_2)_nPPh_2)(AuCl)_2]$ (where $n = 1, 2$ or 3). While complexes of this type are well known, documentation of their spectroscopic characteristics is not readily available and hence this Chapter. The discussion consists of three main parts: 1) the spectral characterization via infrared and multinuclear nmr techniques, 2) the crystal structure determination of the $[PhMe_2PAuCl]$ complex, and 3) a comparative study of triorganophosphinegold(I) chlorides found in the literature with respect to phosphine cone-angles and the P–Au and Au–Cl bond lengths. The last study will aim to determine whether a correlation exists between the cone angles of the triorganophosphine ligands and the lengths of the P–Au and Au–Cl bonds. The spectral characterization will be discussed in terms of how the data relate to analogous information found in the literature, what it describes about the chemical nature of these complexes, and how the information is useful for the characterization of the triorganophosphinegold(I) 6-mercaptopurinate complexes discussed in Chapter 4.

3.2 Spectral characterization of the triorganophosphinegold(I) chloride complexes

3.2.1 Infrared spectroscopy

The spectra were recorded as described in Chapter 2. The appearance of all the spectra were fairly simple, containing a few strong absorptions with a number of very weak ones. These absorptions are due to the vibrational modes associated with the phosphine ligands, since the modes involving the gold(I) and chloride atoms, such as $\nu(\text{P-Au})$ and $\nu(\text{Au-Cl})$, occur below 400 cm^{-1} and were not recorded owing to the limitations of the instrumentation. Table 3.2.1 lists the major absorptions found in each spectrum and the functional group vibrations to which they have been assigned. Infrared studies of these compounds in the literature are usually concerned with the P-Au and Au-Cl stretching modes^{32,33}, so the assignments are based on comparisons with the spectra of free phosphines. Figures 3.2.1 a) - d) show the infrared spectra of some selected compounds, and it is clear that by comparing, for example, the spectra of $[\text{Et}_3\text{PAuCl}]$ and $[\text{Ph}_3\text{PAuCl}]$, the absorption peaks found at 1586 cm^{-1} and at 1480 cm^{-1} for $[\text{Ph}_3\text{PAuCl}]$ and not found for $[\text{Et}_3\text{PAuCl}]$ must be due to aromatic ring vibrations. The main peaks found below *ca* 1450 cm^{-1} are due to $\nu(\text{C-C})$, $\nu(\text{P-C})$ and $\delta(\text{C-H})$ vibrations. These latter absorptions are listed together, as it is not possible to assign them unambiguously due to their overlapping absorption ranges^{34,35}. Comparison with the infrared spectra of the free phosphines³⁶ shows that, for Ph_3P , the $\nu(\text{P-C})$ vibration occurs at 1430 cm^{-1} , but it is unclear how the stretching frequency of this bond is affected when the phosphorus atom binds to the gold(I) atom. The aromatic absorptions for Ph_3P ³⁶ occur at *ca* 1580 cm^{-1} and 1480 cm^{-1} , indicating a slight change upon coordination, related to the change in electron density about the gold centre. The $\nu(\text{C-H})$ absorptions for alkyl C-H groups are more intense than those for aryl C-H groups in the complexes, an observation that has been made before³⁴, which is due to the hydrogen atoms being more tightly bound to the aromatic rings than to alkyl carbon atoms. This is also the reason why the aryl C-H absorptions occur between 50 and 100 wavenumbers higher than alkyl C-H absorptions. The alkyl absorptions are most intense in the spectrum of

Table 3.2.1: Infrared Data For The Triorganophosphinegold(I) Chloride Complexes.

Complex	$\nu(\text{C-H})$	$\nu(\text{C=C})$	$\nu(\text{C-C}), \nu(\text{P-C}), \delta(\text{C-H})$
[Et ₃ PAuCl]	2962s, 2932m, 2905m, 2874m	-	1456s, 1413m, 1384m, 1259w,br,sh, 1044vs
[Cycl ₃ PAuCl]	2922vs, 2852s	-	1447s, 1176m, 1040w
[Ph ₃ PAuCl]	3071w, 3058w,br	1586w, 1480m	1435s, 1180w, 1103s
[(<i>o</i> -Tol) ₃ PAuCl]	3054w, 2971m, 3023w, 2930m	1589m, 1565w, 1468s	1448s, 1376m, 1163m, 1133s, 1070w
[(<i>m</i> -Tol) ₃ PAuCl]	3050w,br, 2911w	1593m, 1578w, 1478s	1447s, 1404m, 1383m, 1309w, 1108s, 1045w
[(<i>p</i> -Tol) ₃ PAuCl]	3014w, 2963w, 2917w	1598m, 1559w, 1498m	1447w,br, 1397m, 1384w, 1310w, 1188m, 1103vs
[PhMe ₂ PAuCl]	3056w, 3041w, 3023w, 2990m,sh, 2850m,br	1587w, 1573w, 1507w, 1490w, 1473w	1437s, 1426m, 1413s, 1384m, 1190m, 1161m, 1112s, 1073m
[dppm(AuCl) ₂]	3050w,br, 2916m, 2853w	1483w	1436vs, 1385m, 1185w, 1160w, 1103s, 1069m
[dppe(AuCl) ₂]	3053m,br, 2906m	1587w, 1573w, 1482m	1435s, 1411m, 1173s, 1106vs, 1071m
[dppp(AuCl) ₂]	3049m,br, 2917m, 2901m, 2851m	1586m, 1571w, 1481s	1434vs, 1405s, 1385s, 1184m, 1158m, 1104vs, 1069m

Note: Units are wavenumbers (cm⁻¹).

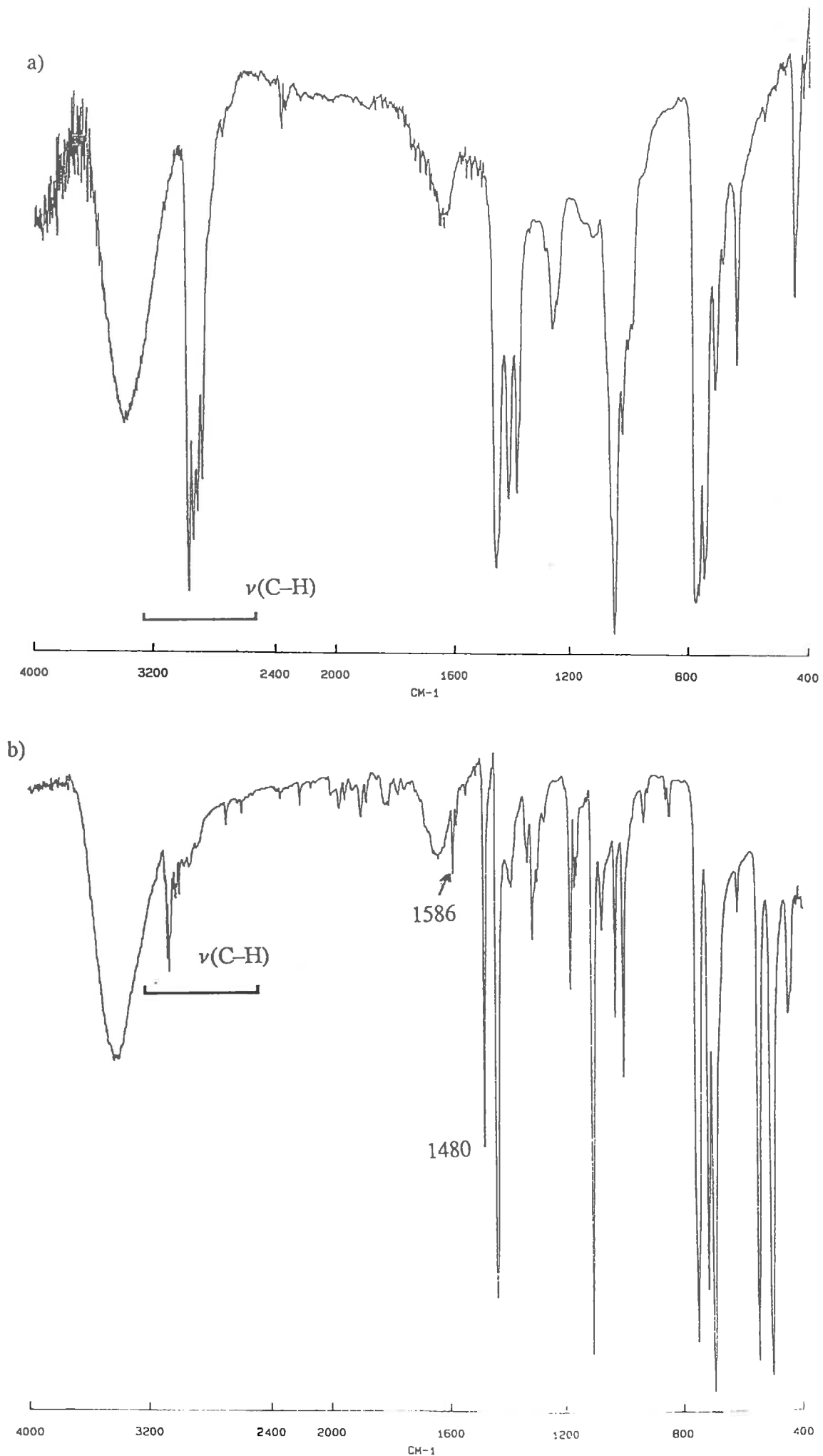


Figure 3.2.1: IR Spectra Of a) Triethylphosphinegold(I) Chloride, $[\text{Et}_3\text{PAuCl}]$,
b) Triphenylphosphinegold(I) Chloride, $[\text{Ph}_3\text{PAuCl}]$.

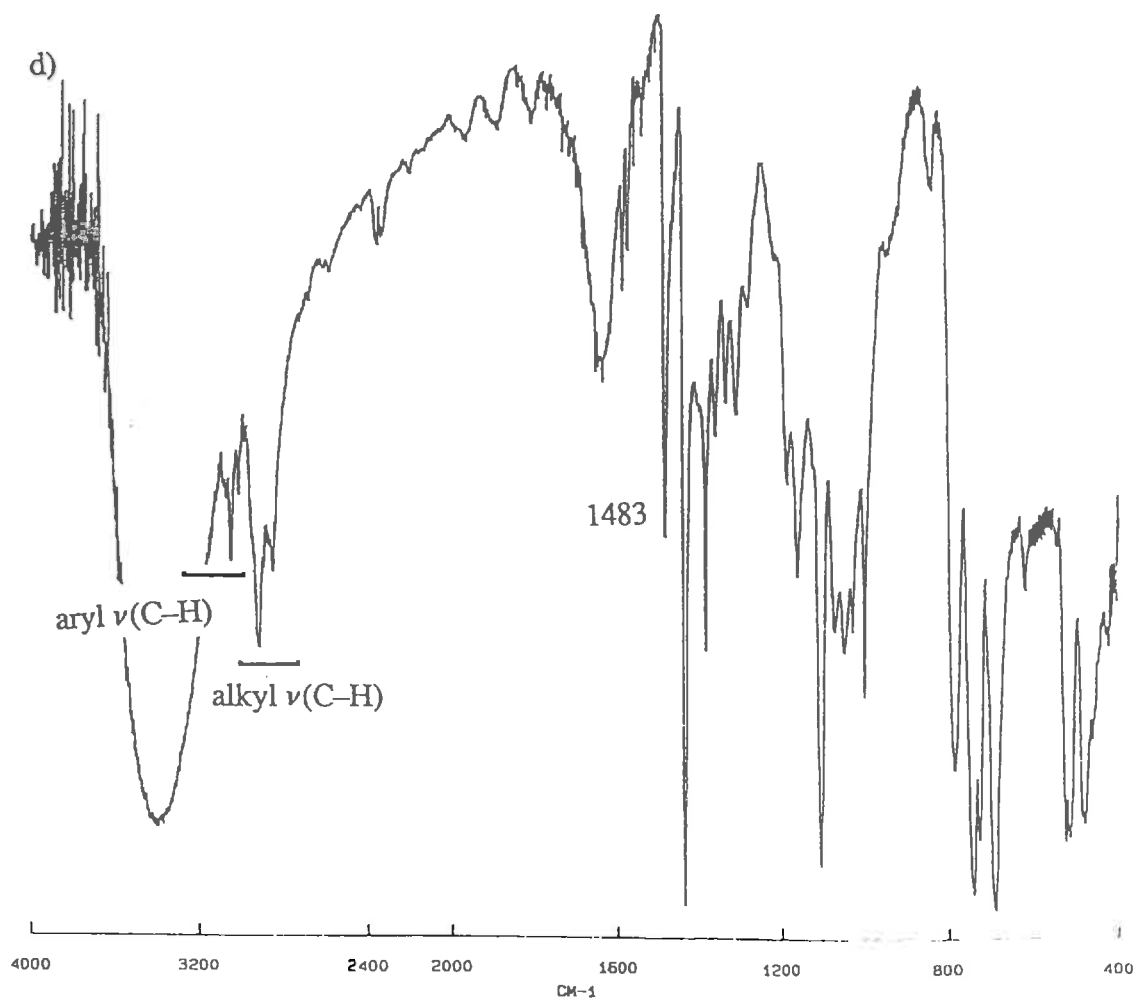
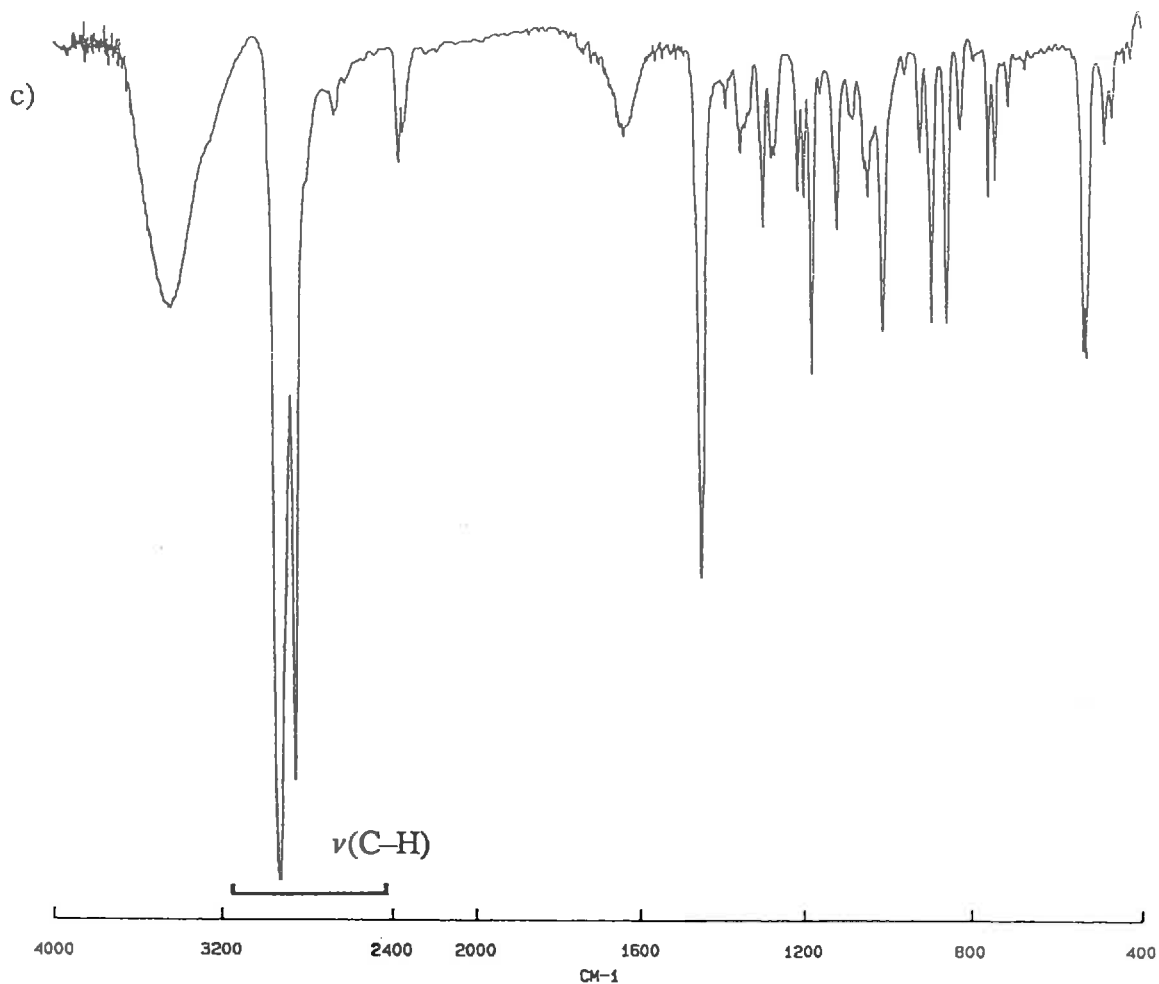


Figure 3.2.1: IR Spectra Of c) *Tricyclohexylphosphinegold(I) Chloride*, $[\text{Cycl}_3\text{PAuCl}]$,
 d) $[\mu\text{-bis(diphenylphosphino)methane}]_2\text{bis(gold(I)chloride)}$, $[(\text{Ph}_2\text{PCH}_2\text{PPh}_2)(\text{AuCl})_2]$,
 d) $[\mu\text{-bis(diphenylphosphino)methane}]_2\text{bis(gold(I)chloride)}$, $[(\text{Ph}_2\text{PCH}_2\text{PPh}_2)(\text{AuCl})_2]$

[Cycl₃PAuCl], in Figure 3.2.3. Figure 3.2.1 d) shows the relative intensities of aryl to alkyl absorptions in the [dppm(AuCl)₂] complex.

The data listed here will be useful in Chapter 4 as it can help assign those absorptions due to the phosphine group in the spectra of the complexes, simplifying the analysis. There are significant differences between the spectra of the triorganophosphinegold(I) chloride complexes and the free phosphines to suggest product formation. However, the respective absorptions between the gold(I) complexes are similar, indicating that the constituent absorptions are largely independent of the nature of the rest of the molecule, as expected.

3.2.2 ¹H NMR spectroscopy

The proton NMR spectra were obtained as described in Chapter 2. Although soluble in chloroform, for reasons of consistency the solvent used for these complexes was d₆-dimethylsulphoxide, as this was the solvent utilized for the 6-mercaptopurinate complexes in Chapter 4. Figure 3.2.2 a) shows the labeling scheme adopted for the carbon atoms and the protons bound to them.

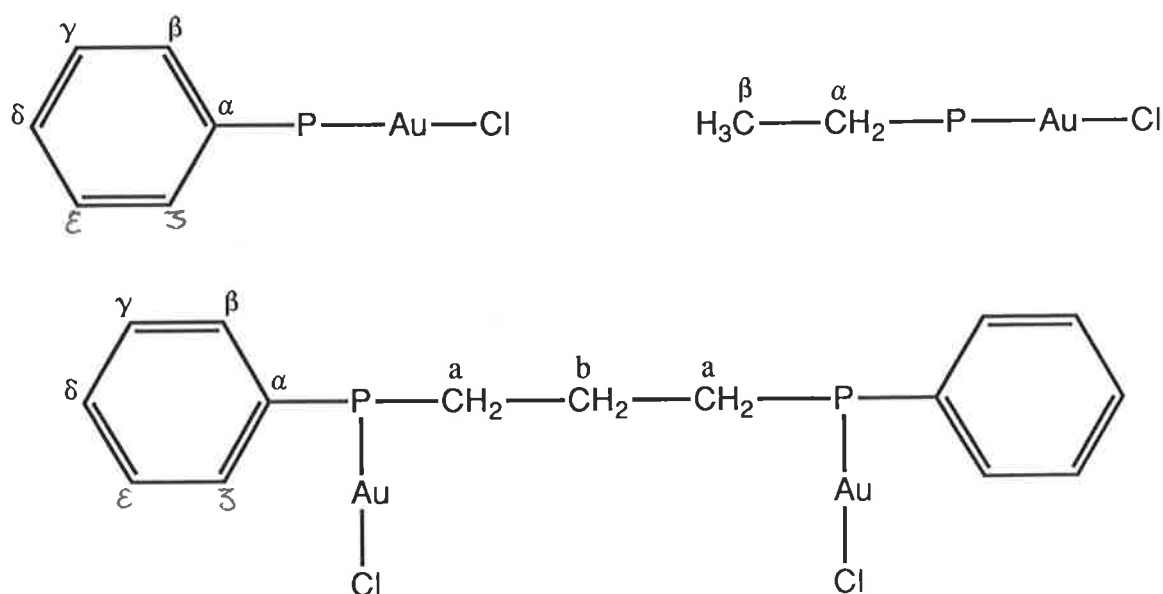


Figure 3.2.2 a): Labeling Scheme Adopted For NMR Assignments.

Table 3.2.2 lists the assignments for the protons in each complex, including the multiplicity and coupling constants. The methyl group of the tolyl phosphine complexes and of $[\text{PhMe}_2\text{PAuCl}]$ are designated by a CH_3 subscript. The resonances for aromatic protons occur as complex multiplets due to the complicated proton-proton and phosphorus-proton coupling combinations, so are given as a range. This range is typical of triphenylphosphine absorptions e.g. as found in $[\text{Ru}(6\text{-MP})_2(\text{PPh}_3)_2]^{2+}$ and Ph_3P ^{37,38}. These resonances occur downfield at *ca* 7.5 ppm because of the deshielding effect of the delocalized electrons in the aromatic moiety. As expected, there is little difference between these regions in the spectra of all the complexes containing this type of group. A similar but broader complex multiplet occurs for the cyclohexyl protons of $[\text{Cycl}_3\text{PAuCl}]$, also due to complex coupling patterns, but upfield due to their alkyl nature. Broad multiplets have been assigned to each proton type, but no coupling could be resolved.

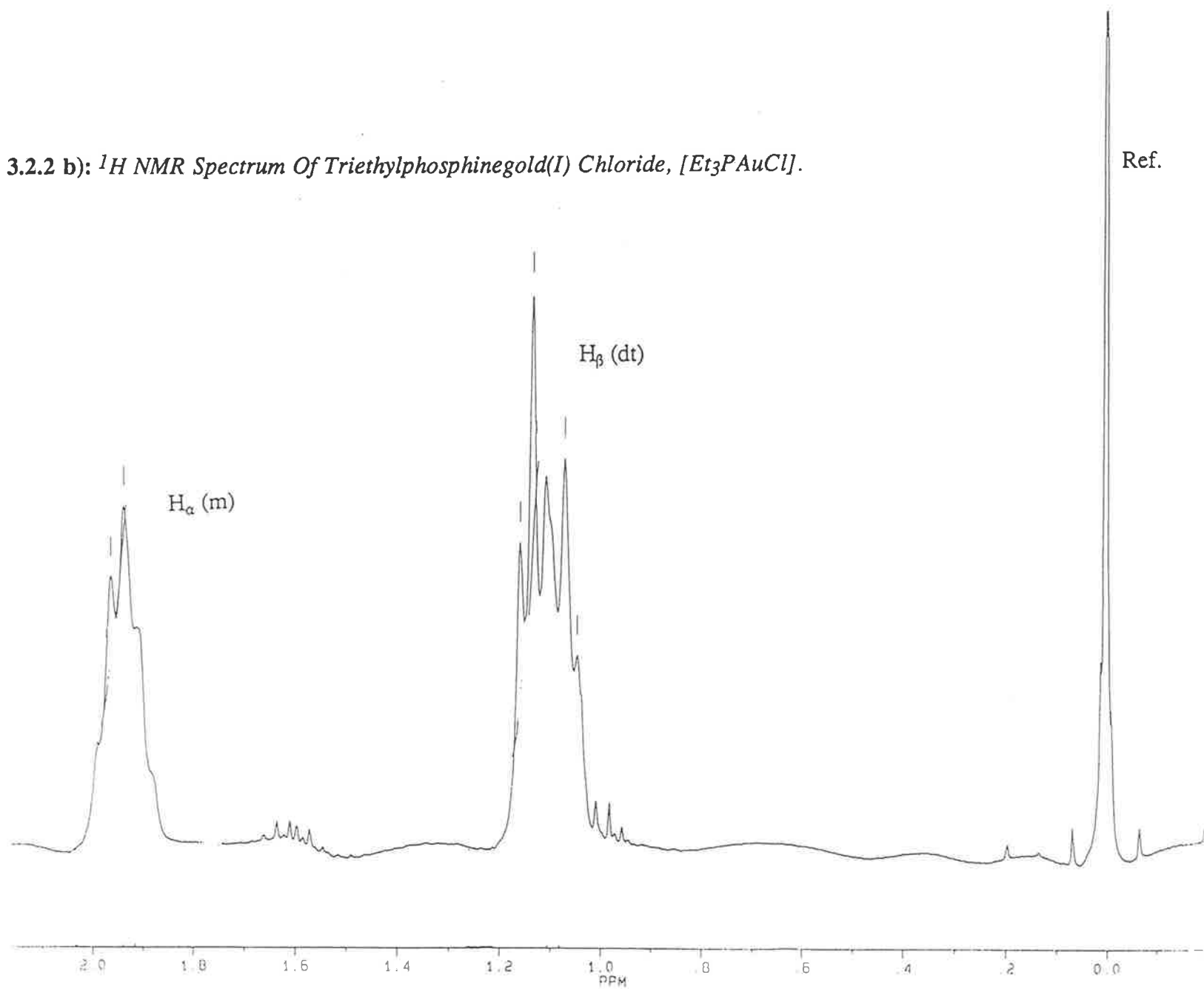
Two-bond indirect spin-spin coupling between the phosphorus atom and the protons was observed for the resonance of the alpha protons in $[\text{Et}_3\text{PAuCl}]$, giving rise to two overlapping quartets, where $^3J_{\text{HH}} = 7.71\text{Hz}$, but $^2J_{\text{PH}}$ was unresolved. The value of $^3J_{\text{PH}} = 18.89\text{Hz}$ for the beta protons results in a doublet of triplets, which is shown in Figure 3.2.2 b). The uniformity of the shapes of these resonances suggests free rotation about the C–C and P–C bonds in the phosphine groups. Phosphorus to proton coupling was manifested in the resonance of H_a for $[\text{dppe}(\text{AuCl})_2]$, appearing as a triplet due to coupling with the two magnetically equivalent phosphorus nuclei. It was not, however, distinguished for the H_a and H_b protons in the spectra of the $[\text{dppe}(\text{AuCl})_2]$ and $[\text{dppp}(\text{AuCl})_2]$ complexes, the resonances being broad and featureless. The reason for this is that the protons involved couple to two phosphorus atoms, and, in the case of H_a , the signs of the coupling constants are opposite, hence giving a net coupling which is unresolved in the spectra, and further complicated by proton–proton coupling. This corresponds with the resonance in the free phosphines, e.g. H_a in dppe is just resolved as a triplet centred at δ 2.10 ppm³⁸, compared to δ 2.97 ppm in $[\text{dppe}(\text{AuCl})_2]$. The latter observation is consistent with the deshielding effect expected at these nuclei upon complexation to the gold centre.

Table 3.2.2: ^1H NMR Chemical Shift Values (ppm) For Triorganophosphinegold(I) Chloride Complexes.

Complex	H_α	H_β	H_γ	H_δ	H_{-CH_3}	H_a	H_b
[Et ₃ PAuCl]	1.94(m) (7.71)	1.10(dt) (7.39) (18.89)	-	-	-	-	-
[Cycl ₃ PAuCl]	2.13(m)	1.79(m)	1.35(m)	1.26(m)	-	-	-
[Ph ₃ PAuCl]		7.63 - 7.51(m)			-	-	-
[(<i>o</i> -Tol) ₃ PAuCl]		7.64 - 6.87(m)			2.59(s)	-	-
[(<i>m</i> -Tol) ₃ PAuCl]		7.52 - 7.23(m)			2.34(s)	-	-
[(<i>p</i> -Tol) ₃ PAuCl]		7.42 - 7.39(m)			2.38(s)	-	-
[PhMe ₂ PAuCl]		7.86 - 7.56(m)			1.94(d) (11.35) ^c	-	-
[dppm(AuCl) ₂]		7.77 - 7.44(m)			-	4.70(t) (12.58) ^c	-
[dppe(AuCl) ₂]		7.79 - 7.59(m)			-	2.97(m)	-
[dppp(AuCl) ₂]		7.76 - 7.51(m)			-	3.05(m)	1.71(m)

Note: All coupling constants, in parentheses, are in units of hertz: a = $^3\text{J}_{\text{HH}}$, b = $^3\text{J}_{\text{PH}}$, c = $^2\text{J}_{\text{PH}}$.

Figure 3.2.2 b): ^1H NMR Spectrum Of Triethylphosphinegold(I) Chloride, $[\text{Et}_3\text{PAuCl}]$.



The data for the proton nmr studies on the complexes is consistent with literature values for the free phosphines and analogous complexes, and are useful for comparison with the 6-mercaptapurinate complexes in the next chapter.

3.2.3 ^{13}C $\{^1\text{H}\}$ NMR spectroscopy

The carbon-13 chemical shifts and phosphorus to carbon coupling constants are listed in Table 3.2.3. The phenyl-type carbon atoms are seen to resonate downfield, owing to the deshielding effect of the delocalized electron density in the ring. Of particular interest is the observation that the beta and gamma carbons in the $[\text{dppm}(\text{AuCl})_2]$ and $[\text{dppe}(\text{AuCl})_2]$ compounds appear as multiplets, where they would normally be expected to appear as doublets due to coupling with one phosphorus atom. This phenomenon has been noted before in the literature, the explanation being based on the fact that the two phosphorus atoms in the complexes are chemically equivalent but magnetically inequivalent due to the isotope effect (where a single ^{13}C atom in the molecule creates isotopic asymmetry)³⁹. Indeed, in the spectrum of $[\text{dppe}(\text{AuCl})_2]$, the beta and gamma carbon resonances resemble triplets, where the coupling value given represents the separation between the peaks. This structure can clearly be seen in Figure 3.2.3. As with the proton spectra, the C_α resonances are seen only as multiplets, but for the spectra of $[\text{dppp}(\text{AuCl})_2]$ a doublet of multiplets can be resolved. The problems with the resolution are more than likely due to the isotope effect.

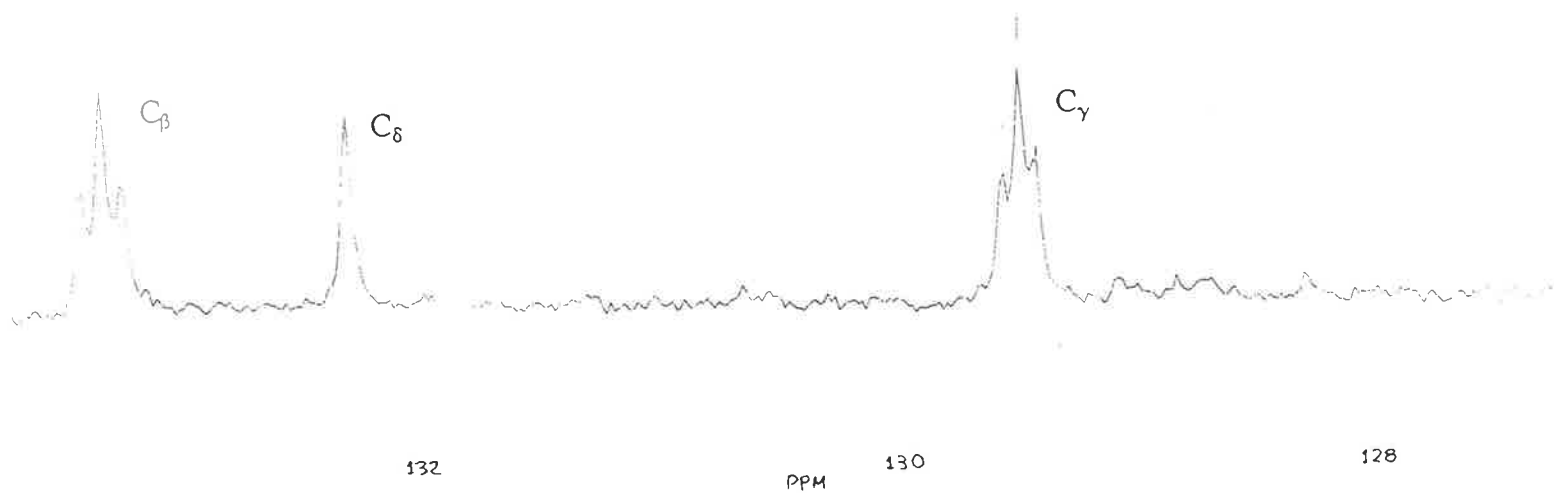
As expected, the phenyl-region chemical shifts for the $[(\mu-1, n\text{-}(\text{diphenyl-phosphino})\text{alkane})\text{bis}(\text{gold}(\text{I}) \text{chloride})]$ complexes are similar to those found for $[\text{Ph}_3\text{PAuCl}]$ and the tolyl compounds, the assignments of which were based on analogous complexes found in the literature. The alpha, beta, gamma and delta resonances for Ph_3P are 137.2, 133.6, 128.5 and 128.4 respectively⁴⁰, so it is apparent that coordination to the gold results in a major shift in the C_α shift value but only minor shifts in the others⁴¹. The former shift is likely to be a reflection of the change in electron density about the phosphorus atom when the P–Au bond is formed: the $^1\text{J}_{\text{PC}}$ value increases from 11.3 Hz in Ph_3P to 61.03 Hz in the $[\text{Ph}_3\text{PAuCl}]$ complex. Similar trends can be observed for the alpha carbon in all phenyl containing

Table 3.2.3: ^{13}C Chemical Shifts Values (ppm) For The Triorganophosphinegold(I) Chloride Complexes.

Complex	C_α	C_β	C_γ	C_δ	C_ϵ	C_ζ	-CH ₃	C_a	C_b
[Et ₃ PAuCl]	16.9(d) (36.23) ^a	9.05	-	-	-	-	-	-	-
[Cycl ₃ PAuCl]	32.3(d) (31.55) ^a	26.2(d) (12.08) ^b	30.3	25.7	30.3	26.2(d) (12.08) ^b	-	-	-
[Ph ₃ PAuCl]	128.3(d) (61.03) ^a	133.9(d) (13.36) ^b	129.8(d) (10.49) ^c	132.4	129.8(d) (10.49) ^c	133.9(d) (13.36) ^b	-	-	-
[(<i>o</i> -Tol) ₃ PAuCl]	124.2(d) (64.70) ^a	141.9 (11.93) ^b	132.6(d) (8.91) ^c	132.4	127.2(d) (10.57) ^c	133.0(d) (9.66) ^b	22.3(d) (11.02) ^c	-	-
[(<i>m</i> -Tol) ₃ PAuCl]	128.2(d) (61.36) ^a	139.2(d) (12.15) ^b	130.9(d) (12.83) ^c	133.1	129.6(d) (12.08) ^c	134.1(d) (14.87) ^b	21.0	-	-
[(<i>p</i> -Tol) ₃ PAuCl]	125.3(d) (64.53) ^a	133.7(d) (14.27) ^b	130.2(d) (12.15) ^c	142.5(d) (2.19)	130.2(d) (12.15) ^c	133.7(d) (14.27) ^b	21.0	-	-
[PhMe ₂ PAuCl]	Obs.	131.8(d) (12.19) ^b	129.1(d) (11.10) ^c	131.6(d)	129.1(d) (11.10) ^c	131.8(d) (12.19) ^b	14.4(d) (39.78) ^a	-	-
[dppm(AuCl) ₂]	128.7(m)	133.4(m)	129.3(m)	132.3	129.3(m)	133.4(m)	-	24.6(m)	-
[dppe(AuCl) ₂]	128.7(d) (58.60) ^a	133.3(t) (6.42)	129.5(t) (5.29)	132.2	129.5(t) (5.29)	133.3(t) (6.42)	-	22.4(m)	-
[dppp(AuCl) ₂]	128.8(d) (59.85) ^a	133.1(d) (13.36) ^b	129.5(d) (11.85) ^c	132.2	129.5(d) (11.85) ^c	133.1(d) (13.36) ^b	-	26.5(d) (52.49) ^a	20.3(m)

Note: ^{31}P - ^{13}C coupling constants, in parentheses, are in units of hertz: a = $^1\text{J}_{\text{PC}}$, b = $^2\text{J}_{\text{PC}}$ and c = $^3\text{J}_{\text{PC}}$.

Figure 3.2.3 : ^{13}C NMR Spectrum Of [μ -1,2-bis(diphenylphosphino)ethane]bis(gold(I) chloride), $[(\text{Ph}_2\text{P}(\text{CH}_2)_2\text{PPh}_2)(\text{AuCl})_2]$.



phosphines in going from the free phosphine to the complex, although the same was not necessarily true for the proton resonances. It should be noted that for [(*o*-Tol)₃PAuCl] and [(*m*-Tolyl)₃PAuCl] the assignments made for the six aromatic carbons are tentative due to the closeness of the resonance values⁴⁰. Assignments in the spectrum of [Cycl₃PAuCl] are based on the data for the free phosphine^{42,43}; the C_α chemical shift remains static, but the ¹J_{PC} value increases from 18.6 Hz to 31.55 Hz in the complex. The other ring carbons altered very little in terms of chemical shifts and coupling constants. A similar result is observed for Et₃P and [Et₃PAuCl]⁴². This suggests that the chemical shifts of phosphorus-bound carbons in an aromatic environment are more sensitive to coordination effects at the phosphorus atom than are alkyl-type carbons.

The spectroscopic data confirm the stoichiometries of the triorganophosphinegold(I) chlorides and provide an essential reference for the interpretation of the spectra for the triorganophosphinegold(I) 6-mercaptopurinate complexes.

3.3 Crystal structure determination of the [PhMe₂PAuCl] complex

The general details concerning the data collection procedure used in the crystal structure determination of this complex have already been given in Chapter 2. What follows here is a discussion of the unit cell and the molecular structure obtained by refinement of this data.

Crystals of the complex were grown from the slow evaporation of a concentrated ethanolic solution of the compound. The [PhMe₂PAuCl] complex crystallizes in the achiral space group *P*2₁2₁2₁ (*D*₂⁴, No. 19)⁴⁴ and the absolute configuration was determined on the basis of the differences between high-angle Friedel pairs included in the data set. Crystal and refinement data are listed in Table 3.3.1 and crystallographic results are summarized in Tables 3.3.2 to 3.3.7. A list of the observed and calculated structure factors is given in the Appendix. The crystallographic numbering scheme drawn with the ORTEP⁴⁵ program is shown in Figure 3.3.1.

Table 3.3.1: Crystallographic Parameters for the [PhMe₂PAuCl] Complex.

Data	[PhMe ₂ PAuCl]
Formula	C ₈ H ₁₁ AuClP
Formula weight	370.6
Crystal shape	block
Crystal dimensions (mm)	0.27 x 0.14 x 0.14
Crystal system	orthorhombic
Space group	<i>P</i> 2 ₁ 2 ₁ 2 ₁ (<i>D</i> ₂ ⁴ , No. 19)
<i>a</i> (Å)	12.639(4)
<i>b</i> (Å)	16.931(6)
<i>c</i> (Å)	9.458(3)
<i>α</i> (°)	90
<i>β</i> (°)	90
<i>γ</i> (°)	90
<i>V</i> (Å ³)	2024(1)
<i>Z</i>	8
ρ_{calc} (g cm ⁻³)	2.432
<i>F</i> (000)	1360
μ (cm ⁻¹)	148.55
θ limits, cell (°)	8.0 to 12.8
θ limits, data (°)	1.5 to 25.4
<i>hkl</i> range	0 to 15, 0 to 20, 0 to 11
Range of transmission factors	0.945 to 1.058
Scan technique	$\omega:2\theta$
No. of data measured	3015
No. of unique data	2143
<i>R</i> _{amal}	0.046
No. of unique data used	1608
Criterion of observability	$I \geq 3.0\sigma(I)$
No. of parameters	199
<i>R</i>	0.035
<i>R</i> _w	0.039
Residual electron density (e Å ⁻³)	-0.99 to 0.94

Table 3.3.2: Fractional Atomic Coordinates For The [PhMe₂PAuCl] Complex.

Atom	x	y	z
Au(1)	0.54801(7)	0.35615(5)	0.52508(10)
Au(2)	0.46393(7)	0.32211(5)	0.21977(10)
Cl(1)	0.6889(5)	0.2949(4)	0.4228(8)
Cl(2)	0.4269(5)	0.4549(3)	0.1866(7)
P(1)	0.4152(5)	0.4157(4)	0.6343(6)
P(2)	0.4450(5)	0.1923(3)	0.2333(6)
C(12)	0.3771(17)	0.3718(15)	0.8006(27)
C(13)	0.2956(18)	0.4149(14)	0.5424(32)
C(22)	0.4590(19)	0.1540(13)	0.4053(21)
C(23)	0.5613(18)	0.1549(14)	0.1520(23)
C(111)	0.4417(18)	0.5158(11)	0.6767(18)
C(112)	0.3623(18)	0.5682(14)	0.7121(28)
C(113)	0.3843(22)	0.6416(17)	0.7472(31)
C(114)	0.4867(23)	0.6699(15)	0.7451(24)
C(115)	0.5677(20)	0.6198(15)	0.7089(27)
C(116)	0.5450(17)	0.5445(13)	0.6746(21)
C(211)	0.3342(18)	0.1439(15)	0.1591(25)
C(212)	0.3288(18)	0.0602(14)	0.1650(26)
C(213)	0.2458(20)	0.0246(12)	0.0926(34)
C(214)	0.1736(20)	0.0645(15)	0.0269(37)
C(215)	0.1819(19)	0.1470(17)	0.0182(33)
C(216)	0.2618(17)	0.1840(13)	0.0777(35)

Table 3.3.3: Anisotropic Thermal Parameters For The [PhMe₂PAuCl] Complex.

Atom	U ₁₁	U ₂₂	U ₃₃	U ₁₂	U ₁₃	U ₂₃
Au(1)	0.0389(5)	0.0414(5)	0.0407(5)	0.0026(5)	0.0041(6)	-0.0044(5)
Au(2)	0.0344(5)	0.0331(4)	0.0337(5)	0.0000(5)	-0.0012(5)	-0.0029(4)
Cl(1)	0.044(4)	0.054(4)	0.090(6)	0.007(3)	0.012(4)	-0.019(4)
Cl(2)	0.038(3)	0.031(3)	0.082(5)	0.001(3)	-0.004(4)	0.001(3)
P(1)	0.039(4)	0.049(4)	0.025(4)	-0.003(3)	0.005(3)	-0.001(3)
P(2)	0.032(3)	0.035(3)	0.027(3)	0.003(3)	-0.006(4)	-0.001(3)
C(12)	0.05(2)	0.07(2)	0.05(2)	-0.00(1)	0.01(2)	0.02(2)
C(13)	0.05(2)	0.07(2)	0.09(2)	0.00(1)	-0.00(2)	-0.06(2)
C(22)	0.07(2)	0.05(1)	0.02(1)	0.01(2)	0.02(1)	0.01(1)
C(23)	0.04(1)	0.05(2)	0.03(1)	0.00(2)	0.03(1)	0.00(1)
C(111)	0.04(1)	0.04(1)	0.003(9)	-0.01(1)	0.00(1)	0.007(9)
C(112)	0.05(2)	0.05(2)	0.03(2)	0.00(1)	0.02(2)	-0.00(2)
C(113)	0.08(2)	0.06(2)	0.06(2)	0.02(2)	0.04(2)	0.04(2)
C(114)	0.10(2)	0.06(2)	^{0.01(1)} 0.00(1)	-0.02(2)	0.00(1)	0.02(1)
C(115)	0.05(2)	0.08(2)	0.04(2)	-0.03(2)	-0.01(2)	-0.01(2)
C(116)	0.03(1)	0.05(1)	0.02(1)	-0.00(1)	-0.02(1)	-0.01(1)
C(211)	0.04(1)	0.04(1)	0.05(2)	-0.01(1)	-0.02(1)	0.01(1)
C(212)	0.05(2)	0.05(2)	0.05(2)	-0.00(1)	-0.02(2)	0.01(1)
C(213)	0.04(1)	0.03(1)	0.12(3)	-0.01(1)	-0.01(2)	-0.01(2)
C(214)	0.06(2)	0.05(2)	0.12(3)	-0.00(2)	-0.05(2)	-0.03(2)
C(215)	0.06(2)	0.08(2)	0.09(2)	0.00(2)	-0.05(2)	0.02(2)
C(216)	0.04(1)	0.03(1)	0.15(3)	0.01(1)	-0.06(2)	0.02(2)

Table 3.3.4: Hydrogen Atom Parameters For The [PhMe₂PAuCl] Complex.

Atom	x	y	z	B(eq)
H(12a)	0.3231	0.4041	0.8451	4.6
H(12b)	0.3493	0.3192	0.7840	4.6
H(12c)	0.4383	0.3685	0.8622	4.6
H(13a)	0.2435	0.4456	0.5940	5.8
H(13c)	0.2709	0.3609	0.5326	5.8
H(15b)	0.3059	0.4378	0.4494	5.8
H(22a)	0.3998	0.1709	0.4630	4.0
H(22b)	0.4605	0.0967	0.4013	4.0
H(22c)	0.5245	0.1731	0.4465	4.4
H(23a)	0.5604	0.0976	0.1553	3.8
H(23b)	0.5639	0.1722	0.0543	3.8
H(23c)	0.6230	0.1744	0.2019	3.8
H(112)	0.2892	0.5507	0.7112	4.3
H(113)	0.3272	0.6766	0.7751	5.7
H(114)	0.5012	0.7246	0.7689	5.1
H(115)	0.6403	0.6384	0.7081	5.1
H(116)	0.6021	0.5093	0.6478	3.6
H(212)	0.3804	0.0293	0.2170	4.6
H(213)	0.2419	-0.0326	0.0912	5.3
H(214)	0.1140	0.0376	-0.0160	7.1
H(215)	0.1283	0.1769	-0.0321	6.8
H(216)	0.2698	0.2404	0.0639	7.0

Table 3.3.5: Bond Distances (\AA) For The $[\text{PhMe}_2\text{PAuCl}]$ Complex.

Atom	Atom	Distance	Atom	Atom	Distance
Au(1)	– Cl(1)	2.277(6)	C(111)	– C(116)	1.39(3)
Au(1)	– P(1)	2.214(6)	C(112)	– C(113)	1.32(4)
Au(2)	– Cl(2)	2.273(5)	C(113)	– C(114)	1.38(3)
Au(2)	– P(2)	2.205(5)	C(114)	– C(115)	1.37(3)
P(1)	– C(12)	1.81(2)	C(115)	– C(116)	1.35(3)
P(1)	– C(13)	1.74(2)	C(211)	– C(212)	1.42(3)
P(1)	– C(111)	1.77(2)	C(211)	– C(216)	1.38(3)
P(2)	– C(22)	1.76(2)	C(212)	– C(213)	1.39(3)
P(2)	– C(23)	1.77(2)	C(213)	– C(214)	1.30(3)
P(2)	– C(211)	1.77(2)	C(214)	– C(215)	1.40(3)
C(111)	– C(112)	1.38(3)	C(215)	– C(216)	1.31(3)

Table 3.3.6: Bond Angles (°) For The [PhMe₂PAuCl] Complex.

Atom	Atom	Atom	Angle	Atom	Atom	Atom	Angle
Cl(1)	- Au(1)	= P(1)	177.2(3)	P(1)	= C(111)	= C(116)	120(2)
Cl(2)	- Au(2)	= P(2)	175.4(2)	C(112)	- C(111)	- C(116)	117(2)
Au(1)	- P(1)	= C(12)	114.9(9)	C(111)	- C(112)	- C(113)	121(2)
Au(1)	- P(1)	= C(13)	114.9(9)	C(112)	- C(113)	- C(114)	122(3)
Au(1)	- P(1)	= C(111)	113.4(8)	C(113)	- C(114)	- C(115)	119(2)
C(12)	- P(1)	= C(13)	102(1)	C(114)	- C(115)	- C(116)	119(2)
C(12)	- P(1)	= C(111)	104(1)	C(111)	- C(116)	- C(115)	122(2)
C(13)	- P(1)	= C(111)	106(1)	P(2)	= C(211)	= C(212)	119(2)
Au(2)	- P(2)	= C(22)	115.2(7)	P(2)	- C(211)	- C(216)	121(2)
Au(2)	- P(2)	= C(23)	111.6(8)	C(212)	- C(211)	- C(216)	119(2)
Au(2)	- P(2)	= C(211)	113.7(8)	C(211)	- C(212)	- C(213)	117(2)
C(22)	- P(2)	= C(23)	101(1)	C(212)	- C(213)	- C(214)	123(2)
C(22)	- P(2)	= C(211)	106(1)	C(213)	- C(214)	- C(215)	120(2)
C(23)	- P(2)	= C(211)	109(1)	C(214)	- C(215)	- C(216)	120(2)
P(1)	- C(111)	= C(112)	122(2)	C(211)	- C(216)	- C(215)	121(2)

Table 3.3.7: Mean Plane Data For The [PhMe₂PAuCl] Complex.

Plane number 1: Least-squares plane through the phenyl ring defined by atoms C(111) to C(116).

Atoms Defining Plane	Distance (Å)	esd (Å)
C(111)	0.0057	0.0171
C(112)	-0.0124	0.0263
C(113)	0.0081	0.0274
C(114)	-0.0008	0.0222
C(115)	0.0006	0.0255
C(116)	-0.0030	0.0200
Additional Atom	Distance (Å)	
P(1)	0.0323	

Mean deviation from plane is 0.0051 Å.

Chi-squared: 0.4.

Plane number 2: Least-squares plane through the phenyl ring defined by atoms C(211) to C(216).

Atoms Defining Plane	Distance (Å)	esd (Å)
C(211)	-0.0241	0.0246
C(212)	0.0434	0.0327
C(213)	-0.0144	0.0323
C(214)	-0.0305	0.0346
C(215)	0.0234	0.0301
C(216)	0.0022	0.0257
Additional Atom	Distance (Å)	
P(2)	0.1208	

Mean deviation from plane is 0.0230 Å.

Chi-squared: 4.3.

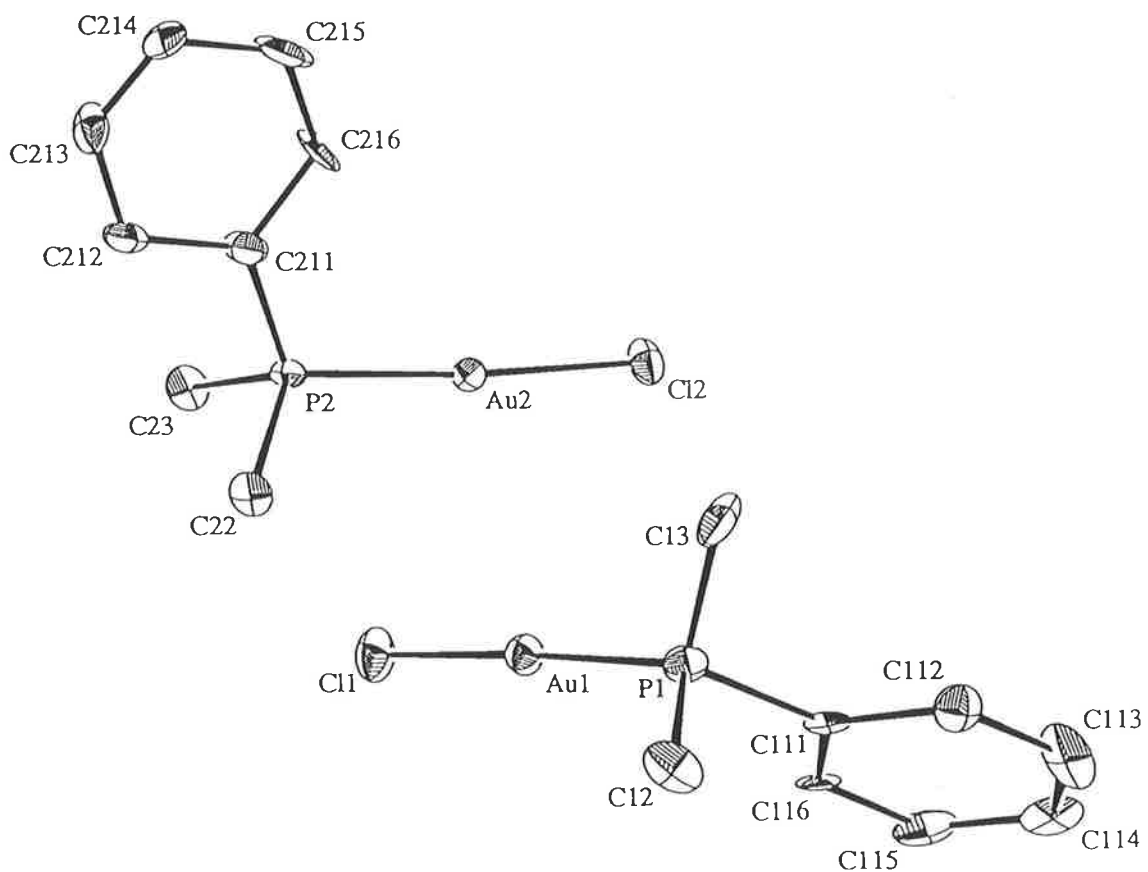


Figure 3.3.1: *Molecular Structure And Crystallographic Numbering Scheme For [PhMe₂PAuCl].*

A diagram of the unit cell is shown in Figure 3.3.2. The unit cell is comprised of eight molecules, the asymmetric unit therefore contains two molecules, labeled molecules 1 and 2 respectively; see Figure 3.3.1. The molecules are arranged in the unit cell in what appears to be a dimeric relationship, known as head-to-tail dimers. Such interactions have been observed before for triorganophosphinegold(I) chlorides where the phosphine is relatively small in size, and are attributable to the presence of close gold to gold interactions in the lattice⁴⁶. The Au(1)...Au(2) interaction is at a distance of 3.262(1) Å, which is less than the sum of the van der Waals radii of 3.40 Å, but is due more to relativistic effects than a significant bonding interaction⁴⁶.

Both Au(1) and Au(2) exist in the expected linear geometry, clearly demonstrated in Figure 3.3.1, with P–Au–Cl angles of 177.2(3) and 175.4(2)° respectively, comparable with 179.63(8)° for [Ph₃PAuCl]⁴⁷. The respective P–Au bond distances are 2.214(6) and 2.205(5) Å, which are equivalent within standard deviation. The Au–Cl bond distances of 2.277(6) and 2.273(5) Å, respectively are also within standard deviation range. The bond distances of P(1)–C(111) and P(2)–C(211) are both 1.77(2) Å, a value lower but consistent with those observed for [Ph₃PAuCl]. The phosphorus to methyl group bond lengths in the range of 1.76(2) to 1.81(2) Å are comparable to those found for [Et₃PAuCl]⁴⁸. Analysis of Tables 3.3.5 and 3.3.6 reveals that all the remaining corresponding intramolecular parameters between molecules 1 and 2 are identical to within standard deviation.

The phenyl rings are planar for both molecules, with mean deviations from planarity of 0.01(2) and 0.02(3) Å for molecules 1 and 2 respectively. The internal carbon to carbon bond distances range from 1.32(4) to 1.39(3) Å for molecule 1 and from 1.30(3) to 1.42(3) Å for molecule 2. These values are typical for phenyl ring systems and are indicative of electron delocalization.

The main difference between the two molecules in the asymmetric crystallographic unit is in the geometry of the gold atoms as manifested by the respective values for the P–Au–Cl angle. This difference might be attributable to the presence of the dimer in the lattice, where the Au(1)...Au(2) interaction gives rise to unequal distortions in the ideal linear geometries about

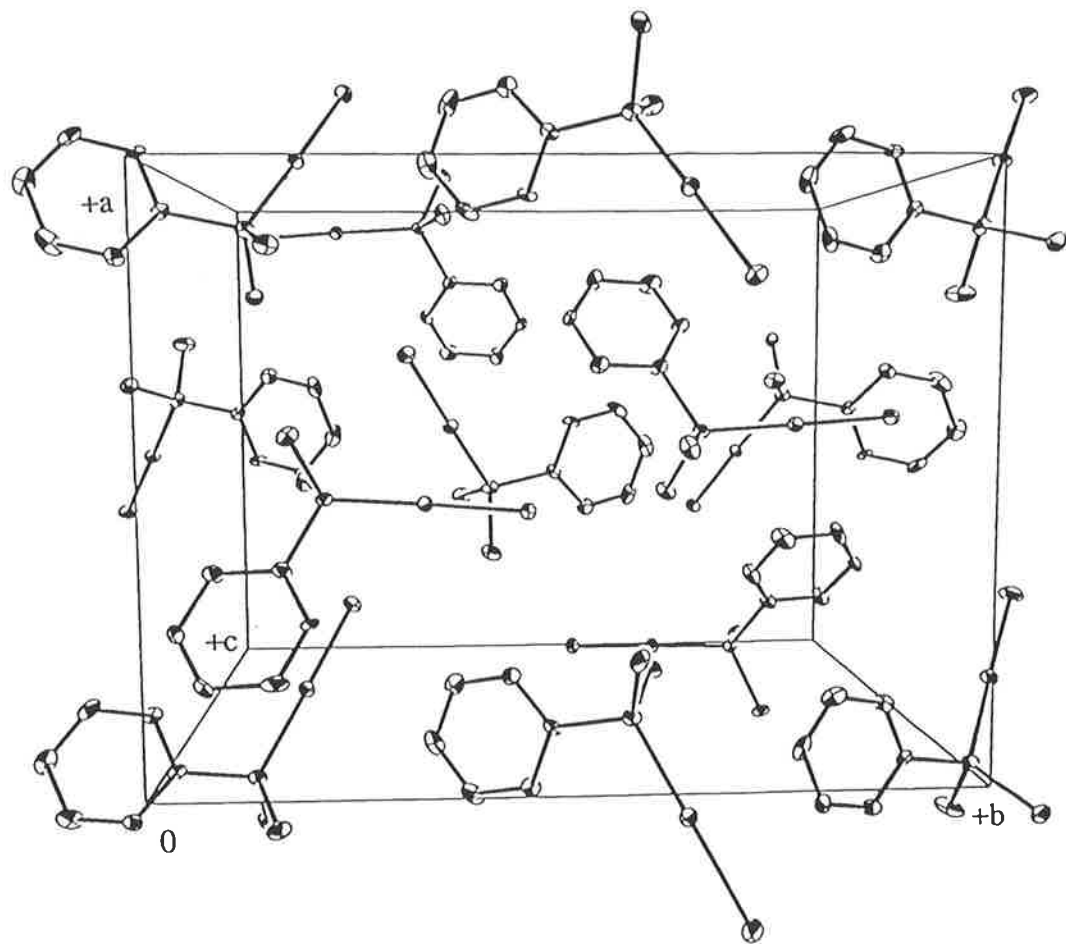


Figure 3.3.2: Unit Cell Diagram For [PhMe₂PAuCl].

the two gold atoms⁴⁶. The two independent molecules also differ from each other in the relative orientations of the phosphine-bound substituents, as shown in the values of the torsion angles Cl(n) / Au(n) / P(n) / C(n2), C(n3), C(n11) of -47, -165 and 72° for n = 1 and -178, 68 and -56° for n = 2, respectively. The next section investigates whether the size of the phosphine as represented by the cone-angle has any effect on the intramolecular parameters associated with the gold atom.

3.4 Cone-angle correlation for triorganophosphinegold(I) chloride complexes

The cone-angle of a phosphine is a structural property first introduced by Tolman⁴⁹, and is useful in defining the volume of space about a phosphine complex within which steric (van der Waals) interactions are likely to occur. Tolman's original work was centred on phosphinenickel carbonyl compounds, in which the nickel atom formed the fourth substituent on the phosphorus atom, to give this centre an approximately tetrahedral geometry. The substituents on the phosphorus can sweep out a maximum volume of space about the P–Ni axis in the shape of a cone whose apex lies on the metal centre, and the angle between the edge of this cone and a line extended from the P–Ni bond axis (the apical axis) is defined as half of the cone-angle⁵⁰. Cone-angles can usually be determined from crystallographic data, although for phosphines containing mixed substituents, further mathematical treatment is required, resulting in a weighted cone-angle. The Tolman cone-angles are hence dependent on the spatial size of the substituents and on the lengths of the P–Ni bonds. Tolman's analysis assumed this bond length to be of a constant value, but in reality the bulkiness of the substituents and the P–Ni bond length is likely to be interdependent⁴⁹.

A recent publication by Brown⁵¹ introduced a new parameter that complements the cone-angle concept. Known as the *ligand repulsive energy*, E_R , this quantity results from energy-minimization studies on phosphinechromium carbonyls, and represents the van der Waals interactions between the phosphine substituents and moieties along the P–Cr bond axis. The main advantage of the E_R quantity is that it distinguishes steric effects from electronic effects, and is more easily applied to phosphines bound to other elements. Using Tolman's calculated

cone-angles (θ) for the nickel compounds, Brown found that the E_R parameter correlated surprisingly well with the cone-angle; i.e. the relationship between E_R and θ was found to be linear for analogous Cr and Ni compounds to within a small margin of error⁵¹.

This study is concerned with the application of θ and E_R to triorganophosphinegold(I) chloride complexes of the general formula $[R_3PAuCl]$. Complexes of this type invariably possess a gold atom of linear geometry i.e. the P–Au–Cl bond angle is close to 180° . Hence the purely steric effects on the chlorine atom that arise from the size of the phosphine are likely to be minimized. Since the cone-angle parameter encompasses electronic effects also^{49,50} (the electronic nature of the groups bound to the phosphorus atom will influence the length of the phosphorus to metal bond and thus the magnitude of the derived or calculated cone-angle), then variations in the P–Au and Au–Cl bonds with a change in θ are likely to be attributable to electronic rather than steric effects. Therefore, a correlation may be drawn between these parameters. The quantity of E_R , which should describe only van der Waal or purely steric effects, should thus correlate less well with the intramolecular parameters if electronic effects are significant. The parameters for a range of triorganophosphinegold(I) chlorides have been collated from literature reports and are shown, together with calculated θ and E_R values, in Table 3.4.

Both θ and E_R have been plotted against the P–Au and Au–Cl bonds; the results are shown in Figure 3.4. At first inspection there seem to be no linear relationships present between the parameters and bond lengths. However, closer analysis of some of the plots reveals that a common trend does exist, albeit with some exceptions. The plot of P–Au versus cone-angle contains points that appear to form a linear distribution, with the anomalous points representing $[(PhO)_3PAuCl]$ and $[PhMe_2PAuCl]$. The analogous plot against E_R demonstrates the same linear appearance, with the same two exceptions. The values of P–Au for those complexes on the linear progression are in fact all equivalent within experimental error. Hence, there is no noticeable electronic effect due to the phosphine ligand on the length of P–Au. The values for θ can thus be regarded as absolute, similarly to Tolman's assumption of a constant P–Ni bond for various phosphine ligands. The plot of P–Au versus E_R displays a similar appearance to the

Table 3.4: Cone Angles (θ), Ligand Repulsive Energies (E_R) And Intramolecular Parameters Of The $[R_3PAuCl]$ Complexes.

R_3P	θ ($^\circ$)	E_R (kcal mol $^{-1}$)	P–Au–Cl ($^\circ$)	P–Au (\AA)	Au–Cl (\AA)	Ref.
Et $_3$ P	132	61	178.5(3)	2.232(9)	2.305(8)	[48]
			178.9(3)	2.231(8)	2.306(8)	
Cycl $_3$ P	170	116	177.0(2)	2.242(4)	2.279(5)	[52]
Ph $_3$ P	145	75	179.63(8)	2.235(3)	2.279(3)	[47]
Cycl $_2$ PhP	162	105	178.3(1)	2.234(2)	2.281(3)	[53]
(PhO) $_3$ P	128	65	178.5(2)	2.192(5)	2.273(5)	[54]
(<i>o</i> -Tol) $_3$ P	194	113	179.4(1)	2.243(2)	2.281(3)	[55]
(<i>m</i> -Tol) $_3$ P	145	79	175.1(1)	2.235(2)	2.288(2)	[56]
PhMe $_2$ P	122	44	177.2(3)	2.214(6)	2.277(6)	This work
			175.4(2)	2.205(5)	2.273(5)	

Note: There are two molecules in the asymmetric crystallographic units for each of $[Et_3PAuCl]$ and $[PhMe_2PAuCl]$.

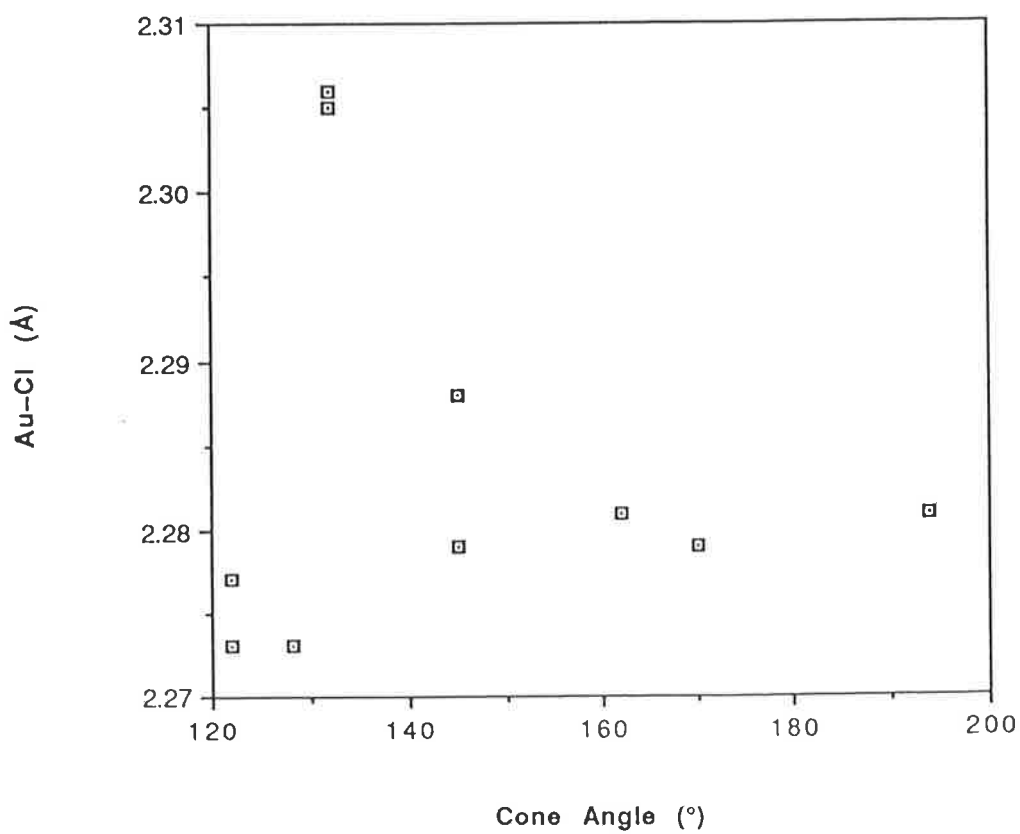
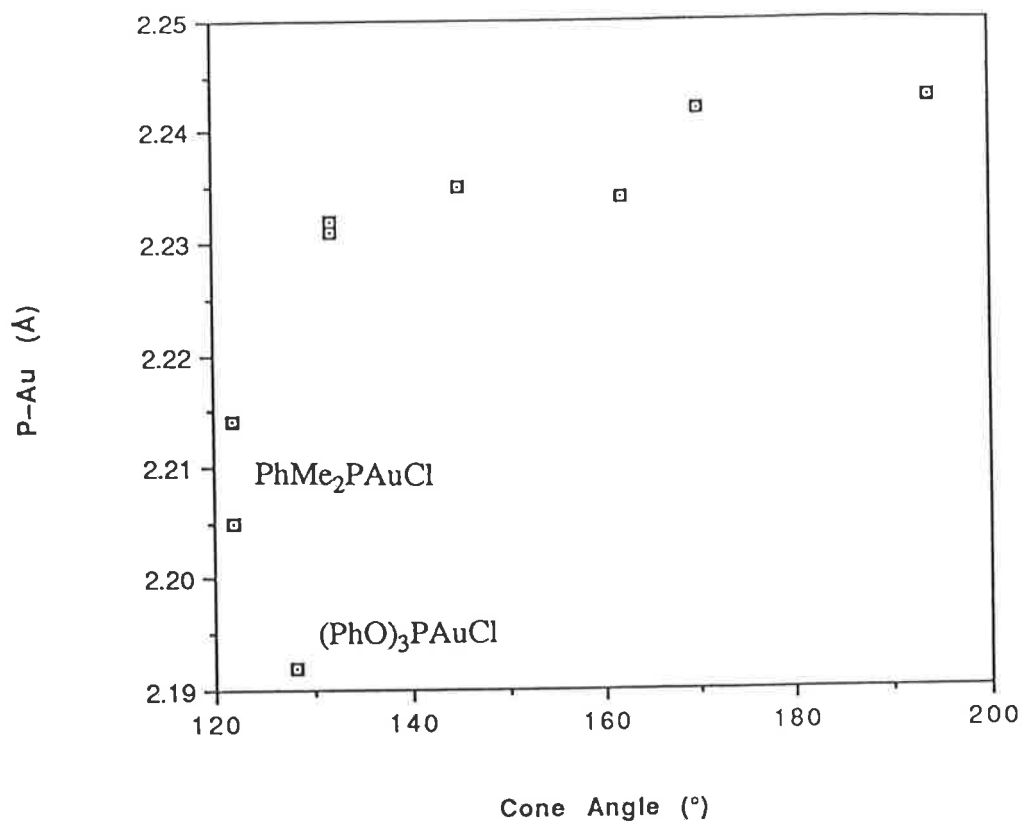


Figure 3.4: Plots Of The Parameters θ , And E_R Versus P-Au, Au-Cl And P-Au-Cl..

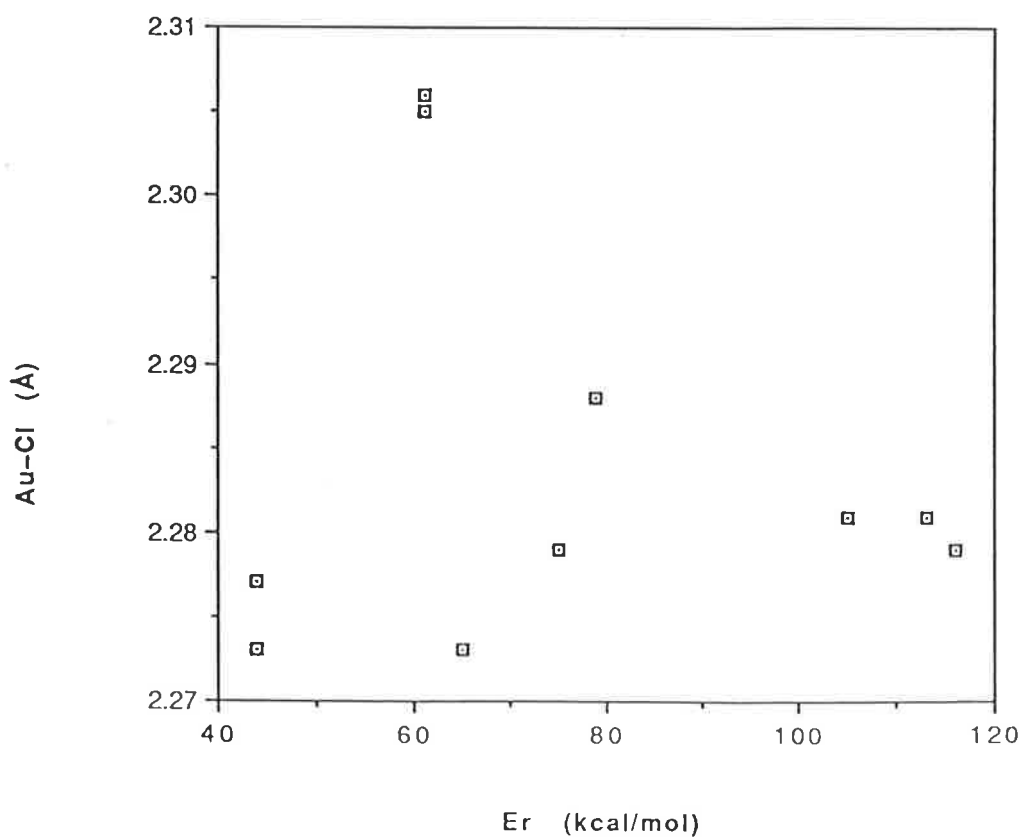
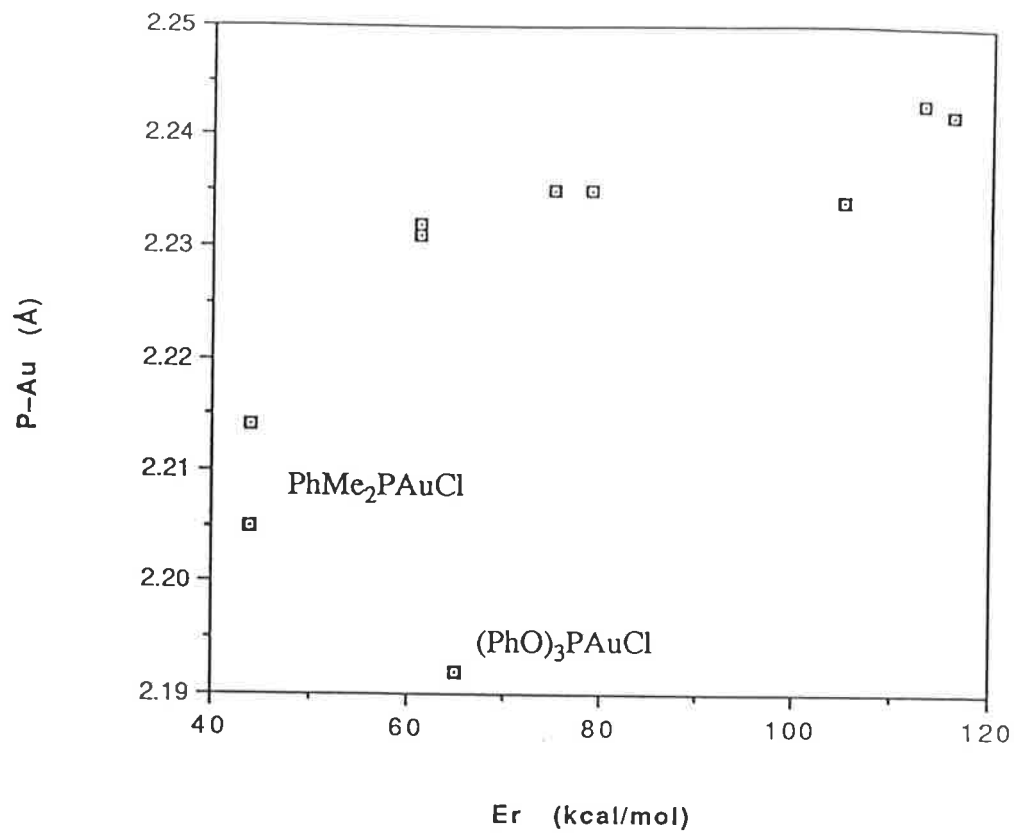


Figure 3.4 (continued)

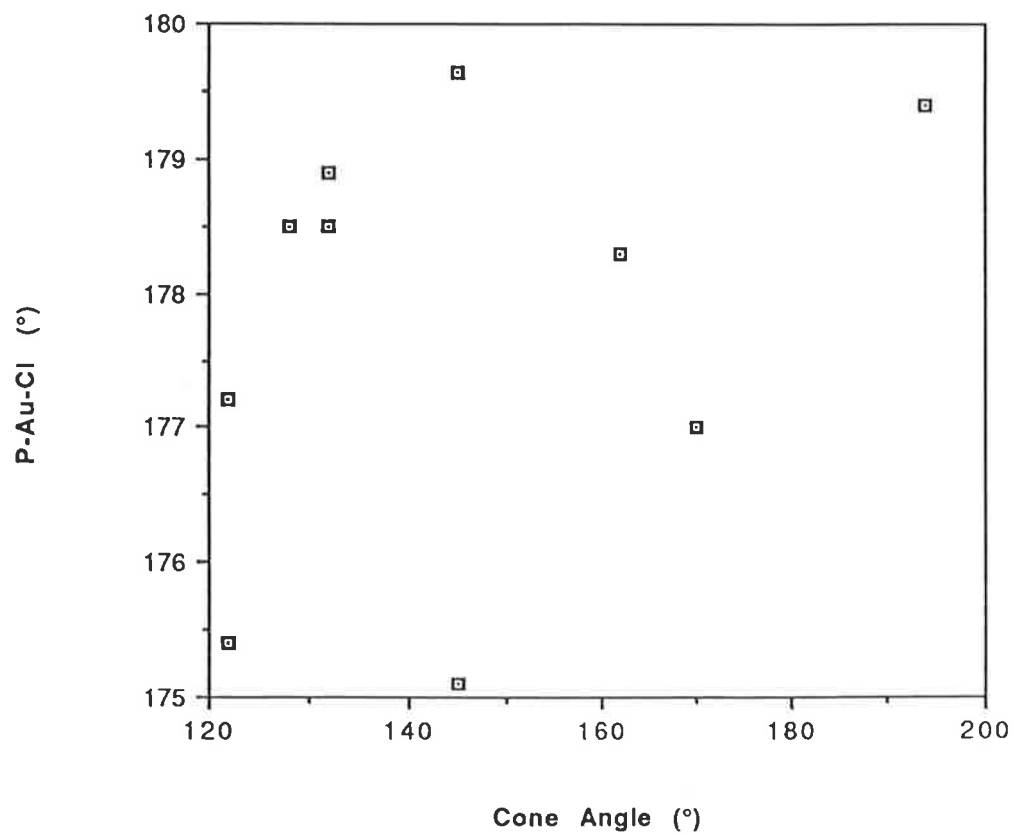
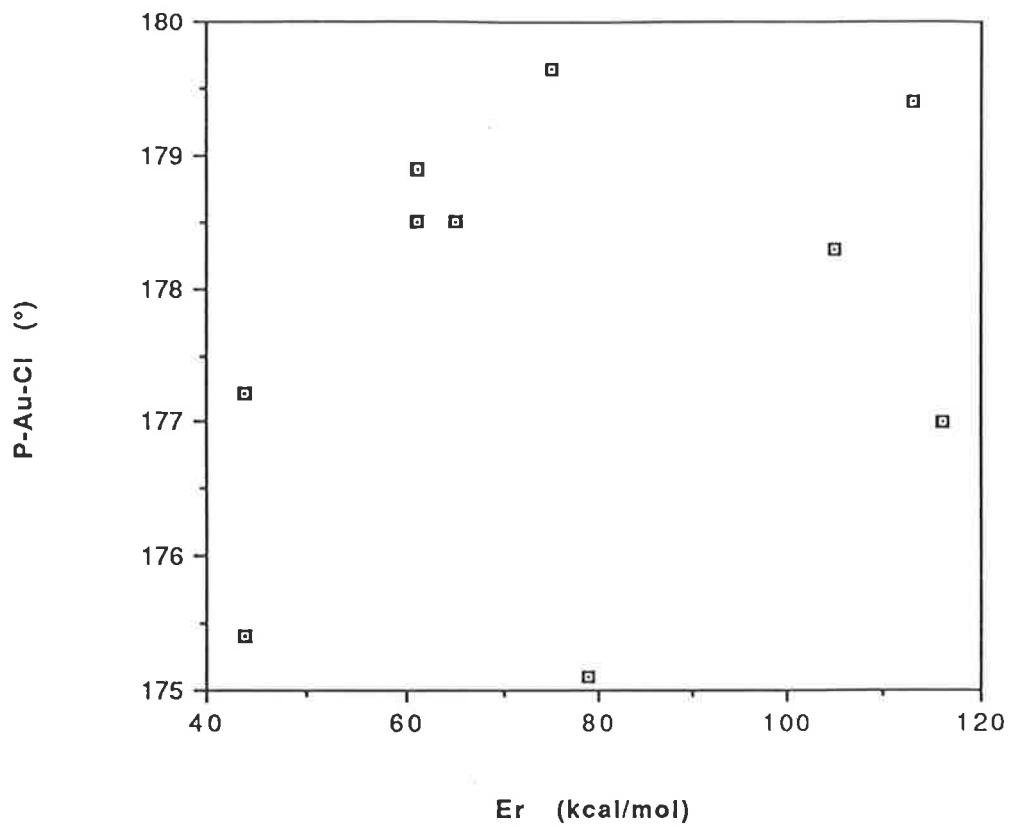


Figure 3.4 (continued)

P–Au versus θ plot, indicating that steric effects on the P–Au bond in the $[R_3PAuCl]$ complexes are also negligible.

The two exceptions to the general trends were found for the $[(PhO)_3PAuCl]$ and $[PhMe_2PAuCl]$ complexes. In the paper by Brown⁵¹, phosphite ligands were noted as exceptions due to the flexibility of C–O–P linkages, which can easily 'absorb' steric effects by adopting a variety of conformations not available to phosphines. Hence, neither θ nor E_R can adequately express the steric or electronic factors of the $(PhO)_3P$ ligand. $PhMe_2P$ is classed as a small phosphine, and gold(I) chloride complexes of these phosphines often crystallize in such an arrangement to allow the closest Au...Au interaction possible in the lattice, usually resulting in dimers⁴⁶. As discussed earlier in this chapter, the structure determination of $[PhMe_2PAuCl]$ revealed that such dimers were present, with an Au(1)...Au(2) interaction of 3.262(1) Å, less than the sum of the van der Waals radii (3.40 Å). Such a small distance is ~~often not~~ ^{not often} observed for complexes of larger phosphines, and so in this complex the intermolecular interaction may be a determining factor for the anomalous P–Au bond length. $[Et_3PAuCl]$ also contains a small phosphine⁴⁸, but the cone angle is considerably greater. The closest Au...Au interaction here is 3.615(2) Å, significantly larger than 3.40 Å, and $[Et_3PAuCl]$ is consistent with the linear trend of the graph. Thus, it can be postulated that, since $[PhMe_2PAuCl]$ is anomalous for both E_R and θ , electronic factors from intermolecular Au...Au interactions have ^a ~~s~~ significant influence on the length of the P–Au bond in complexes of small phosphines. Neither E_R nor θ are adequate quantities for accommodating such intermolecular interactions.

The Au–Cl bond length should manifest electronic effects of the phosphine ligand on the P–Au–Cl ~~chromophore~~ ^{moiety} more than steric effects, as it occupies a position *trans* to the phosphorus atom. The plots involving Au–Cl (i.e. Figure 3.4) show that there is no clear trend. In fact, all the values lie within experimental error. This indicates that, even in complexes with significant intermolecular interactions in the lattice, the Au–Cl bond length is largely independent of the steric and electronic effects of the phosphine ligand.

The P–Au–Cl bond angle displays a relatively large spread of values. However, there is no noticeable trend based on the nature of the phosphine ligand coordinated to the gold centre. The values for [(PhO)₃PAuCl] and [PhMe₂PAuCl] are not exceptional, indicating that conformations and intermolecular interactions do not affect this aspect of the chromophore in any special way.

Based on the examples studied, the important results are: 1) the P–Au bond length is independent of the steric and electronic effects of the phosphine ligand, but can reflect the electronic environment about the gold atom if significant intermolecular Au...Au interactions are present; 2) the Au–Cl bond length remains invariant, regardless of intermolecular interactions and the steric or electronic profiles of the phosphine ligand; 3) the P–Au–Cl bond angle shows a range of values, but this angle is not clearly related to any steric or electronic property of the phosphine ligand; and 4) the ligand repulsive energy, E_R , as calculated by Brown from energy-minimization models, has a good correlation with the cone-angle of phosphines but is no better than this angle for observing how the steric nature of a phosphine affects the intramolecular characteristics of triorganophosphinegold(I) chlorides.

CHAPTER 4

**Spectroscopic Characterization of the Triorganophosphinegold(I)
6-mercaptapurinate Complexes**

4.1 Introduction

In this chapter the results obtained from the spectroscopic analysis of the complexes with the general formulae $[R_3PAu(6-MP)]$ (where $R_3P = Et_3P, Cyc_3P, PhMe_2P, Ph_3P, (o-Tol)_3P, (m-Tol)_3P$ or $(p-Tol)_3P$), $[(Ph_2P(CH_2)_nPPH_2)(AuCl)(Au(6-MP))]$ (where $n = 2$ or 3) and $[(Ph_2P(CH_2)_nPPH_2)(Au(6-MP))_2]$ (where $n = 1, 2$ or 3) will be tabulated and discussed, with emphasis on how these results indicate product formation and what they reveal about the coordinated 6-mercaptapurinate ligand. A general introduction concerning the 6-mercaptapurine ligand itself and its chemistry as related to this project will be given first to help illustrate the points discussed later.

4.2 An overview of the 6-mercaptapurine ligand

6-Mercaptopurine, 6-MPH, shown in Figure 4.2.1, is a member of the class of nucleobases known as purines. These compounds consist of a six-membered heterocyclic ring fused to a

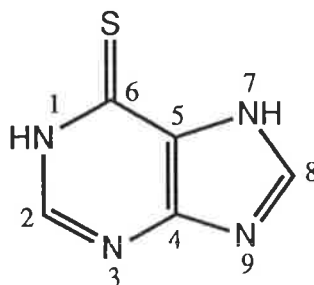


Figure 4.2.1: Labelling Scheme For 6-mercaptapurine.

five membered ring at the 4 and 5 positions, as shown in Figure 4.2.1, thus resulting in a nine membered skeletal structure. Purines are characterized by the 1, 3, 7, 9 substitution pattern of nitrogens for carbons.

The systematic name for 6-MPH is 1,7-dihydro-6H-purine-6-thione, which indicates that the two protonated nitrogens are at positions 1 and 7. This finding is based on crystal structure determinations^{57,58}, but determinations on complexes containing 6-MP show that it is possible for the protons to be bound to the other nitrogens alternatively. Tautomeric studies on the free ligand in solution⁵⁹ have shown that the most dominant tautomer is in fact the 1,9-dihydro-form, suggesting the protons may be sufficiently labile to suit the requirements of a given reaction mechanism: four tautomers of 6MPH are shown in Figure 4.2.2. It is therefore also possible for the 6-thione group to exist as the thiol, although in a proportionately low quantity⁵⁹.

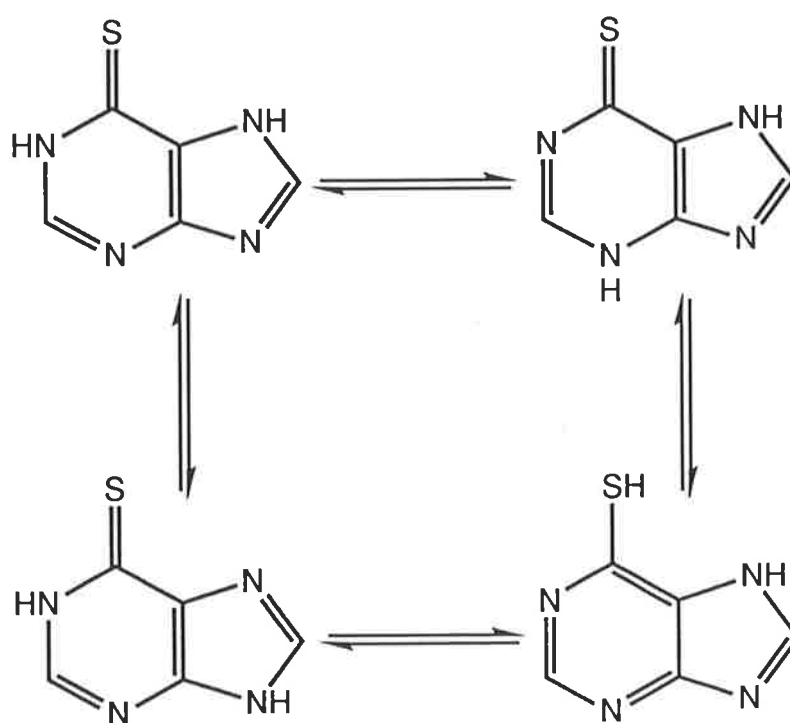


Figure 4.2.2: Tautomeric Structures For 6-mercaptapurine.

For the preparation of the compounds in this thesis, a metathetical reaction involving base was utilized. This involves the exchange of a chloride ion for the 6-MP ligand, though it is not clear whether the reaction is S_N1 or S_N2 . In either mechanism, the 6-MPH molecule must be

deprotonated. Studies on the pK_a values for the protonated nitrogen centres indicate that N¹ is the more acidic atom⁵⁹. Deprotonation here can result in the formal negative charge being located on the sulphur atom through resonance (see Figure 4.2.3), thus leading to coordination to the gold centre. This is verified by the crystal structure determinations presented in Chapter 5.

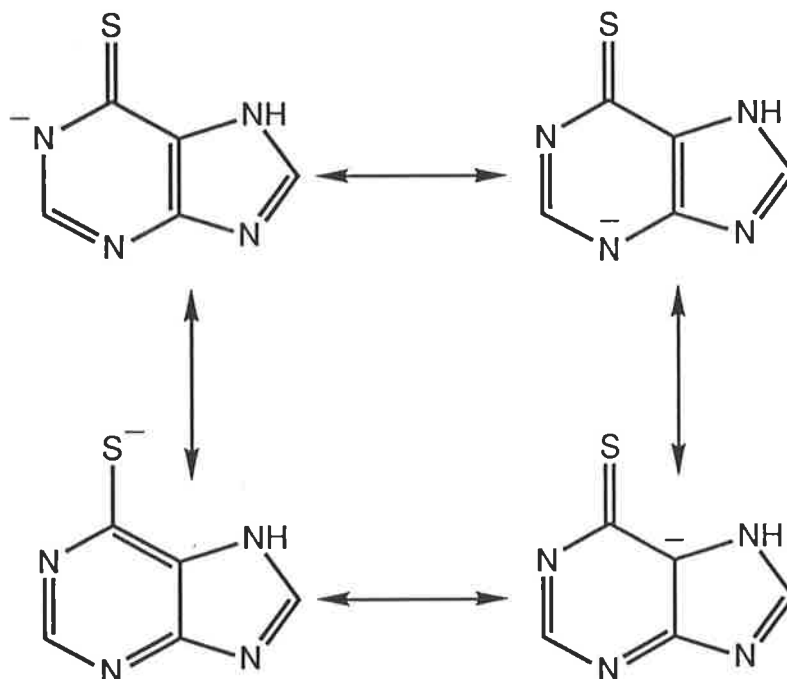


Figure 4.2.3: *Resonance Contributors To The 6-mercaptapurine Anion.*

The electron density in the 6-mercaptapurine ring system is likely to be altered when gold(I) is coordinated by the sulphur atom, leading to greater delocalization of electron density in the six-membered ring system. The changes in electron density should be detectable via spectroscopy.

4.3 Infrared spectroscopy

The heterocyclic, aromatic nature of the 6-mercaptapurine ring system lends itself to a great many possible absorptions in the infrared spectrum. Chromophores such as C=N, C=C and other groups have all been observed to absorb strongly in purines and related thionucleobases^{60,61,62}. Due to the overlapping of many of the absorptions, coupling effects, as noted by Thakur and Singh⁶³, are likely to occur in 6-mercaptapurine and its complexes, leading to complicated spectra. Reports in the literature often conflict with each other as to where certain absorptions occur in the spectra, and on the identities of certain strong

absorptions^{64,65}. The 'thioamide bands'^{63,66}, which result from the coupling of various absorption modes of delocalized N–C–S groups, have not been generally reported for complexes of 6-mercaptapurine, although they are noted for other thionucleobases^{60,61,62}. The thioamide absorption is usually divided into four bands (I, II, III and IV)⁶³, which differ in their frequency ranges due to the constituent absorptions that are coupled. Bands III and IV are usually observed at lower wavenumbers, around the finger-print region, and thus they are not considered here as their assignments would be too tentative. The major contribution to band I is from the $\nu(\text{C}=\text{N})$ vibrational mode, and band II from $\nu(\text{C}=\text{S})$. Hence coordination of 6-mercaptapurine to gold via sulphur should be observable through changes to the absorption frequencies of these two bands.

The lack of definitive data in the literature has resulted in uncertainty as to how the infrared data for the complexes in this thesis should be presented. The scheme thus chosen is to list the data in two tables: 1) Table 4.3.1 presenting the absorptions due to the phosphine moieties (from comparisons with the spectra of the corresponding triorganophosphinegold(I) chloride species); and 2) Table 4.3.2 presenting the absorptions due to the purine group. In the latter table, the absorptions for the thioamide bands I and II, $\nu(\text{N}-\text{C}-\text{S})$, have been resolved from the phosphine peaks by comparison with the spectra of the corresponding gold(I) chloride complexes, but the other absorptions are listed together under the classification of 'purine ring vibrations'. Bands I and II assignments are those peaks occurring in the range where this absorption is observed for analogous heterocycles e.g. for thionucleobase gold(I) complexes^{60,61,62}. These should not be regarded as definitive assignments, but they are likely to be representative of the important thioamide chromophore.

A broad absorption band occurs in many of the spectra at *ca* 3400 cm^{-1} , due to water molecules of hydration, or moisture in the spectrometer environment. This band, when present, overlaps with that of the N–H stretching mode. The $\nu(\text{N}-\text{H})$ absorption in all the spectra occurs over a broad range of approximately 3300 to 2200 cm^{-1} of medium intensity; such a wide range has been observed before in complexes of 6-MP via sulphur coordination^{64,65}. However, the $\nu(\text{C}-\text{H})$ absorption also occurs in this range^{34,35}, so the broad absorption in this region should

Table 4.3.1: Infrared Data For The Phosphine Moieties.

Complex	$\nu(\text{C-H})$	$\nu(\text{C=C})$	$\nu(\text{P-C}), \nu(\text{C-C})$ and $\delta(\text{C-H})$
[Et ₃ PAu(6-MP)]	3099m, 3055m, 2964s, 2931s	-	1456m, 1418m, 1384m, 1268m
[Cycl ₃ PAu(6-MP)]	3099w, 3041w, 2926vs, 2853s	-	1446m, 1417w, 1385s, 1176w
[PhMe ₂ PAu(6-MP)]	3053s, 2952s, 2925s	1465w	1436m, 1419m, 1384m, 1110w
[Ph ₃ PAu(6-MP)]	3054m, 2967m, 2929m	1479m	1434vs, 1421m, 1381s, 1181w, 1100s
[(<i>o</i> -Tol) ₃ PAu(6-MP)]	3054m, 2964m, 2927m	1590s, 1470m	1449s, 1410m, 1384s
[(<i>m</i> -Tol) ₃ PAu(6-MP)]	3035m, 2921m	1477m	1447m, 1418m, 1384m, 1108m
[(<i>p</i> -Tol) ₃ PAu(6-MP)]	3097s, 3035s, 2917s	1597s, 1497s	1441m, 1418s, 1397s, 1384s, 1187s, 1102vs
[dppe(AuCl)(Au(6-MP))]	3055m, 2924s	1575m, 1481m	1435vs, 1384s,sh, 1173s, 1104s
[dppp(AuCl)(Au(6-MP))]	3056m, 2924s, 2851m	1481m	1434vs, 1404m, 1384s, 1104s
[dppm(Au(6-MP)) ₂]	3050s, 2927s, 2923s, 2853s	1483m, 1467m	1436vs, 1421s, 1384vs, 1191m, 1102s
[dppe(Au(6-MP)) ₂]	3050m, 2972m	1481w, 1468w	1436s, 1410m, 1384vs, 1171w, 1104m
[dppp(Au(6-MP)) ₂]	3050s, 2924s, 2852m	1482m	1436vs, 1420s, 1384s, 1186m, 1104s

Note: Units are wavenumbers (cm⁻¹).

Table 4.3.2: Infrared Data For The 6-mercaptopurinate Moieties.

Complex	$\nu(\text{N-C-S})$		Purine
	band I	band II	vibrational modes.
6-MPH	1615s, 1575s, 1558m	1346s	1224s, 1156m, 1147m, 1123m, 1276m, 1529s, 1472m, 1409vs
[Et ₃ PAu(6-MP)]	1586s, 1557vs	1320m	1238m, 1208m, 1155w, 1132w
[Cycl ₃ PAu(6-MP)]	1585m, 1554s	1320m	1269w, 1214w, 1207w, 1176w, 1132w, 1113w
[PhMe ₂ PAu(6-MP)]	1586m, 1558s	1319s	1269w, 1238w, 1208w, 1132w
[Ph ₃ PAu(6-MP)]	1591s, 1559vs	1321s	1244s, 1236s, 1210m, 1157w, 1133w
[(<i>o</i> -Tol) ₃ PAu(6-MP)]	1617m, 1554vs	1322m	1277w, 1241m, 1229m, 1211m, 1164w
[(<i>m</i> -Tol) ₃ PAu(6-MP)]	1587m, 1555s	1319m	1268w, 1234m, 1207m
[(<i>p</i> -Tol) ₃ PAu(6-MP)]	1585s, 1556vs	1319s	1269m, 1234s, 1208s, 1120w
[dppe(AuCl)(Au(6-MP))]	1616s, 1595s, 1558m	1335m	1311m, 1276m, 1217m, 1159m, 1122m
[dppp(AuCl)(Au(6-MP))]	1594m, 1558m	1318w	1347w, 1274w, 1240w, 1209w
[dppm(Au(6-MP)) ₂]	1586s, 1558vs	1320s	1270m, 1236s, 1207m, 1161w
[dppe(Au(6-MP)) ₂]	1587s, 1558s	1319m	1275w, 1236m, br, 1209m, 1122w
[dppp(Au(6-MP)) ₂]	1585s, 1558vs	1319s	1269m, 1238s, 1208m

Note: Units are wavenumbers (cm⁻¹).

be assigned to the overlapping vibrational modes of both $\nu(\text{N-H})$ and $\nu(\text{C-H})$. This absorption can be seen most clearly for 6-mercaptapurine in Figure 4.3.1 and for $[\text{Ph}_3\text{PAu}(6\text{-MP})]$ in Figure 4.3.2. In the latter spectra the $\nu(\text{C-H})$ absorptions from the phosphine moiety can also be observed as sharper peaks around 3000 to 2900 cm^{-1} , an observation that is suggestive of product formation. Further confirmation arises from inspection of the thioamide peaks in these two spectra; the differences in absorption frequencies for band I at 1575 cm^{-1} for 6-MPH and 1591 cm^{-1} for $[\text{Ph}_3\text{PAu}(6\text{-MP})]$ for example. This suggests an increase in bond order of the C=N chromophore, which is expected since the electron density in the aromatic system increases upon coordination. The C=S bond approaches single bond character when the sulphur atom coordinates to gold, and this is observed via band II; the 1346 cm^{-1} peak moves down to 1321 cm^{-1} in the $[\text{Ph}_3\text{PAu}(6\text{-MP})]$ complex. Similar trends are observed for these peaks in all spectra, suggesting product formation in all cases. The deduced changes in bond lengths will be verified in the crystallographic examples in Chapter 5; N(1)–C(6) undergoes a decrease in length of around 0.040 to 0.045 Å, and C(6)–S(6) increases by approximately 0.039 to 0.052 Å upon coordination.

As expected, the frequencies of the phosphine vibrational modes are observed to be independent of the coordination of the purine molecule to the gold atom. From the foregoing discussion, infrared spectroscopy confirms product formation via the appearance of both phosphine and 6-mercaptapurine absorptions in the spectra, and by the frequency shifts of the thioamide absorption bands, which also confirm the expected increase in electron density in the ring system upon complexation.

4.4 ^1H NMR spectroscopy

The data obtained from the proton NMR spectra of the complexes studied in this thesis are reported in Tables 4.4.1 and 4.4.2. The solvent utilized was d_6 -dimethylsulphoxide as this proved to be the only common solvent into which 6-mercaptapurine and the gold complexes would dissolve. NMR studies on these and other analogous complexes were also performed in other solvents (e.g. $(\text{CD}_3)_2\text{CO}$ and CDCl_3) which showed that there was no obvious solvent

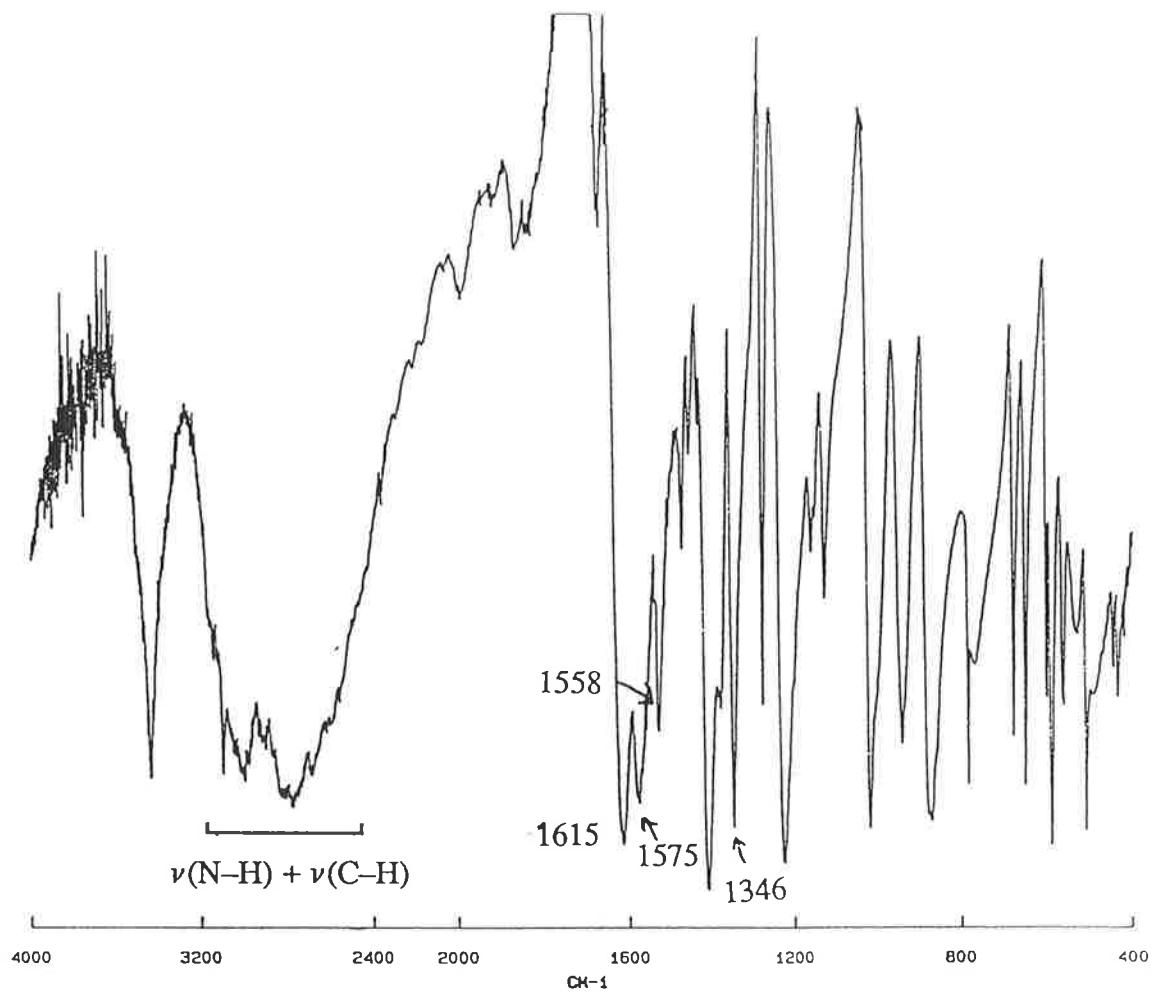


Figure 4.3.1: IR Spectrum Of 6-mercaptapurine, 6-MPH.

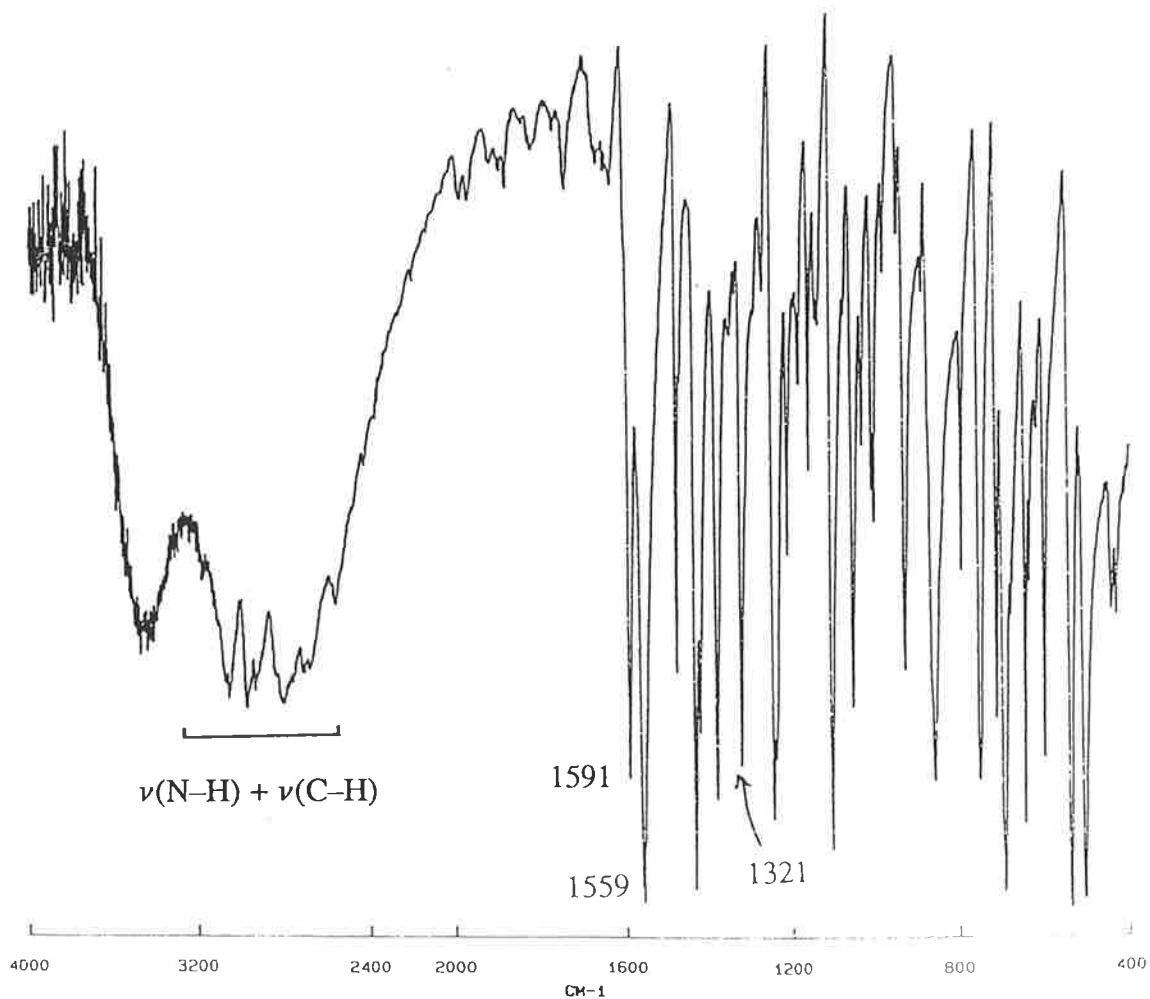


Figure 4.3.2: IR Spectrum Of 6-mercaptapurinato(triphenylphosphine)gold(I), $[\text{Ph}_3\text{PAu}(6\text{-MP})]$.

Table 4.4.1: ^1H NMR Chemical Shift Values (ppm) For The $[\text{R}_3\text{PAu}(6\text{-MP})]$ Complexes.

R_3P	H^2	H^9	H^8	H_α	H_β	H_γ	H_δ	H_{CH_3}
6-MPH	8.45(s)	13.62(br,s)	8.26(s)	-	-	-	-	-
Et ₃ P	8.41(s)	13.21(br,s)	8.31(s)	1.95(m)	1.19(dt) (7.68) ^a (18.35) ^b	-	-	-
Cycl ₃ P	8.34(s)	13.12(br,s)	8.19(s)	2.17(m)	1.89(m)	1.57(m)	1.32(m)	-
Ph ₃ P	8.51(s)	13.35(br,s)	8.31(s)	-	7.73 - 7.64(br,m)			-
(<i>o</i> -Tol) ₃ P	8.35(s)	13.14(br,s)	7.87(s)	-	7.61 - 7.02(br,m)			2.64(s)
(<i>m</i> -Tol) ₃ P	8.45(s)	13.32(br,s)	8.25(s)	-	7.53 - 7.38(br,m)			2.28(s)
(<i>p</i> -Tol) ₃ P	8.46(s)	13.32(br,s)	8.32(s)	-	7.58 - 7.40(br,m)			2.37(s)
PhMe ₂ P	8.45(s)	13.20(br,s)	8.33(s)	-	8.02 - 7.58(br,m)			1.95(d) (10.63) ^c

Note: Coupling constants, in parentheses, are in units of Hertz: a: $^3\text{J}_{\text{H-H}}$, b: $^3\text{J}_{\text{P-H}}$ and c: $^2\text{J}_{\text{P-H}}$.

Table 4.4.2: ^1H NMR Chemical Shift Values (ppm) For The $[(\text{Ph}_2\text{P}(\text{CH}_2)_n\text{PPh}_2)(\text{AuCl})(\text{Au}(6\text{-MP}))]$ And $[(\text{Ph}_2\text{P}(\text{CH}_2)_n\text{PPh}_2)(\text{Au}(6\text{-MP}))_2]$ Complexes.

Complex	H^2	H^9	H^8	Phenyl Protons	H_a	H_b
6-MPH	8.45(s)	13.62(br,s)	8.26(s)	-	-	-
[dppe(AuCl)(Au(6-MP))]	8.36(s)	13.12(br,s)	8.17(s)	7.95 - 7.37(br,m)	3.04(m)	-
[dppp(AuCl)(Au(6-MP))]	8.36(s)	13.25(br,s)	8.09(s)	7.87 - 7.47(br,m)	3.02(m)	-
[dpmm(Au(6-MP)) ₂]	8.37(s)	13.13(br,s)	8.24(s)	7.89 - 7.49(br,m)	4.64(m)	-
[dppe(Au(6-MP)) ₂]	8.40(s)	13.22(br,s)	8.20(s)	7.74 - 7.50(br,m)	3.07(m)	1.73(m)
[dppp(Au(6-MP)) ₂]	8.38(s)	13.18(br,s)	8.19(s)	7.83 - 7.42(br,m)	3.16(m)	1.89(m)

Note: Coupling constants, in parentheses, are in units of Hertz: a: $^3\text{J}_{\text{H-H}}$, b: $^3\text{J}_{\text{P-H}}$ and c: $^2\text{J}_{\text{P-H}}$.

effects⁶². The labelling scheme adopted is the same as that described in Chapter 3 for the phosphine moieties, and the protons of the 6-MP moiety are labelled as in Figure ^{4.2.1}~~4.2.2~~. Crystal structure determinations on $[\text{Ph}_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$ and $[(o\text{-Tol})_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$ (see Chapter 5) revealed that the protonated imadazole nitrogen in these complexes is H⁹, whereas it was H⁷ in 6-mercaptopurine. Therefore the H⁹ resonance given for 6-mercaptopurine in Tables 4.4.1 and 4.4.2 corresponds to H⁷.

As for the triorganophosphinegold(I) chloride analogues, the phenyl regions are not very informative due to the complicated coupling patterns which lead to broad multiplets. This applies also to the spectrum of $[\text{Cycl}_3\text{PAu}(6\text{-MP})]$, where broad and complex multiplets are still observed. If there is an effect on coupling and chemical shifts in the phosphine ligand when 6-MP substitutes for the chloride on the gold centre, then it would best be observed in the spectrum of $[\text{Et}_3\text{PAu}(6\text{-MP})]$, where phosphorus-proton coupling is most clearly illustrated. Table 4.4.3 shows the coupling constants involved with $[\text{Et}_3\text{PAuCl}]$ and $[\text{Et}_3\text{PAu}(6\text{-MP})]$.

Table 4.4.3: *Coupling Constants (Hz) Associated With The Et₃P Moiety.*

Complex	³ J _{PH}	³ J _{HH}	δH _α	δH _β
$[\text{Et}_3\text{PAuCl}]$	18.89	7.71	1.94(m)	1.10(dt)
$[\text{Et}_3\text{PAu}(6\text{-MP})]$	18.35	7.68	1.95(m)	1.19(dt)

It was not possible to determine clearly the values for ²J_{PH} (which are typically in the order of 10-11 Hz)^{60,62}, but the data from the table suggest there is little effect on the coupling and shielding factors of the phosphine moiety between chloride and 6-MP substituted complexes. Hence, the phosphine chemical shifts between the corresponding chloride and 6-MP substituted complexes are fairly similar.

The proton resonances arising from the 6-mercaptopurine moiety give rise to simple peaks in the spectra. The free ligand contains two N-H protons, nominally defined as H¹ and H⁷, which occur together as a very broad resonance centred at δ 13.62 ppm. This phenomenon has been

reported in the literature before, at δ 13.5 ppm in d_6 -dms⁶⁷. On complexation H¹ is removed, but the single N–H resonance for H⁹ is still broad, and is indeed almost unresolved in some of the spectra. Where observed, it resonates in the approximate range of δ 13.10 to δ 13.35 ppm. An example in the literature where 6-MP is complexed via the sulphur atom to a tungsten atom, [W(CO)₅(6-MP)]⁶⁸, gives this proton resonance at δ 14.37 ppm. The high value of the chemical shift is indicative of both the aromatic environment and of the electronegative nitrogen, which both give rise to a deshielding effect. The H² and H⁸ resonances in the free ligand occur as singlets at δ 8.45 and δ 8.26 ppm respectively (the literature example gives δ 8.35 and δ 8.15 ppm)⁶⁷, with the H⁸ resonance being slightly broader and smaller. This is illustrated in Figure 4.4.1 for [(*m*-Tol)₃PAu(6-MP)], where the broad appearance of H⁹ and the complex phenyl region can also be observed. The small difference in shape allowed identification of these peaks in the complexes, where the H² resonance was still observed to be further downfield (the tungsten complex has values of δ 9.23 and δ 8.60 ppm, respectively). The values for both these protons vary between the complexes, and so is no real guide as to whether complexation has occurred; however, the appearance of both phosphine and purine resonances in each spectra and the resultant integration allowed the confirmation of product formation.

4.5 ¹³C {¹H} NMR spectroscopy

The labelling system used is the same as that for the triorganophosphinegold(I) chloride precursors in Chapter 3 and that for the purine in Figure ^{4.2.1}~~4.2.2~~. Tables 4.5.1 and 4.5.2 show the assigned resonances and coupling constants.

The resonances due to the phosphine moiety are consistent with those obtained for the triorganophosphinegold(I) chloride precursors in terms of chemical shift values. In the complexes, the alpha carbon resonates approximately 1 ppm further downfield when R = Ph, *o*-Tol, *m*-Tol or *p*-Tol, and around 0.4 ppm for the other complexes when compared with the triorganophosphinegold(I) chloride precursors. The beta, gamma and delta carbon resonances show no significant changes. However, it is still notable that C₈ for the dpmm, dppe and dppp complexes experiences a general upfield shift of 0.5 to 0.7 ppm. The phenyl resonances in

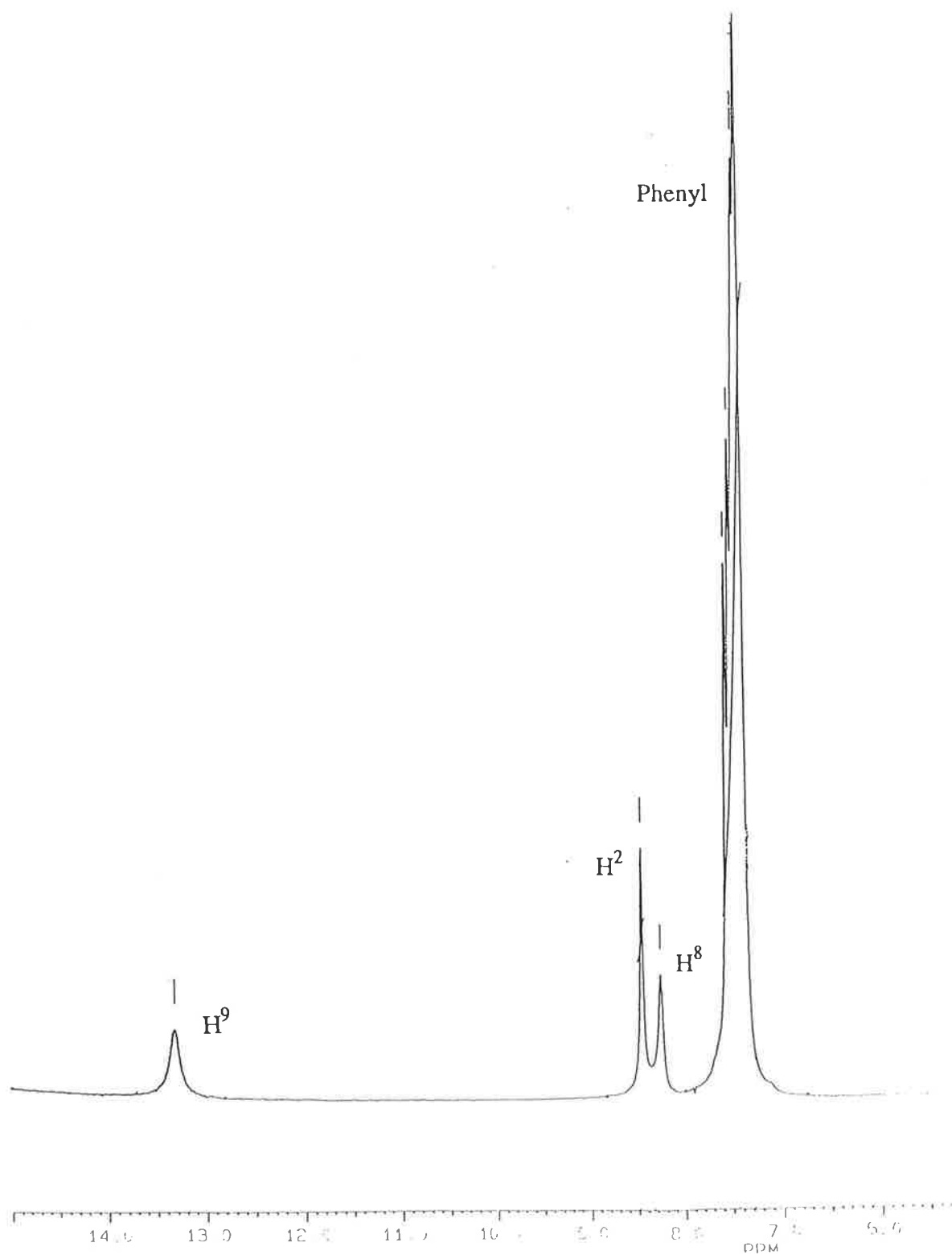


Figure 4.4.1: ^1H NMR Spectrum Of 6-mercaptopurinato[tri(m-tolyl)phosphine]gold(I), $[(m\text{-Tol})_3\text{PAu}(6\text{-MP})]$.

Table 4.5.1: ^{13}C $\{^1\text{H}\}$ And $^{31}\text{P}\{^1\text{H}\}$ Chemical Shifts (ppm) For The Complexes Of The General Formula $[\text{R}_3\text{PAu}(6\text{-MP})]$.

Complex	C ²	C ⁴	C ⁵	C ⁶	C ⁸	C _α	C _β	C _γ	C _δ	C _ε	C _ζ	-CH ₃	³¹ P
6-mercaptopurine	144.6	151.0	127.9	170.8	144.6	-	-	-	-	-	-	-	-
[Et ₃ PAu(6-MP)]	141.3	151.1	132.7	171.0	149.1	17.3(d) (33.81) ^a	8.92	-	-	-	-	-	37.52
[Cycl ₃ PAu(6-MP)]	140.8	150.8	132.7	170.8	149.1	32.6(d) (28.53) ^a	26.4(d) (12.08) ^b	30.1	25.5	30.1	26.4(d) (12.08) ^b	-	57.23
[Ph ₃ PAu(6-MP)]	141.6	151.4	132.4	170.1	149.2	129.4(d) (56.38) ^a	133.9(d) (14.19) ^b	129.6(d) (11.24) ^c	132.0	129.6(d) (11.24) ^c	133.9(d) (14.19) ^b	-	37.53
[(<i>o</i> -Tol) ₃ PAu(6-MP)]	140.6	150.6	Obs.	169.5	148.7	125.4(d) (60.99) ^a	142.3(d) (12.53) ^b	132.3(d) (8.30) ^c	131.9	126.9(d) (9.66) ^c	133.1(d) (9.43) ^b	22.3(d) (10.94) ^c	19.15
[(<i>m</i> -Tol) ₃ PAu(6-MP)]	141.3	151.0	132.4	170.2	149.0	129.2(d) (57.06) ^a	138.8(d) (11.85) ^b	130.9(d) (12.91) ^c	132.5	129.2(d) (11.85) ^c	134.1(d) (14.94) ^b	21.0	37.36
[(<i>p</i> -Tol) ₃ PAu(6-MP)]	141.5	150.9	132.4	170.4	149.0	126.3(d) (59.70) ^a	133.6(d) (14.19) ^b	130.0(d) (11.93) ^c	141.9	130.0(d) (11.93) ^c	133.6(d) (14.19) ^b	20.9	35.61
[PhMe ₂ PAu(6-MP)]	141.4	150.9	Obs.	170.4	149.2	132.9(d) (56.15) ^a	131.9(d) (13.21) ^b	128.9(d) (11.02) ^c	131.3(d) (1.43) ^d	128.9(d) (11.02) ^c	131.9(d) (13.21) ^b	14.8(d) (36.53) ^a	9.21

Note: ^{31}P - ^{13}C coupling constants, in parentheses, are in Hertz, where: a = $^1\text{J}_{\text{PC}}$; b = $^2\text{J}_{\text{PC}}$; c = $^3\text{J}_{\text{PC}}$; d = $^4\text{J}_{\text{PC}}$.

Table 4.5.2: ^{13}C $\{^1\text{H}\}$ And ^{31}P $\{^1\text{H}\}$ Chemical Shifts (ppm) For The Complexes Of The General Formulae
 $[(\text{Ph}_2\text{P}(\text{CH}_2)_n\text{PPh}_2)(\text{AuCl})(\text{Au}(6\text{-MP}))]$ And $[(\text{Ph}_2\text{P}(\text{CH}_2)_n\text{PPh}_2)(\text{Au}(6\text{-MP}))_2]$.

Complex	C ²	C ⁴	C ⁵	C ⁶	C ⁸	C _α	C _β	C _γ	C _δ	C _a	C _b	³¹ P
6-mercaptopurine	144.6	151.0	127.9	170.8	144.6	-	-	-	-	-	-	-
[dppe(AuCl)(Au(6-MP))]	141.5	150.6	Obs.	170.1	149.0	129.2(d) (56.00) ^a	133.2(t) (6.87)	129.2(t) (5.66)	131.9	22.7(t) (19.93) ^a	-	31.6
[dppp(AuCl)Au(6-MP)]	142.0	150.9	Obs.	171.3	150.9	129.1(d) (56.83) ^a	133.0(d) (13.13) ^b	129.3(d) (11.10) ^c	131.9	26.5(dd) (36.08) ^a (12.76) ^c	19.9(m)	28.5
[dppm(Au(6-MP)) ₂]	141.0	150.7	132.5	170.5	149.1	Obs.	133.5(t) (6.72)	128.9(t) (5.38)	131.6	24.9(m)	-	31.3
[dppe(Au(6-MP)) ₂]	142.5	150.3	133.0	169.0	149.0	129.3(d) (56.08) ^a	133.2(t) (6.87)	129.2(t) (5.59)	131.9	22.6(m)	-	33.6
[dppp(Au(6-MP)) ₂]	141.3	150.9	132.7	170.4	149.8	129.0(d) (55.85) ^a	133.1(d) (13.28) ^b	129.1(d) (11.25) ^c	131.6	26.6(d) (30.52) ^a	19.5(m)	31.4

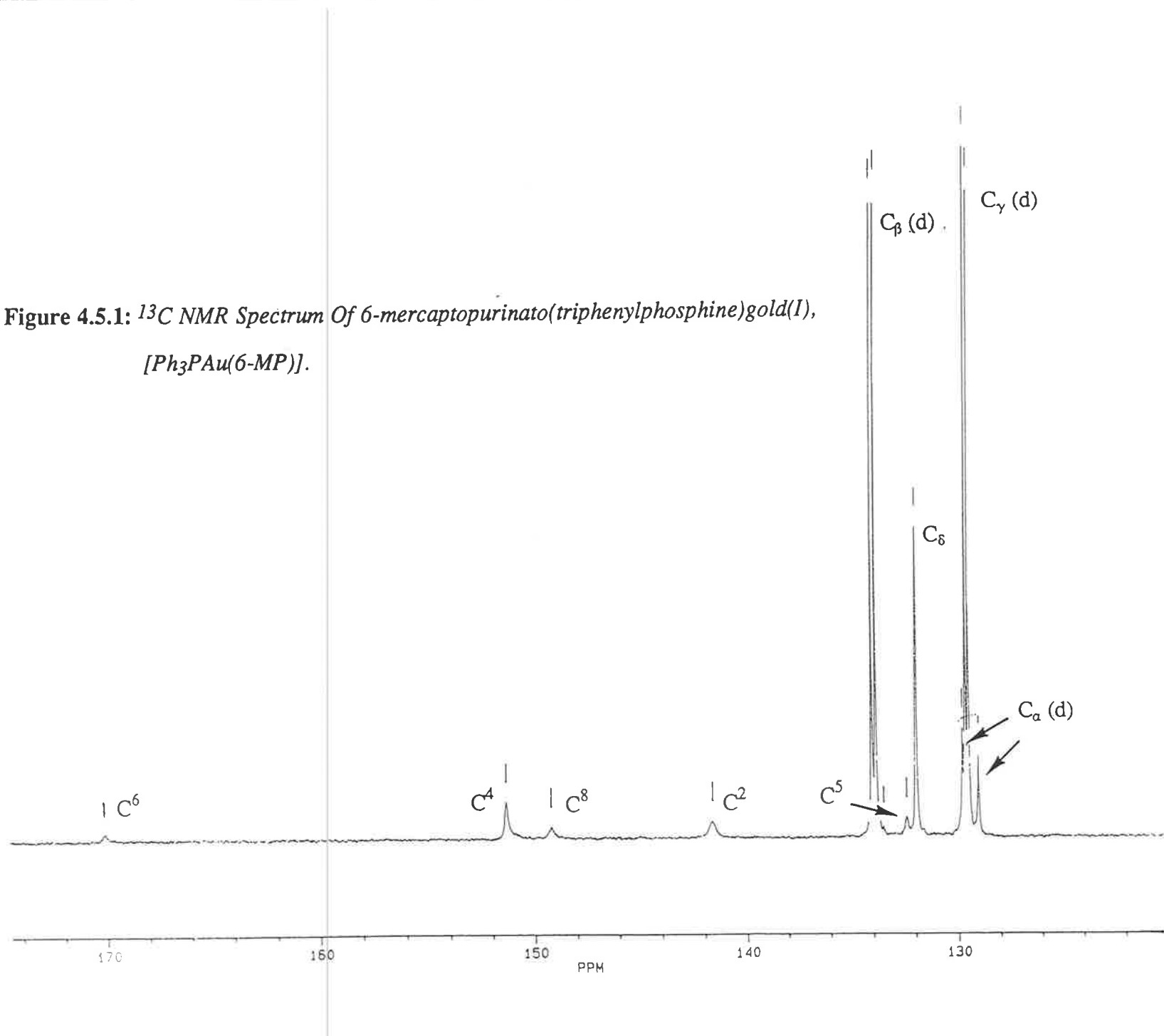
Note: ^{31}P - ^{13}C coupling constants, in parentheses, are in Hertz, where: a = $^1\text{J}_{\text{PC}}$; b = $^2\text{J}_{\text{PC}}$; c = $^3\text{J}_{\text{PC}}$. Obs. indicates obscured.

these complexes again display the coupling pattern due to isotopic asymmetry observed for the triorganophosphinegold(I) chloride precursors. It is also notable that the complex multiplet that might have been expected in the $[(\text{Ph}_2\text{P}(\text{CH}_2)_n\text{PPh}_2)(\text{AuCl})(\text{Au}(6\text{-MP}))]$ complexes due to overlapping phenyl regions is not observed, indicating that all the phenyl rings give equivalent resonance values in solution, hence suggesting that the gold ligands are fluxional; this is supported by evidence from the phosphorus-31 nmr studies (see later). The resonances for C_a and C_b are usually observed as complex multiplets which, like the methyl groups on the tolyl phosphine complexes, do not experience significant changes in chemical shift or coupling constants after complexation.

The five carbon resonances for the 6-MP moiety occur as broad peaks at greater than δ 130 ppm due to the deshielding environment of the aromatic system. The relevant spectral region for $[\text{Ph}_3\text{PAu}(6\text{-MP})]$ is shown in Figure 4.5.1, which demonstrates how the purine resonances have a small and broad appearance. This feature is due to the Nuclear-Overhauser Effect⁶⁹, and arises because the rigid structure of the ring system does not allow efficient dissipation of the spin energies via thermal motion, causing these nuclei to have long relaxation times and hence to occur as broad resonances. The resonances for 6-mercaptapurine are given in Tables 4.4.1 and 4.4.2, and agree with the literature reports that indicate that C^2 and C^8 occur at the same chemical shift value, i.e. δ 144.8 ppm⁷⁰. These peaks are resolved in the complexes as the resonance for C^8 is shifted downfield and that for C^2 is shifted upfield. The chemical shifts of the C^4 and C^6 atoms do not change significantly upon complexation, whereas C^5 is shifted downfield in value by approximately 5 ppm. What follows is a rationalization of the shifts in resonance values observed for each nuclei in going from the free purine to the complexes.

Complexation of 6-mercaptapurine to the gold atom via the sulphur atom causes significant delocalization within the six membered ring system, which results in C^2 experiencing a greater share of electron density, the effect being a slight shielding of the nucleus and so a shift upfield in the spectra. The resonance for C^6 might be expected to change for the same reasons, however, the reduction of double bond character of the $\text{C}=\text{S}$ bond has an opposing effect on

Figure 4.5.1: ^{13}C NMR Spectrum Of 6-mercaptopurinato(triphenylphosphine)gold(I),
[$\text{Ph}_3\text{PAu}(6\text{-MP})$].



electron density, and so the chemical shift value undergoes no significant change. The resonances for C⁴ and C⁵ will experience a similar even distribution of electron density, however, although the C⁵ resonance is indeed observed to move upfield, the signal for C⁴ remains fairly static. The delocalization of electron density in the six-membered ring creates an electron withdrawing effect on the atoms in the imadazole ring, thus resulting in a downfield shift for the C⁸ resonance.

In fact, the shifts in resonance observed are really only minor, but are large enough to suggest that complexation to gold via sulphur has occurred for all the complexes; the resonances for C⁵ and C⁸ are the most indicative of product formation. The occurrence of both phosphine and purine resonances in the spectra is further verification for product formation.

4.6 ³¹P {¹H} NMR spectroscopy

The chemical shifts observed in the ³¹P proton decoupled spectra of the complexes are listed in Tables 4.5.1 and 4.5.2. All the resonances are singlets, as coupling to carbon-13 nuclei is not observed owing to the low natural abundance of this isotope. The spectra are also diagnostic in terms of purity of the complexes.

The resonances can often appear broad: Figures 4.6.1 and 4.6.2 show the spectra of [Ph₃PAu(6-MP)] and [(*m*-Tol)₃PAu(6-MP)], the latter resonance having the broader appearance. The chemical shift values vary appreciably between the complexes, as the phosphorus nucleus is sensitive to changes outside the coordination sphere⁷¹, such as the cone-angle of the phosphine and the electronic environment of the adjacent atoms.

Of note are the spectra of the complexes [dppe(AuCl)(Au(6-MP))] and [dppp(AuCl)(Au(6-MP))], where two types of phosphorus nuclei occur in the molecule but only a single resonance is observed at ambient temperature. Low temperature NMR studies on a representative molecule i.e. [dppe(AuCl)(Au(6-MP))] resolved this resonance into two peaks that differed by approximately 4 ppm at 190 K, suggesting that 6-mercaptopurinate and chloride

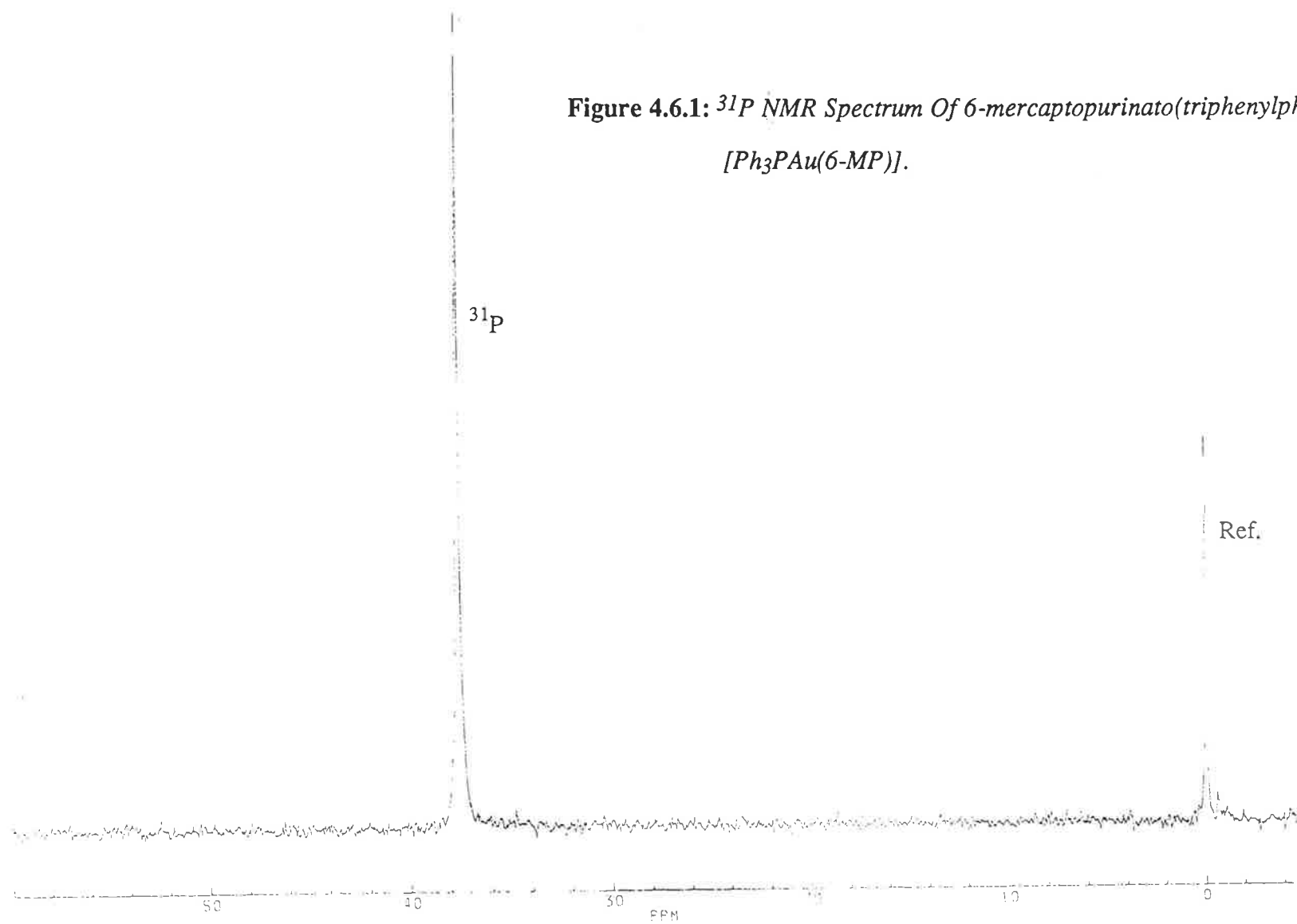
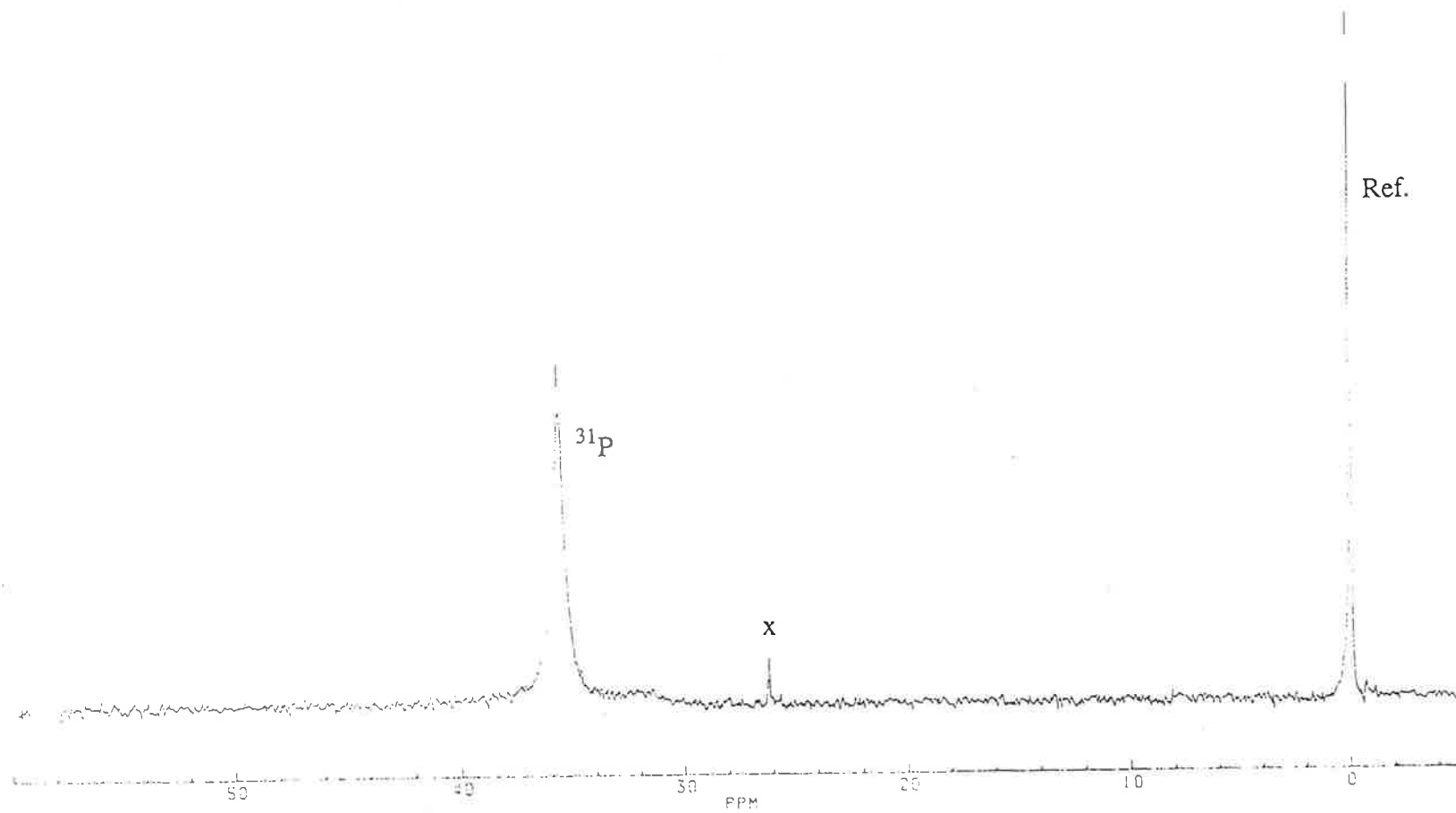


Figure 4.6.1: ^{31}P NMR Spectrum Of 6-mercaptopurinato(triphenylphosphine)gold(I), $[\text{Ph}_3\text{PAu}(6\text{-MP})]$.

Figure 4.6.2: ^{31}P NMR Spectrum Of 6-mercaptapurinato(tri(*p*-tolyl)phosphine)gold(I),
[(*p*-Tol) $_3$ PAu(6-MP)].

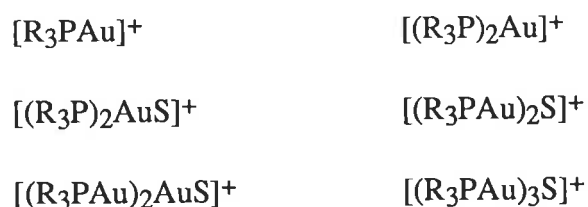


ligands are fluxional at ambient temperatures. The results from the carbon-13 NMR studies on the phenyl regions of these complexes supports this hypothesis.

4.7 Fast Atom Bombardment mass spectroscopy

The technique of Fast Atom Bombardment mass spectroscopy (FAB-MS) involves the bombardment of gaseous molecules by high energy particles, to produce a series of ions that can be observed on a spectrum with respect to their mass/charge ratio and their intensity. The positively charged ions produced are accelerated through a magnetic field in a gaseous state, and so collide with each other and fragment or aggregate according to the stabilities of the resulting ions. The detector in the spectrometer records those ions that reach it as a peak on the spectrum, the most intense of which is assigned an intensity of 100%.

Previous work on monomeric triorganophosphinegold(I) thionucleobases of the general formula $[R_3PAuSR']^{60,62}$ revealed that high nuclearity aggregates involving phosphorus, gold and sulphur can occur in relatively high abundance, indicating not only that such ions are stable but also that both sulphur and phosphorus have an affinity for gold(I), as discussed in Chapter 1. Examples of the types of ions observed are given below:



Tables 4.7.1 and 4.7.2 list the m/e values, intensity (as a percentage of the height of the most intense peak) and the assignment of the most significant ions found in the FAB spectra of the complexes. For those complexes of the general formula $[R_3PAu(6-MP)]$ where $R_3P = Et_3P$, $Cycl_3P$, $PhMe_2P$, Ph_3P , $(o-Tol)_3P$, $(m-Tol)_3P$ or $(p-Tol)_3P$, Table 4.7.1 shows that the molecular ion, $[M]^+$, is observed for all the complexes, suggesting product formation. A selection of the type of fragments listed above are observed in these spectra, suggesting that the binding force between the sulphur and gold atoms is independent of the nature of the thionucleobase. However, fragments analogous to $[(R_3PAu)_2(6-MP)]^+$ were not observed for

Table 4.7.1: Mass Spectral Data For The Complexes Of The General Formula $[R_3PAu(6-MP)]$.

R_3P	$[R_3PAu]^+$	$[M]^+$	$[(R_3P)_2Au]^+$	$[(R_3PAu)_2(6-MP)]^+$	$[(R_3PAu)_2S]^+$	$[(R_3PAu)_3S]^+$
Et ₃ P	315, 77%	466, 96%	433, 100%	780, 79%	-	977, 13%
Cycl ₃ P	477, 60%	628, 90%	757, 100%	1104, 90%	-	1463, 20%
PhMe ₂ P	335, 62%	486, 19%	-	-	702, 8%	1037, 15%
Ph ₃ P	459, 59%	610, 100%	721, 27%	1068, 27%	-	-
(<i>o</i> -Tol) ₃ P	501, 24%	652, 33%	805, 100%	1152, 10%	-	-
(<i>m</i> -Tol) ₃ P	501, 100%	652, 90%	805, 51%	1152, 62%	-	-
(<i>p</i> -Tol) ₃ P	501, 40%	652, 51%	805, 100%	1152, 6%	-	-

Table 4.7.2: Mass Spectral Data For The Complexes Of The General Formulae $[(Ph_2P(CH_2)_nPPh_2)(AuCl)(Au(6-MP))]$ and $[(Ph_2P(CH_2)_nPPh_2)(Au(6-MP))_2]$.

Complex	$[dppnAu]^+$	$[dppnAuS]^+$	$[PhP_2(CH_2)_{(n+1)}Au]^+$	$[(dppnAu)_2]^+$	$[dppnAu_2S]^+$
$[dppm(Au(6-MP))_2]$	581, 8%	-	-	-	-
$[dppe(AuCl)(Au(6-MP))]$	-	-	378, 5%	-	824, 25%
$[dppe(Au(6-MP))_2]$	-	-	-	-	824, 8%
$[dppp(AuCl)(Au(6-MP))]$	609, 11%	641, 4%	-	-	838, 8%
$[dppp(Au(6-MP))_2]$	609, 16%	-	392, 87%	1218, 3%	-

Table 4.7.2: (cont.)

Complex	$[dppnAu_2(6-MP)]^+$	$[Ph_3P(CH_2)_nAu_2S(6-MP)]^+$
$[dppm(Au(6-MP))_2]$	929, 100%	-
$[dppe(AuCl)(Au(6-MP))]$	943, 100%	898, 5%
$[dppe(Au(6-MP))_2]$	943, 100%	-
$[dppp(AuCl)(Au(6-MP))]$	957, 100%	912, 15%
$[dppp(Au(6-MP))_2]$	957, 56%	912, 6%

Note: The symbol n represents m, e, or p for the corresponding complex in that row. $n+1$ indicates an extra methylene group. When used as a subscript, $n = 1$ for dppm complexes, $n = 2$ for dppe complexes and so on.

other thionucleobases, suggesting that the stability of the S-C⁶ bond is higher for 6-mercaptopurine under the conditions of the spectrometer.

The results from Table 4.7.1 suggest, from inspection of the composition of the fragments, that phosphorus forms a strong interaction with gold(I) more often in a 1:1 ratio, but that sulphur has the ability to coordinate to as many as three gold atoms.

The spectra for the complexes of the general formula [(Ph₂P(CH₂)_nPPh₂)(AuCl)(Au(6-MP))] and [(Ph₂P(CH₂)_nPPh₂)(Au(6-MP))₂], the results of which are summarized in Table 4.7.2, show a variety of peaks, the results having less in common between the compounds than observed in Table 4.7.1. The molecular ion is never observed with any significant abundance: for the mono-substituted complexes, this might be attributable to the chloride ion being readily lost in the conditions of the spectrometer. Notable is the appearance of the [dppnAu₂(6-MP)]⁺ fragment in high abundance in all the spectra. This fragment is analogous to the [(R₃PAu)₂(6-MP)]⁺ fragment observed in Table 4.7.1, perhaps indicative of a strong interaction between the gold atom and the 6-mercaptopurine moiety.

The combined spectroscopic evidence presented in this chapter is in accordance with the data for analogous compounds in the literature and suggests product formation for the complexes with the general formulae [R₃PAu(6-MP)] (where R₃P = Et₃P, Cycl₃P, PhMe₂P, Ph₃P, (*o*-Tol)₃P, (*m*-Tol)₃P or (*p*-Tol)₃P), [(Ph₂P(CH₂)_nPPh₂)(AuCl)(Au(6-MP))] (where n = 2 or 3) and [(Ph₂P(CH₂)_nPPh₂)(Au(6-MP))₂] (where n = 1, 2 or 3). The results also indicate the presence of the P-Au-S chromophore with 6-mercaptopurine as the sulphur donor. This chromophore is usually observed to be linear in related structures; the next chapter discusses the structure determinations of two of the complexes listed above, and confirms the spectroscopic results.

CHAPTER 5

Crystallographic Investigations of the Complexes $[\text{Ph}_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$ and $[(o\text{-Tol})_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$

5.1 Introduction

In this section the crystal structures of the complexes $[\text{Ph}_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$ and $[(o\text{-Tol})_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$ and will be discussed. The methods concerned with the data collection and refinement procedures for these structures have already been outlined in Chapter 2. The effects of coordination to gold on the 6-mercaptapurinate moiety will be discussed as a separate section.

5.2 Crystal structure of $[\text{Ph}_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$

Crystals of the complex $[\text{Ph}_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$ were grown from the slow evaporation of a concentrated ethanolic solution of $[\text{Ph}_3\text{PAu}(6\text{-MP})]$. The complex crystallizes in the triclinic space group $P\bar{1}$ (C_1^1 , No. 2)⁴⁴; crystal and refinement data are listed in Table 5.2.1 and the derived results are given in Tables 5.2.2 to 5.2.7. Unlike the other structures featured in this thesis, the structure was solved by Patterson methods, using the program DIRDIF92 PATTY⁷². Calculated and observed structure factors are listed in the Appendix. The crystallographic numbering scheme as drawn with the ORTEP⁴⁵ program is shown in Figure 5.2.1.

Figure 5.2.2 shows the unit cell contents with 15 % thermal ellipsoids. The triclinic space group $P\bar{1}$ is centrosymmetric, and this is demonstrated clearly in Figure 5.2.2. There are two

Table 5.2.1: Crystallographic Parameters for the $[Ph_3PAu(6-MP)].C_2H_5OH$ Complex.

Data	$[Ph_3PAu(6-MP)].C_2H_5OH$
Formula	$C_{25}H_{24}AuSPON_4$
Formula weight	656.5
Crystal shape	octahedral
Crystal dimensions (mm)	0.07 x 0.11 x 0.29
Crystal system	triclinic
Space group	$P\bar{1}$ (C_1^1 , No. 2)
a (Å)	11.066(3)
b (Å)	13.552(3)
c (Å)	8.705(2)
α (°)	91.51(2)
β (°)	113.06(2)
γ (°)	89.69(2)
V (Å ³)	1200.8(5)
Z	2
$\rho_{calc.}$ (g cm ⁻³)	1.816
F(000)	640
μ (cm ⁻¹)	63.27
θ limits, cell (°)	7.51 to 12.82
θ limits, data (°)	1.5 to 28.7
hkl range	0 to 14, -18 to 18, -11 to 10
Range of transmission factors	0.964 to 1.016
Scan technique	$\omega:2\theta$
No. of data measured	6297
No. of unique data	5989
R_{amal}	0.033
No. of unique data used	3978
Criterion of observability	$I \geq 3.0\sigma(I)$
No. of parameters	298
R	0.034
R_w	0.029
Residual electron density (e Å ⁻³)	-0.98 to 0.72

Table 5.2.2: Fractional Atomic Coordinates For The $[Ph_3PAu(6-MP)].C_2H_5OH$ Complex.

Atom	x	y	z
Au	0.10466(2)	-0.22959(2)	-0.00005(3)
S(6)	-0.0379(2)	-0.3205(1)	-0.2211(2)
P(1)	0.2556(2)	-0.1366(1)	0.1981(2)
O(41)	0.7564(5)	0.4593(3)	0.4445(5)
N(1)	-0.2015(5)	-0.4645(4)	-0.2436(5)
N(3)	-0.2628(6)	-0.5303(4)	-0.0324(6)
N(7)	-0.0271(5)	-0.3431(3)	0.1643(5)
N(9)	-0.1475(5)	-0.4520(4)	0.2358(5)
C(2)	-0.2659(7)	-0.5257(5)	-0.1846(8)
C(4)	-0.1778(6)	-0.4651(4)	0.0697(7)
C(5)	-0.1043(5)	-0.3985(4)	0.0263(6)
C(6)	-0.1182(5)	-0.3987(4)	-0.1382(6)
C(8)	-0.0567(6)	-0.3789(5)	0.2828(7)
C(11)	0.4109(5)	-0.1995(4)	0.2864(6)
C(12)	0.4960(7)	-0.1823(5)	0.4499(7)
C(13)	0.6143(7)	-0.2308(5)	0.5115(8)
C(14)	0.6469(7)	-0.2948(5)	0.4109(9)
C(15)	0.5622(7)	-0.3118(5)	0.2515(8)
C(16)	0.4439(6)	-0.2646(5)	0.1883(7)
C(21)	0.2909(6)	-0.0202(4)	0.1290(7)
C(22)	0.4141(6)	0.0105(5)	0.1581(8)
C(23)	0.4355(8)	0.1011(6)	0.1034(9)
C(24)	0.3327(9)	0.1588(5)	0.0194(9)
C(25)	0.2100(8)	0.1286(5)	-0.0102(9)
C(26)	0.1878(7)	0.0396(5)	0.0435(9)

Table 5.2.2 (continued)

C(31)	0.2123(6)	-0.1065(5)	0.3719(7)
C(32)	0.2333(8)	-0.0148(5)	0.4489(9)
C(33)	0.2036(9)	0.0031(7)	0.5858(10)
C(34)	0.1534(9)	-0.0708(8)	0.6458(10)
C(35)	0.1368(11)	-0.1604(8)	0.5731(11)
C(36)	0.1639(9)	-0.1781(6)	0.4345(9)
C(41)	0.5805(11)	0.3773(8)	0.2306(12)
C(42)	0.6319(12)	0.4197(7)	0.3923(11)

Table 5.2.3: Anisotropic Thermal Parameters For The $[Ph_3PAu(6-MP)].C_2H_5OH$ Complex.

Atom	U ₁₁	U ₂₂	U ₃₃	U ₁₂	U ₁₃	U ₂₃
Au	0.0484(1)	0.0500(1)	0.0367(1)	-0.0138(1)	0.01005(9)	-0.00523(9)
S(6)	0.059(1)	0.064(1)	0.0324(7)	-0.0208(8)	0.0102(7)	-0.0032(7)
P(1)	0.050(1)	0.048(1)	0.0369(8)	-0.0140(7)	0.0137(7)	-0.0072(7)
O(41)	0.093(4)	0.084(3)	0.045(3)	-0.029(3)	0.025(3)	-0.005(2)
N(1)	0.067(3)	0.061(3)	0.036(3)	-0.023(3)	0.010(2)	-0.007(2)
N(3)	0.086(4)	0.069(4)	0.044(3)	-0.035(3)	0.017(3)	-0.003(3)
N(7)	0.064(3)	0.054(3)	0.032(3)	-0.015(3)	0.010(2)	-0.007(2)
N(9)	0.070(3)	0.058(3)	0.035(3)	-0.010(3)	0.017(2)	-0.001(2)
C(2)	0.085(5)	0.072(5)	0.046(4)	-0.047(4)	0.014(4)	-0.010(3)
C(4)	0.062(4)	0.043(3)	0.037(3)	0.001(3)	0.015(3)	0.002(3)
C(5)	0.047(3)	0.038(3)	0.033(3)	-0.000(3)	0.006(2)	-0.002(2)
C(6)	0.045(3)	0.039(3)	0.038(3)	0.000(3)	0.011(3)	0.003(2)
C(8)	0.070(4)	0.061(4)	0.039(3)	-0.003(3)	0.016(3)	-0.005(3)
C(11)	0.050(3)	0.044(3)	0.037(3)	-0.014(3)	0.010(3)	-0.001(3)
C(12)	0.070(5)	0.056(4)	0.044(4)	-0.007(4)	0.001(3)	-0.013(3)
C(13)	0.071(5)	0.068(4)	0.049(4)	0.002(4)	-0.003(3)	0.005(3)
C(14)	0.064(5)	0.057(4)	0.073(5)	-0.002(4)	0.014(4)	0.006(4)
C(15)	0.070(5)	0.062(4)	0.058(4)	0.001(4)	0.019(4)	-0.005(3)
C(16)	0.068(4)	0.056(4)	0.039(3)	-0.018(3)	0.017(3)	-0.006(3)
C(21)	0.056(4)	0.038(3)	0.039(3)	-0.010(3)	0.018(3)	-0.007(3)
C(22)	0.057(4)	0.059(4)	0.059(4)	-0.011(3)	0.021(3)	0.001(3)
C(23)	0.083(5)	0.066(5)	0.087(6)	-0.024(4)	0.039(5)	0.001(4)
C(24)	0.121(7)	0.048(4)	0.080(5)	-0.011(5)	0.059(5)	-0.003(4)
C(25)	0.094(6)	0.053(5)	0.087(6)	0.016(4)	0.037(5)	0.013(4)
C(26)	0.064(4)	0.061(5)	0.076(5)	-0.004(4)	0.029(4)	0.001(4)

Table 5.2.3 (continued)

C(31)	0.054(4)	0.072(4)	0.039(3)	-0.015(3)	0.016(3)	-0.009(3)
C(32)	0.107(6)	0.069(5)	0.066(5)	-0.019(4)	0.046(5)	-0.019(4)
C(33)	0.127(8)	0.101(7)	0.090(6)	-0.014(6)	0.061(6)	-0.033(5)
C(34)	0.126(8)	0.148(9)	0.066(5)	-0.046(7)	0.059(6)	-0.034(6)
C(35)	0.21(1)	0.17(1)	0.088(7)	-0.12(1)	0.100(8)	-0.052(7)
C(36)	0.165(9)	0.106(6)	0.068(5)	-0.078(6)	0.071(6)	-0.036(5)
C(41)	0.19(1)	0.15(1)	0.092(7)	-0.083(8)	0.037(7)	-0.029(7)
C(42)	0.20(1)	0.120(8)	0.071(6)	-0.078(8)	0.057(7)	-0.035(5)

Table 5.2.4: *Hydrogen Atom Parameters For The [Ph₃PAu(6-MP)].C₂H₅OH Complex.*

Atom	x	y	z	B(eq)
H(2)	-0.3228	-0.5735	-0.2649	6.4
H(8)	-0.0163	-0.3550	0.3975	6.1
H(9)	-0.1830	-0.4871	0.3047	12.4
H(12)	0.4724	-0.1366	0.5211	6.0
H(13)	0.6743	-0.2192	0.6267	6.4
H(14)	0.7310	-0.3282	0.4538	6.4
H(15)	0.5856	-0.3578	0.1806	6.1
H(16)	0.3834	-0.2776	0.0738	5.3
H(22)	0.4880	-0.0314	0.2177	5.3
H(23)	0.5240	0.1231	0.1258	7.0
H(24)	0.3467	0.2223	-0.0202	6.9
H(25)	0.1359	0.1706	-0.0704	6.8
H(26)	0.0986	0.0186	0.0210	6.1
H(32)	0.2695	0.0377	0.4057	6.8
H(33)	0.2179	0.0682	0.6391	7.1
H(34)	0.1302	-0.0589	0.7410	9.8
H(35)	0.1052	-0.2142	0.6194	12.7
H(36)	0.1475	-0.2429	0.3806	9.7
H(41b)	0.5784	0.4268	0.1507	13.2
H(41c)	0.6354	0.3229	0.2238	9.2
H(41a)	0.4923	0.3538	0.2057	13.2
H(42a)	0.6350	0.3691	0.4699	10.8
H(42b)	0.5740	0.4722	0.3975	10.8

Table 5.2.5: Bond Distances (Å) For The [Ph₃PAu(6-MP)].C₂H₅OH Complex.

Atom	Atom	Distance	Atom	Atom	Distance
Au	– S(6)	2.287(1)	C(11)	– C(16)	1.358(8)
Au	– P(1)	2.237(2)	C(12)	– C(13)	1.375(9)
S(6)	– C(6)	1.728(5)	C(13)	– C(14)	1.362(9)
P(1)	– C(11)	1.804(6)	C(41)	– C(42)	1.40(1)
P(1)	– C(21)	1.801(6)	C(14)	– C(15)	1.351(8)
P(1)	– C(31)	1.794(6)	C(15)	– C(16)	1.368(9)
O(41)	– C(42)	1.38(1)	C(21)	– C(22)	1.352(8)
N(1)	– C(2)	1.332(7)	C(21)	– C(26)	1.366(8)
N(1)	– C(6)	1.338(6)	C(22)	– C(23)	1.383(9)
N(3)	– C(2)	1.315(7)	C(23)	– C(24)	1.34(1)
N(3)	– C(4)	1.329(7)	C(24)	– C(25)	1.34(1)
N(7)	– C(5)	1.376(6)	C(25)	– C(26)	1.363(9)
N(7)	– C(8)	1.303(7)	C(31)	– C(32)	1.373(8)
N(9)	– C(4)	1.358(6)	C(31)	– C(36)	1.337(9)
N(9)	– C(8)	1.352(7)	C(32)	– C(33)	1.371(9)
C(4)	– C(5)	1.373(7)	C(33)	– C(34)	1.36(1)
C(5)	– C(6)	1.379(7)	C(34)	– C(35)	1.33(1)
C(11)	– C(12)	1.379(7)	C(35)	– C(36)	1.37(1)

Table 5.2.6: Bond Angles (°) For The $[Ph_3PAu(6-MP)].C_2H_5OH$ Complex.

Atom	Atom	Atom	Angle	Atom	Atom	Atom	Angle
S(6)	– Au	– P(1)	173.71(6)	P(1)	– C(11)	– C(16)	118.7(4)
Au	– S(6)	– C(6)	105.9(2)	C(12)	– C(11)	– C(16)	119.8(6)
Au	– P(1)	– C(11)	111.4(2)	C(11)	– C(12)	– C(13)	119.6(6)
Au	– P(1)	– C(21)	114.9(2)	C(12)	– C(13)	– C(14)	120.0(6)
Au	– P(1)	– C(31)	113.3(2)	C(13)	– C(14)	– C(15)	119.9(7)
C(11)	– P(1)	– C(21)	105.7(3)	C(14)	– C(15)	– C(16)	120.9(6)
C(11)	– P(1)	– C(31)	105.2(3)	C(11)	– C(16)	– C(15)	119.8(6)
C(21)	– P(1)	– C(31)	105.4(3)	P(1)	– C(21)	– C(22)	123.0(5)
C(2)	– N(1)	– C(6)	118.7(5)	P(1)	– C(21)	– C(26)	118.1(5)
C(2)	– N(3)	– C(4)	111.0(5)	C(22)	– C(21)	– C(26)	118.8(6)
C(5)	– N(7)	– C(8)	102.8(5)	C(21)	– C(22)	– C(23)	120.5(6)
C(4)	– N(9)	– C(8)	104.0(5)	C(22)	– C(23)	– C(24)	119.7(7)
N(1)	– C(2)	– N(3)	128.9(5)	C(23)	– C(24)	– C(25)	120.0(7)
N(3)	– C(4)	– N(9)	126.5(6)	C(24)	– C(25)	– C(26)	120.8(7)
N(3)	– C(4)	– C(5)	126.2(5)	C(21)	– C(26)	– C(25)	120.2(7)
N(9)	– C(4)	– C(5)	107.3(5)	P(1)	– C(31)	– C(32)	122.9(5)
N(7)	– C(5)	– C(4)	110.1(5)	P(1)	– C(31)	– C(36)	118.4(5)
N(7)	– C(5)	– C(6)	132.2(5)	C(32)	– C(31)	– C(36)	118.6(6)
C(4)	– C(5)	– C(6)	117.7(5)	C(31)	– C(32)	– C(33)	120.6(7)
S(6)	– C(6)	– N(1)	116.9(4)	C(32)	– C(33)	– C(34)	119.5(8)
S(6)	– C(6)	– C(5)	125.6(4)	C(33)	– C(34)	– C(35)	119.5(8)
N(1)	– C(6)	– C(5)	117.5(5)	C(34)	– C(35)	– C(36)	121.1(8)
N(7)	– C(8)	– N(9)	115.8(5)	C(31)	– C(36)	– C(35)	120.6(8)
P(1)	– C(11)	– C(12)	121.5(5)	O(41)	– C(42)	– C(41)	116.1(9)

Table 5.2.7: Mean Plane Data For The $[Ph_3PAu(6-MP)].C_2H_5OH$ Complex.

Plane number 1: Least-squares plane through the 6-mercaptopurinate moiety.

Atoms Defining Plane	Distance (Å)	esd (Å)
S(6)	-0.0024	0.0018
N(1)	0.0192	0.0054
N(3)	-0.0226	0.0062
N(7)	0.0110	0.0051
N(9)	-0.0094	0.0052
C(2)	0.0202	0.0079
C(4)	-0.0124	0.0059
C(5)	-0.0003	0.0054
C(6)	0.0085	0.0054
C(8)	0.0148	0.0065
Additional Atoms	Distance (Å)	
Au	0.1465	

Mean deviation from plane is 0.0121 Å.

Chi-squared: 56.4.

Table 5.2.7 (continued)

Plane number 2: Least-squares plane through the phenyl ring defined by the atoms C(11) to C(16).

Atoms Defining Plane	Distance (Å)	esd (Å)
C(11)	0.0047	0.0051
C(12)	-0.0023	0.0067
C(13)	-0.0050	0.0074
C(14)	0.0060	0.0067
C(15)	-0.0003	0.0068
C(16)	-0.0048	0.0057
Additional Atom	Distance (Å)	
P(1)	0.0342	

Mean deviation from plane is 0.0039 (Å).

Chi-squared: 2.8.

Plane number 3: Least-squares plane through the phenyl ring defined by the atoms C(21) to C(26).

Atoms Defining Plane	Distance (Å)	esd (Å)
C(21)	0.0014	0.0051
C(22)	-0.0023	0.0062
C(23)	0.0020	0.0073
C(24)	-0.0005	0.0068
C(25)	0.0001	0.0073
C(26)	-0.0010	0.0067
Additional Atom	Distance (Å)	
P(1)	0.0118	

Mean deviation from plane is 0.0012 Å.

Chi-squared: 0.3.

Table 5.2.7 (*continued*)

Plane number 4: Least-squares plane through the phenyl ring defined by the atoms C(31) to C(36).

Atoms Defining Plane	Distance (Å)	esd (Å)
C(31)	-0.0031	0.0061
C(32)	0.0062	0.0081
C(33)	-0.0002	0.0094
C(34)	-0.0111	0.0100
C(35)	0.0177	0.0126
C(36)	-0.0024	0.0101
Additional Atom	Distance (Å)	
P(1)	0.0790	

Mean deviation from plane is 0.0068 Å.

Chi-squared: 3.5.

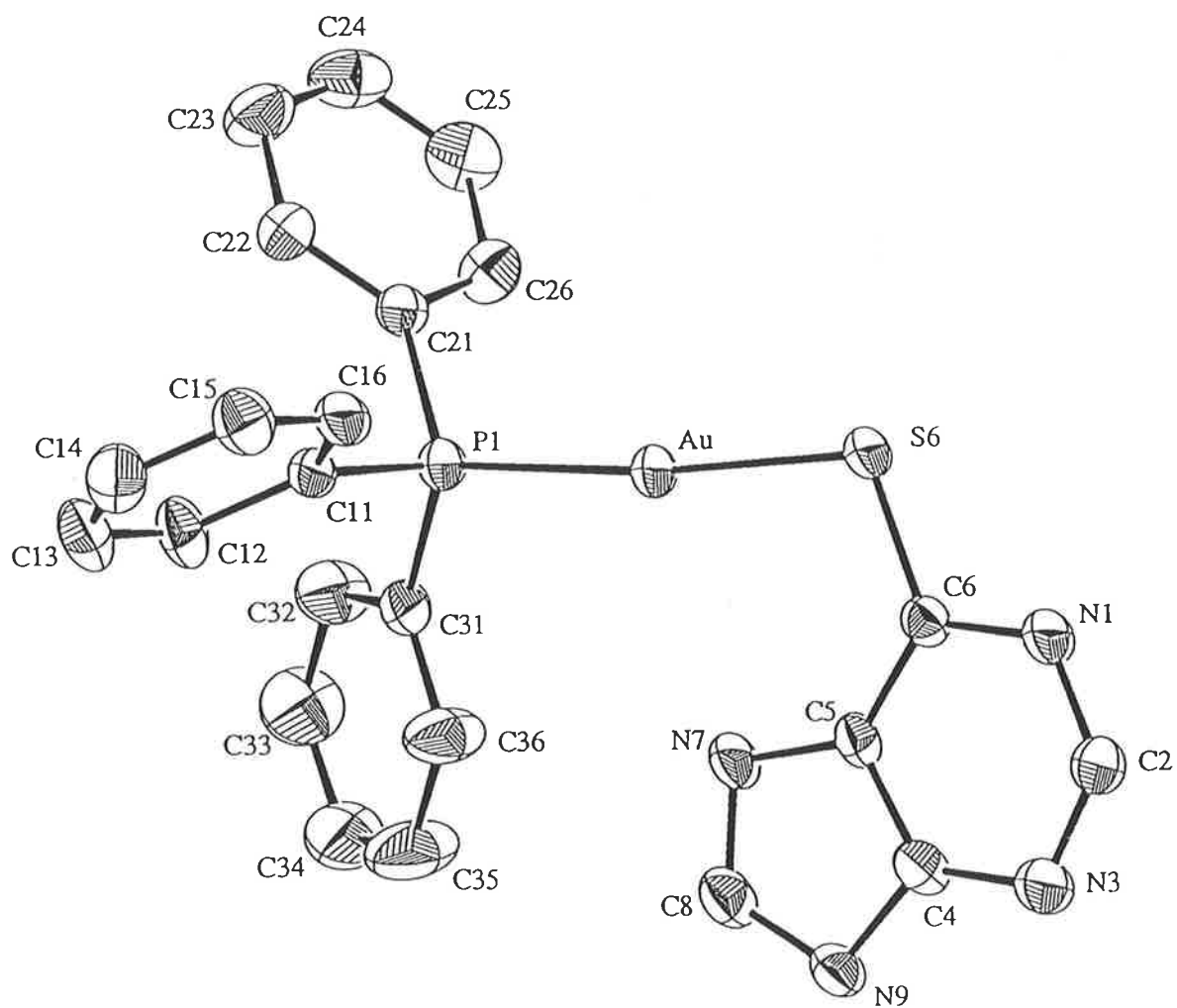


Figure 5.2.1: *Molecular Structure And Crystallographic Numbering Scheme For [Ph₃PAu(6-MP)].*

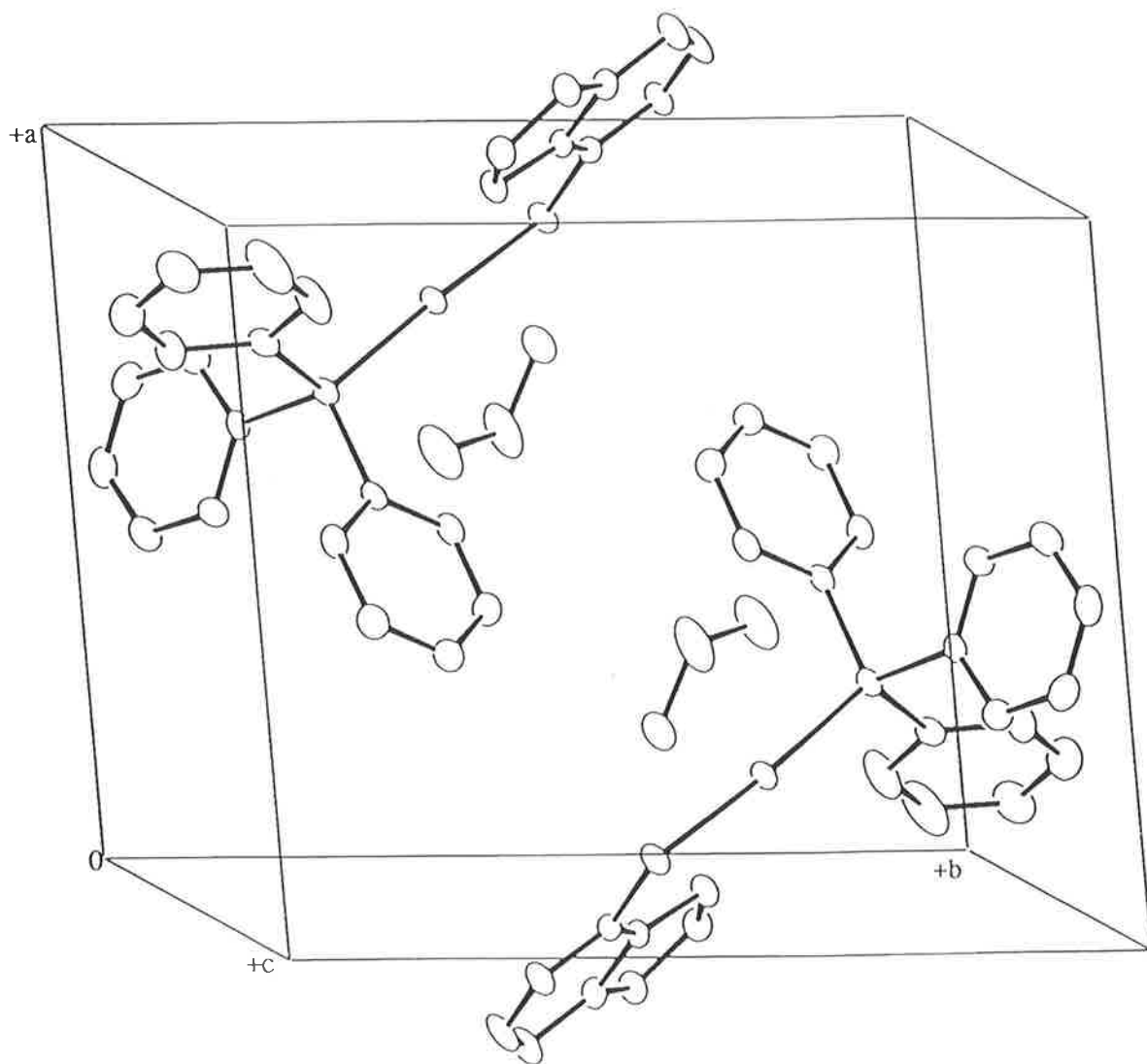


Figure 5.2.2: Unit Cell Diagram Of $[\text{Ph}_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$.

ethanol solvent molecules of crystallization in the unit cell such that the ratio of complex to ethanol molecules is 1:1. Hydrogen bonding is observed to occur between O(41) of the ethanol molecule and H(9) of the 6-mercaptapurinate moiety, at a distance of 1.77 Å; this is demonstrated in the extended lattice diagram of Figure 5.2.3. Hydrogen bonding might be expected to occur between nearby purine moieties; however, the planes of any two purine rings that are nearest neighbours are only parallel, not co-planar. Figure 5.2.3 also indicates how the molecules pack in such a way in the lattice such that the phosphine groups form a layer, and the more polar gold(I), purine and ethanol regions form another layer. There are no significant Au...Au interactions in the lattice: the closest such interaction is 6.628(1) Å, which is larger than range of Au...Au contacts of 2.75 to 3.25 Å normally considered to indicate a 'significant' interaction⁴⁶.

Figure 5.2.1 is an ORTEP⁴⁵ diagram of [Ph₃PAu(6-MP)] plotted with 30% probability ellipsoids. The gold atom exists in the expected linear geometry defined by the P(1) atom of the triphenylphosphine ligand and the S(6) atom derived from the 6-mercaptapurinate moiety, with a P–Au–S angle of 173.71(6)°. The deviation from ideal geometry may be related to the presence of the close intramolecular Au...N(7) contact. The intramolecular distance is 2.884(5) Å, less than the sum of the van der Waals radii of Au and N of 3.25 Å, but not suggestive of a significant interaction. The length of the P(1)–Au bond is found to be 2.237(1) Å, which is equivalent within standard deviation to the value of 2.235(3) Å for [Ph₃PAuCl]⁴⁷. Comparison between the Ph₃P moieties of [Ph₃PAuCl] and [Ph₃PAu(6-MP)] reveals that this region is virtually identical in both complexes; hence, coordination of the 6-MP ligand to gold has a negligible effect on the electronic factors of the Ph₃P ligand.

As expected, the phenyl rings of the phosphine moiety are all planar; the maximum mean deviation is 0.01(1) Å for the ring defined by atoms C(31) to C(36). The maximum deviation of P(1) out of this ring is 0.079 Å. The internal bond distances and bond angles are all typical of electron delocalization in the six-membered aromatic system.

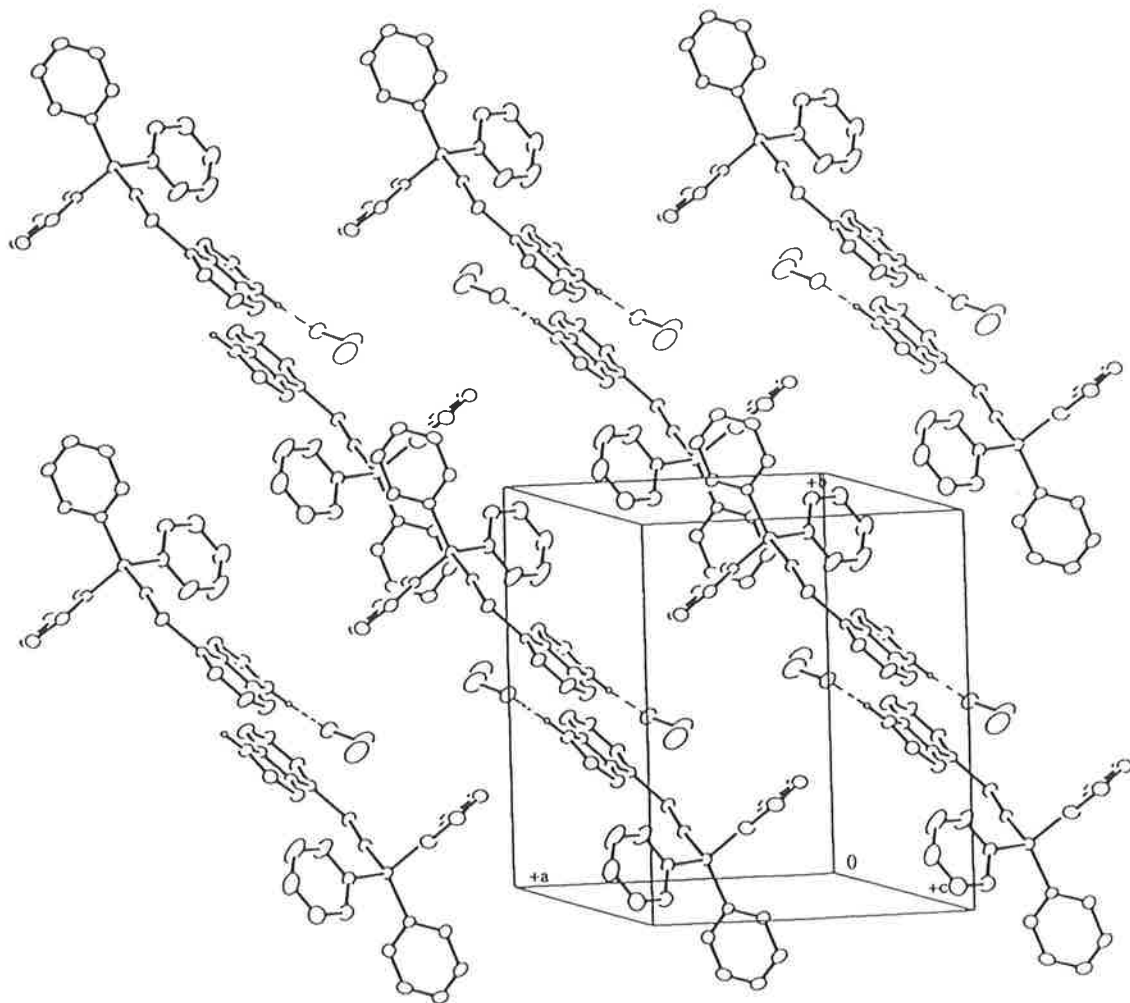


Figure 5.2.3: *Lattice Diagram Of $[\text{Ph}_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$.*

The Au–S bond distance is 2.287(1) Å, which is a typical value for triorganophosphinegold(I) thionucleobase complexes; examples of these will be given in Chapter 6. The intramolecular parameters of the 6-mercaptapurinate moiety will be discussed in detail in section 5.4.

5.3 Crystal Structure of [(*o*-Tol)₃PAu(6-MP)].C₂H₅OH.

Crystals of the complex [(*o*-Tol)₃PAu(6-MP)].C₂H₅OH, grown from the slow evaporation of a concentrated ethanolic solution of the compound, crystallize in the monoclinic space group P2₁/n (C_{2h}⁵, No. 14)⁴⁴. Crystal and refinement data are listed in Table 5.3.1 and the derived parameters are given in Tables 5.3.2 to 5.3.7. Calculated and observed structure factors can be found in the Appendix. The crystallographic numbering scheme ^{is} shown in the ORTEP diagram in Figure 5.3.1.

The contents of the unit cell are illustrated in Figure 5.3.2. The ethanol and 6-mercaptapurinate regions are observed to be associated as a layer, as are the phosphine moieties. Hydrogen bonding occurs between H(9) and N(3) of nearby purine rings, at a intermolecular distance of 1.93 Å. This association is possible as the purine ring systems are close to co-planar; Figure 5.3.3 shows this interaction in detail. The closest Au...Au contact in the lattice is at a distance of 7.821(2) Å; as for [Ph₃PAu(6-MP)].C₂H₅OH this distance is not indicative of a significant interaction.

Figure 5.3.1 shows an ORTEP diagram of the complex with 30% thermal ellipsoids. Notable is the similarity in appearance with [Ph₃PAu(6-MP)]. As expected, the P–Au–S chromophore is nearly linear, with an angle of 177.03(8)°. This is closer to linearity than for [Ph₃PAu(6-MP)]; this may be due to the different hydrogen bonding interactions, or is perhaps due to the larger cone angle of (*o*-Tol)₃P and the possible consequent steric interactions. The N(7) atom is directed towards the gold centre at a distance of 2.860(7) Å, again less than the sum of the van der Waals radii, but not suggestive of any significant bonding interaction⁴⁶.

Table 5.3.1: Crystallographic Parameters for the [(*o*-Tol)₃PAu(6-MP)].C₂H₅OH Complex.

Data	[(<i>o</i> -Tol) ₃ PAu(6-MP)].C ₂ H ₅ OH
Formula	C ₂₈ H ₃₀ AuSPON ₄
Formula weight	698.6
Crystal shape	hexagonal
Crystal dimensions (mm)	0.07 x 0.11 x 0.35
Crystal system	monoclinic
Space group	P2 ₁ /n (C _{2h} ⁵ , No. 14)
<i>a</i> (Å)	10.067(2)
<i>b</i> (Å)	10.518(2)
<i>c</i> (Å)	25.416(4)
β (°)	98.42(2)
<i>V</i> (Å ³)	2662.1(9)
Z	4
$\rho_{\text{calc.}}$ (g cm ⁻³)	1.778 1.743
F(000)	1404 1376
μ (cm ⁻¹)	57.16
θ limits, cell (°)	7.7 to 12.8
θ limits, data (°)	1.5 to 27.9
<i>hkl</i> range	0 to 12, 0 to 13, -33 to 33
Range of transmission factors	0.939 to 1.070
Scan technique	ω :2 θ
No. of data measured	6280
No. of unique data	5913
R _{amal}	0.025
No. of unique data used	4183
Criterion of observability	$I \geq 3.0\sigma(I)$
No. of parameters	325
R	0.040
R _w	0.041
Residual electron density (e Å ⁻³)	3.51

Table 5.3.2: Fractional Atomic Coordinates For The [(*o*-Tol)₃PAu(6-MP)].C₂H₅OH Complex.

Atom	x	y	z
Au	0.34432(3)	0.51852(3)	0.14775(1)
S(6)	0.1721(2)	0.6422(2)	0.1099(1)
P(1)	0.5073(2)	0.3918(2)	0.1878(1)
O(41)	0.1220(7)	0.2139(6)	-0.0267(3)
N(1)	0.1579(7)	0.8334(6)	0.0445(3)
N(3)	0.3315(8)	0.9469(7)	0.0122(3)
N(7)	0.4869(7)	0.6950(6)	0.0912(3)
N(9)	0.5481(7)	0.8552(7)	0.0425(3)
C(2)	0.2072(10)	0.9250(9)	0.0176(4)
C(4)	0.4130(9)	0.8625(8)	0.0393(4)
C(5)	0.3778(8)	0.7632(7)	0.0693(3)
C(6)	0.2430(8)	0.7502(7)	0.0717(3)
C(8)	0.5863(9)	0.7533(9)	0.0744(4)
C(11)	0.5543(7)	0.4320(7)	0.2573(3)
C(12)	0.5587(8)	0.3363(8)	0.2951(4)
C(13)	0.5845(10)	0.3647(9)	0.3476(4)
C(14)	0.6032(10)	0.4863(11)	0.3635(4)
C(15)	0.6004(9)	0.5808(8)	0.3272(4)
C(16)	0.5769(8)	0.5574(7)	0.2743(4)
C(17)	0.5825(8)	0.6639(7)	0.2363(4)
C(21)	0.4648(8)	0.2248(6)	0.1868(3)
C(22)	0.5554(8)	0.1387(7)	0.1696(3)
C(23)	0.5256(10)	0.0122(8)	0.1666(4)
C(24)	0.4045(10)	-0.0294(8)	0.1792(4)

Table 5.3.2 (continued)

C(25)	0.3166(9)	0.0545(8)	0.1955(4)
C(26)	0.3445(8)	0.1820(7)	0.2006(3)
C(27)	0.2468(8)	0.2677(8)	0.2212(4)
C(31)	0.6580(8)	0.4081(7)	0.1575(3)
C(32)	0.7749(8)	0.4418(7)	0.1869(3)
C(33)	0.8892(9)	0.4634(9)	0.1673(5)
C(34)	0.8828(11)	0.4515(9)	0.1135(6)
C(35)	0.7668(11)	0.4161(9)	0.0814(4)
C(36)	0.6526(9)	0.3935(8)	0.1036(4)
C(37)	0.5280(11)	0.3565(9)	0.0668(4)
C(41)	0.1443(14)	0.2360(12)	0.0655(5)
C(42)	0.1396(12)	0.3009(12)	0.0152(5)

Table 5.3.3: Anisotropic Thermal Parameters For The [(*o*-Tol)₃PAu(6-MP)].C₂H₅OH Complex.

Atom	U ₁₁	U ₂₂	U ₃₃	U ₁₂	U ₁₃	U ₂₃
Au	0.0425(2)	0.0445(2)	0.0442(2)	0.0005(2)	0.0064(2)	0.0099(1)
S(6)	0.040(1)	0.073(2)	0.099(2)	0.004(1)	0.011(1)	0.045(2)
P(1)	0.038(1)	0.037(1)	0.034(1)	-0.0008(9)	0.008(1)	0.0038(9)
O(41)	0.082(5)	0.104(5)	0.042(5)	0.014(4)	0.012(4)	-0.009(4)
N(1)	0.063(5)	0.058(4)	0.048(5)	0.005(4)	0.016(4)	0.017(4)
N(3)	0.066(5)	0.067(5)	0.062(6)	0.005(4)	0.017(5)	0.025(4)
N(7)	0.042(4)	0.064(5)	0.064(6)	0.002(4)	0.009(4)	0.008(4)
N(9)	0.063(5)	0.066(5)	0.050(6)	-0.011(4)	0.015(4)	0.009(4)
C(2)	0.071(7)	0.064(6)	0.058(7)	0.003(5)	0.009(6)	0.019(5)
C(4)	0.058(6)	0.059(5)	0.039(6)	-0.006(5)	0.018(5)	-0.001(4)
C(5)	0.053(5)	0.049(5)	0.035(6)	-0.004(4)	0.006(4)	0.003(4)
C(6)	0.051(5)	0.054(5)	0.042(6)	0.000(4)	0.013(5)	0.009(4)
C(8)	0.058(6)	0.071(6)	0.050(7)	0.001(5)	0.014(5)	0.010(5)
C(11)	0.036(4)	0.045(4)	0.028(5)	-0.001(3)	0.010(4)	-0.001(3)
C(12)	0.058(6)	0.053(5)	0.042(6)	-0.006(4)	0.008(5)	0.006(4)
C(13)	0.074(7)	0.084(7)	0.031(6)	-0.015(6)	-0.000(5)	0.012(5)
C(14)	0.071(6)	0.108(8)	0.029(6)	-0.015(6)	0.010(5)	-0.015(6)
C(15)	0.064(6)	0.067(6)	0.036(6)	-0.006(5)	0.012(5)	-0.017(5)
C(16)	0.040(5)	0.047(5)	0.043(6)	-0.003(4)	0.009(4)	-0.004(4)
C(17)	0.059(6)	0.043(5)	0.062(7)	0.002(4)	0.007(5)	-0.006(4)
C(21)	0.049(5)	0.035(4)	0.034(5)	-0.004(3)	0.003(4)	0.001(3)
C(22)	0.054(5)	0.046(5)	0.044(6)	0.002(4)	0.012(5)	-0.000(4)
C(23)	0.090(7)	0.047(5)	0.057(7)	0.013(5)	0.025(6)	-0.005(5)
C(24)	0.088(7)	0.042(5)	0.048(6)	-0.011(5)	0.007(6)	-0.002(4)

Table 5.3.3 (continued)

C(25)	0.067(6)	0.054(5)	0.055(7)	-0.011(5)	0.017(5)	0.011(5)
C(26)	0.053(5)	0.044(4)	0.036(6)	-0.002(4)	0.013(4)	0.006(4)
C(27)	0.055(6)	0.063(5)	0.062(7)	-0.008(4)	0.034(5)	0.004(5)
C(31)	0.044(5)	0.042(4)	0.033(5)	-0.005(4)	0.019(4)	0.004(4)
C(32)	0.048(5)	0.047(4)	0.045(6)	0.001(4)	0.021(4)	-0.000(4)
C(33)	0.053(6)	0.068(6)	0.086(9)	-0.013(5)	0.017(6)	-0.005(6)
C(34)	0.074(8)	0.071(7)	0.10(1)	-0.016(6)	0.051(7)	-0.007(7)
C(35)	0.092(8)	0.079(7)	0.058(8)	-0.013(6)	0.045(7)	-0.007(6)
C(36)	0.057(6)	0.056(5)	0.052(7)	-0.009(4)	0.021(5)	-0.001(5)
C(37)	0.093(8)	0.095(7)	0.038(7)	-0.011(6)	0.017(6)	-0.005(5)
C(41)	0.16(1)	0.13(1)	0.07(1)	0.00(1)	0.02(1)	-0.022(8)
C(42)	0.088(9)	0.13(1)	0.056(9)	-0.014(7)	0.011(7)	0.004(8)

Table 5.3.4: Hydrogen Atom Parameters For The [(*o*-Tol)₃PAu(6-MP)].C₂H₅OH Complex.

Atom	x	y	z	B(eq)
H(2)	0.1419	0.9838	-0.0006	6.1
H(8)	0.6791	0.7269	0.0837	5.6
H(9)	0.6060	0.9110	0.0256	5.6
H(12)	0.5431	0.2487	0.2839	4.8
H(13)	0.5895	0.2972	0.3738	6.0
H(14)	0.6189	0.5063	0.4012	6.6
H(15)	0.6156	0.6677	0.3394	5.2
H(17a)	0.6714	0.7016	0.2420	5.2
H(17b)	0.5633	0.6321	0.2002	5.2
H(17c)	0.5164	0.7277	0.2420	5.2
H(22)	0.6394	0.1690	0.1598	4.5
H(23)	0.5892	-0.0480	0.1557	6.0
H(24)	0.3819	-0.1191	0.1764	5.6
H(25)	0.2314	0.0234	0.2038	5.5
H(27a)	0.2804	0.3543	0.2222	5.5
H(27b)	0.1613	0.2635	0.1981	5.5
H(27c)	0.2351	0.2413	0.2568	5.5
H(32)	0.7764	0.4510	0.2250	4.3
H(33)	0.9718	0.4862	0.1900	6.5
H(34)	0.9626	0.4686	0.0974	7.4
H(35)	0.7657	0.4071	0.0434	6.9
H(37a)	0.4551	0.3428	0.0872	7.1
H(37b)	0.5038	0.4240	0.0412	7.1
H(37c)	0.5446	0.2788	0.0482	11.0
H(41a)	0.0606	0.1908	0.0662	11.0

Table 5.3.4 *(continued)*

H(41b)	0.2182	0.1758	0.0697	11.0
H(41c)	0.1574	0.2974	0.0942	11.0
H(42a)	0.2230	0.3468	0.0148	8.6
H(42b)	0.0652	0.3605	0.0110	8.6

Table 5.3.5: Bond Distances (Å) For The [(*o*-Tol)₃PAu(6-MP)].C₂H₅OH Complex.

Atom	Atom	Distance	Atom	Atom	Distance
Au	– S(6)	2.266(2)	C(14)	– C(15)	1.36(1)
Au	– P(1)	2.239(2)	S(6)	– C(6)	1.715(8)
C(15)	– C(16)	1.35(1)	P(1)	– C(11)	1.809(8)
P(1)	– C(21)	1.808(7)	C(16)	– C(17)	1.49(1)
P(1)	– C(31)	1.808(8)	O(41)	– C(42)	1.40(1)
N(1)	– C(2)	1.32(1)	N(1)	– C(6)	1.343(9)
C(21)	– C(22)	1.40(1)	N(3)	– C(2)	1.30(1)
C(21)	– C(26)	1.38(1)	N(3)	– C(4)	1.33(1)
C(22)	– C(23)	1.36(1)	N(7)	– C(5)	1.36(1)
N(7)	– C(8)	1.30(1)	C(23)	– C(24)	1.38(1)
N(9)	– C(4)	1.35(1)	N(9)	– C(8)	1.37(1)
C(24)	– C(25)	1.36(1)	C(25)	– C(26)	1.37(1)
C(4)	– C(5)	1.37(1)	C(5)	– C(6)	1.37(1)
C(26)	– C(27)	1.49(1)	C(11)	– C(12)	1.39(1)
C(11)	– C(16)	1.40(1)	C(12)	– C(13)	1.36(1)
C(31)	– C(32)	1.35(1)	C(31)	– C(36)	1.37(1)
C(13)	– C(14)	1.35(1)	C(32)	– C(33)	1.34(1)
C(33)	– C(34)	1.37(1)	C(34)	– C(35)	1.37(1)
C(35)	– C(36)	1.37(1)	C(36)	– C(37)	1.50(1)
C(41)	– C(42)	1.44(2)			

Table 5.3.6: Bond Angles (°) For The [(*o*-Tol)₃PAu(6-MP)].C₂H₅OH Complex.

Atom	Atom	Atom	Angle	Atom	Atom	Atom	Angle
S(6)	– Au	– P(1)	177.03(8)	C(11)	– C(12)	– C(13)	120.3(8)
Au	– S(6)	– C(6)	105.3(3)	C(12)	– C(13)	– C(14)	120.2(9)
Au	– P(1)	– C(11)	112.0(2)	C(13)	– C(14)	– C(15)	120.3(9)
Au	– P(1)	– C(21)	114.6(3)	C(14)	– C(15)	– C(16)	121.8(8)
Au	– P(1)	– C(31)	110.7(3)	C(11)	– C(16)	– C(15)	118.5(8)
C(11)	– P(1)	– C(21)	105.6(4)	C(11)	– C(16)	– C(17)	122.1(8)
C(11)	– P(1)	– C(31)	106.1(4)	C(15)	– C(16)	– C(17)	119.4(8)
C(21)	– P(1)	– C(31)	107.3(4)	P(1)	– C(21)	– C(22)	118.0(6)
C(2)	– N(1)	– C(6)	118.9(8)	P(1)	– C(21)	– C(26)	121.7(6)
C(2)	– N(3)	– C(4)	110.9(8)	C(22)	– C(21)	– C(26)	120.2(7)
C(5)	– N(7)	– C(8)	103.4(7)	C(21)	– C(22)	– C(23)	120.0(8)
C(4)	– N(9)	– C(8)	105.8(7)	C(22)	– C(23)	– C(24)	119.5(8)
N(1)	– C(2)	– N(3)	128.7(9)	C(23)	– C(24)	– C(25)	120.2(8)
N(3)	– C(4)	– N(9)	127.4(8)	C(24)	– C(25)	– C(26)	122.1(9)
N(3)	– C(4)	– C(5)	127.4(9)	C(21)	– C(26)	– C(25)	117.9(8)
N(9)	– C(4)	– C(5)	105.2(8)	C(21)	– C(26)	– C(27)	122.8(7)
N(7)	– C(5)	– C(4)	111.7(8)	C(25)	– C(26)	– C(27)	119.3(8)
N(7)	– C(5)	– C(6)	132.3(8)	P(1)	– C(31)	– C(32)	120.5(6)
C(4)	– C(5)	– C(6)	115.9(8)	P(1)	– C(31)	– C(36)	120.2(6)
S(6)	– C(6)	– N(1)	115.8(6)	C(32)	– C(31)	– C(36)	119.2(8)
S(6)	– C(6)	– C(5)	125.9(7)	C(31)	– C(32)	– C(33)	124.6(9)
N(1)	– C(6)	– C(5)	118.2(8)	C(32)	– C(33)	– C(34)	116(1)
N(7)	– C(8)	– N(9)	113.8(8)	C(33)	– C(34)	– C(35)	122.0(9)
P(1)	– C(11)	– C(12)	118.9(6)	C(34)	– C(35)	– C(36)	120(1)
P(1)	– C(11)	– C(16)	122.1(6)	C(31)	– C(36)	– C(35)	118.5(9)
C(12)	– C(11)	– C(16)	118.9(8)	C(31)	– C(36)	– C(37)	123.9(8)
C(35)	– C(36)	– C(37)	118(1)	O(41)	– C(42)	– C(41)	110(1)

Table 5.3.7 (continued)

Plane number 2: Least-squares plane through the *ortho*-tolyl ring defined by the atoms C(11) to C(16).

Atoms Defining Plane	Distance (Å)	esd (Å)
C(11)	-0.0041	0.0071
C(12)	-0.0017	0.0086
C(13)	0.0106	0.0100
C(14)	-0.0089	0.0099
C(15)	-0.0011	0.0091
C(16)	0.0056	0.0076
Additional Atoms	Distance (Å)	
P(1)	-0.1466	
C(17)	0.0901	

Mean deviation from plane is 0.0053 (Å).

Chi-squared: 2.5.

Table 5.3.7: Mean Plane Data For The [(*o*-Tol)₃PAu(6-MP)].C₂H₅OH Complex.

Plane number 1: Least-squares plane through the 6-mercaptopurine moiety.

Atoms Defining Plane	Distance (Å)	esd (Å)
S(6)	0.0134	0.0031
N(1)	-0.0479	0.0076
N(3)	0.0324	0.0083
N(7)	-0.0363	0.0077
N(9)	0.0282	0.0076
C(2)	0.0058	0.0104
C(4)	0.0196	0.0086
C(5)	-0.0223	0.0082
C(6)	-0.0503	0.0086
C(8)	0.0016	0.0097
Additional Atom	Distance (Å)	
Au	-0.0444	

Mean deviation from plane is 0.0258 (Å).

Chi-squared: 156.3.

Table 5.3.7: (continued)

Plane number 3: Least-squares plane through the *ortho*-tolyl ring defined by the atoms C(21) to C(26).

Atoms Defining Plane	Distance (Å)	esd (Å)
C(21)	-0.0039	0.0078
C(22)	-0.0062	0.0084
C(23)	0.0131	0.0098
C(24)	-0.0029	0.0092
C(25)	-0.0105	0.0094
C(26)	0.0111	0.0082
Additional Atoms	Distance (Å)	
C(27)	0.0711	
P(1)	-0.0856	

Mean deviation from plane is 0.0080 (Å).

Chi-squared: 5.3.

Table 5.3.7 (*continued*)

Plane number 4: Least-squares plane through the *ortho*-tolyl ring defined by the atoms C(31) to C(36).

Atoms Defining Plane	Distance (Å)	esd (Å)
C(31)	-0.0048	0.0070
C(32)	0.0001	0.0074
C(33)	0.0086	0.0093
C(34)	-0.0093	0.0100
C(35)	-0.0004	0.0099
C(36)	0.0070	0.0083
Additional Atoms	Distance (Å)	
P(1)	-0.1298	
C(37)	0.0010	

Mean deviation from plane is 0.0050 (Å).

Chi-squared: 2.6.

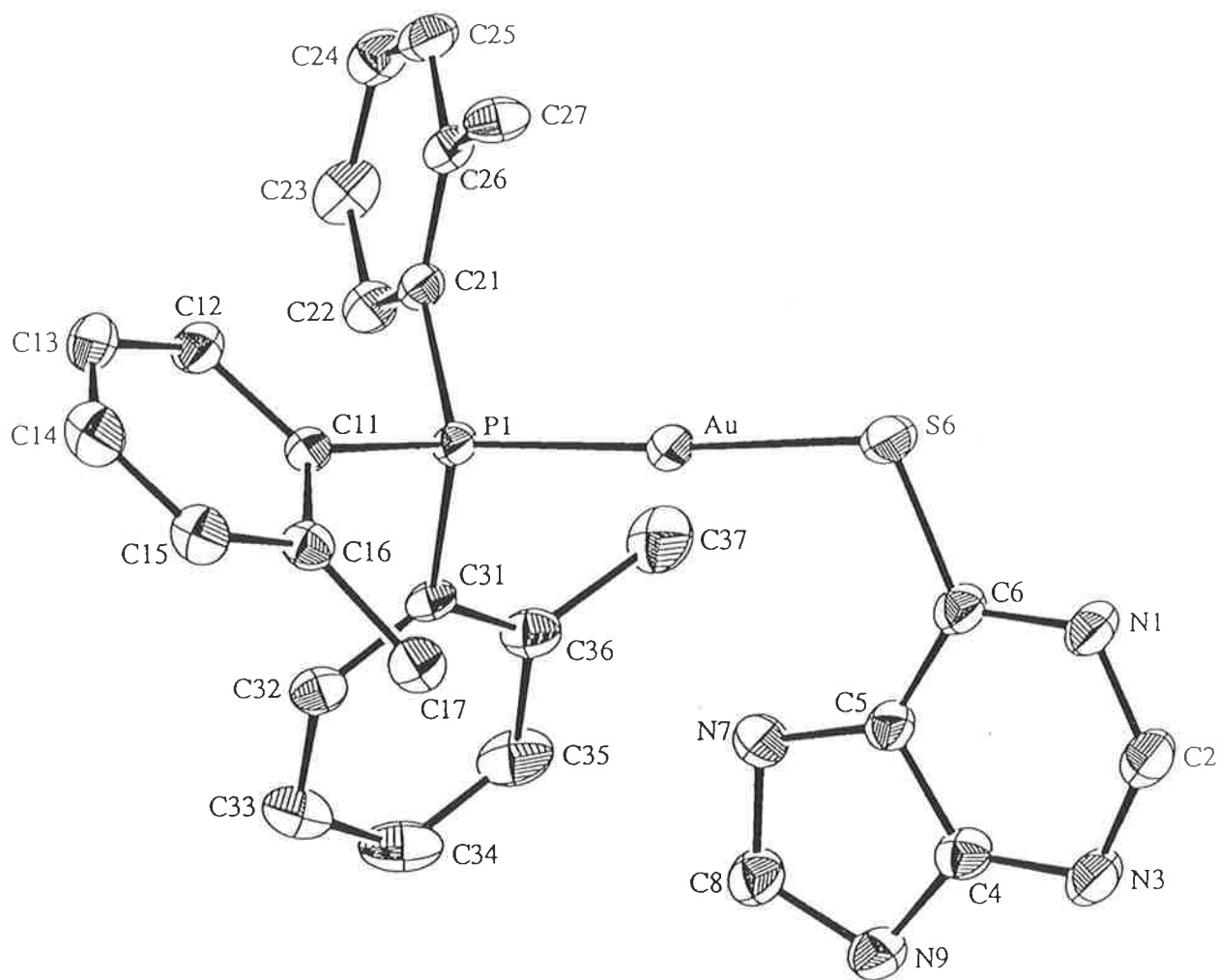


Figure 5.3.1: *Molecular Structure And Crystallographic Numbering Scheme For $[(o\text{-Tol})_3\text{PAu}(6\text{-MP})]$.*

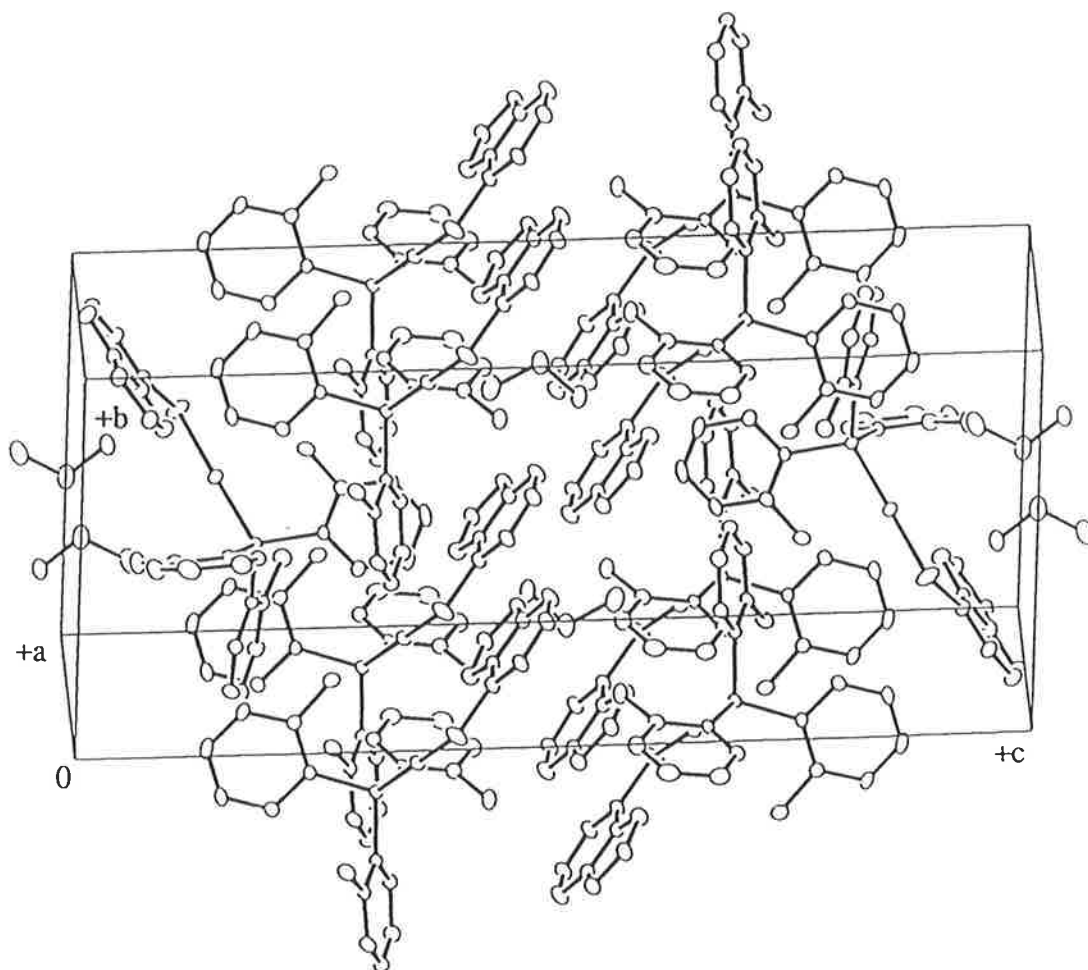


Figure 5.3.2: Unit Cell Diagram Of $[(o\text{-Tol})_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$.

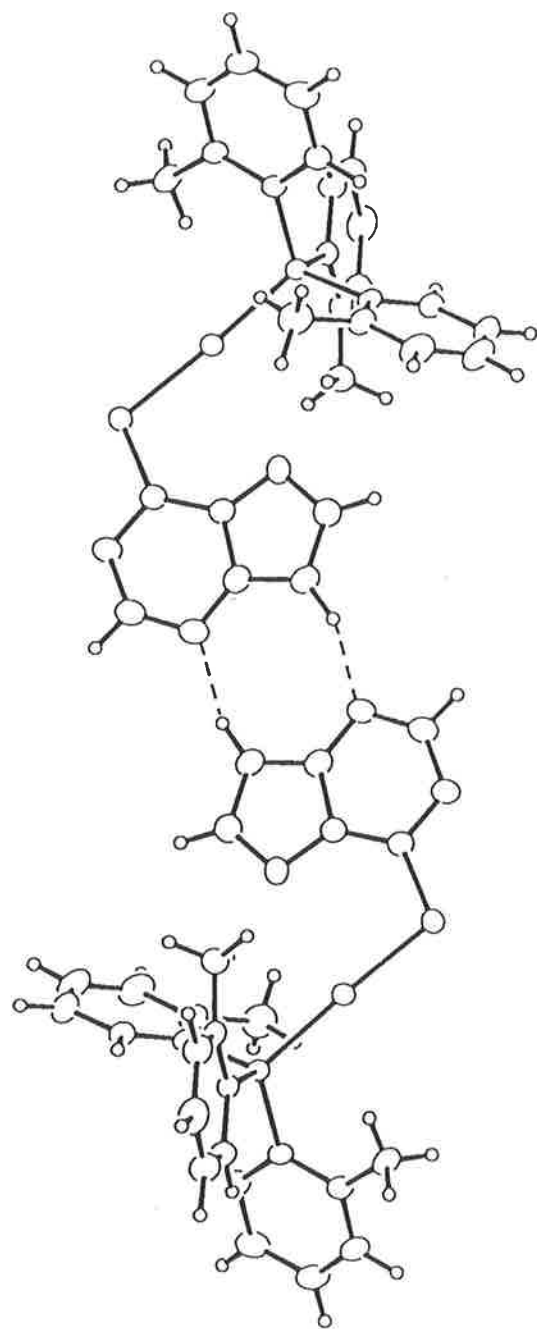


Figure 5.3.3: *Diagram Illustrating Intermolecular Hydrogen Bonding (indicated by dashed lines).*

The P(1)–Au distance is 2.239(2) Å, which is equivalent within standard deviation to the value of 2.243(2) Å for [(*o*-Tol)₃PAuCl]⁵⁵. It is also equivalent to the corresponding value of 2.237(2) Å for [Ph₃PAu(6-MP)], suggesting that not only does the coordination of the 6-MP ligand to the gold centre have little electronic effect on the phosphine moiety but also that the size of the phosphine cone angle affects the P–Au bond to a negligible extent. This latter observation is in accordance with the results obtained from the cone angle correlation for triorganophosphinegold(I) chlorides in Chapter 3. The P–C bond distances also display equivalent values.

The *ortho*-tolyl rings are all planar; the maximum deviation from planarity is observed for the ring defined by the atoms C(21) to C(26), of the value 0.008(9) Å. The phosphorus atom lies out of the plane to a maximum value of 0.147 Å for the ring defined by the atoms C(11) to C(16). The methyl groups on all three ring systems only deviate from the plane by less than 0.1 Å, which is expected as the C(16), C(26) and C(36) atoms are all sp² hybridized in the aromatic system, as evidenced by the corresponding angles all being equal to 120° within standard deviation.

The Au–S bond length is 2.266(2) Å, a value which is significantly shorter than 2.287(1) Å for [Ph₃PAu(6-MP)]. Again, this may be attributable to the relative cone-angles of (*o*-Tol)₃P and Ph₃P, which are 194 and 145° respectively⁴⁹, possibly giving rise to differing steric interactions. The effect of coordination on the 6-mercaptapurinate moiety is discussed in the next section.

5.4 Comparison between the 6-mercaptapurinate moiety in the free ligand and in the complexes.

Crystallographic data for the free 6-mercaptapurine ligand has been obtained from the reference of Sletten, Sletten and Jensen⁵⁷ on the analysis of 6-mercaptapurine monohydrate. Tables 5.4.1 and 5.4.2 give the bond distances and bond angles for 6-mercaptapurine monohydrate. These values are also given pictorially in Figures 5.4.1 and 5.4.2. Alongside

Table 5.4.1: Bond Distances (Å) For 6-mercaptopurine Monohydrate⁵⁷.

Atom	Atom	Distance	Atom	Atom	Distance
N(1)	– C(2)	1.350(2)	C(6)	– S(6)	1.676(2)
C(2)	– N(3)	1.307(2)	C(5)	– N(7)	1.370(2)
N(3)	– C(4)	1.364(2)	N(7)	– C(8)	1.346(2)
C(4)	– C(5)	1.397(2)	C(8)	– N(9)	1.326(2)
C(5)	– C(6)	1.396(2)	N(9)	– C(4)	1.363(2)
C(6)	– N(1)	1.384(2)			

Table 5.4.2: Bond Angles (°) For 6-mercaptopurine Monohydrate⁵⁷.

Atom	Atom	Atom	Angle	Atom	Atom	Atom	Angle
C(6)	– N(1)	– C(2)	125.4(1)	C(6)	– C(5)	– N(7)	132.2(1)
N(1)	– C(2)	– N(3)	125.1(1)	C(4)	– C(5)	– N(7)	105.9(1)
C(2)	– N(3)	– C(4)	113.0(1)	C(5)	– N(7)	– C(8)	106.1(1)
N(3)	– C(4)	– C(5)	124.2(1)	N(7)	– C(8)	– N(9)	113.6(1)
C(4)	– C(5)	– C(6)	121.9(1)	C(8)	– N(9)	– C(4)	104.5(1)
C(5)	– C(6)	– N(1)	110.4(1)	N(9)	– C(4)	– C(5)	109.9(1)
C(5)	– C(6)	– S(6)	127.0(1)	N(9)	– C(4)	– N(3)	125.9(1)
N(1)	– C(6)	– S(6)	122.6(1)				

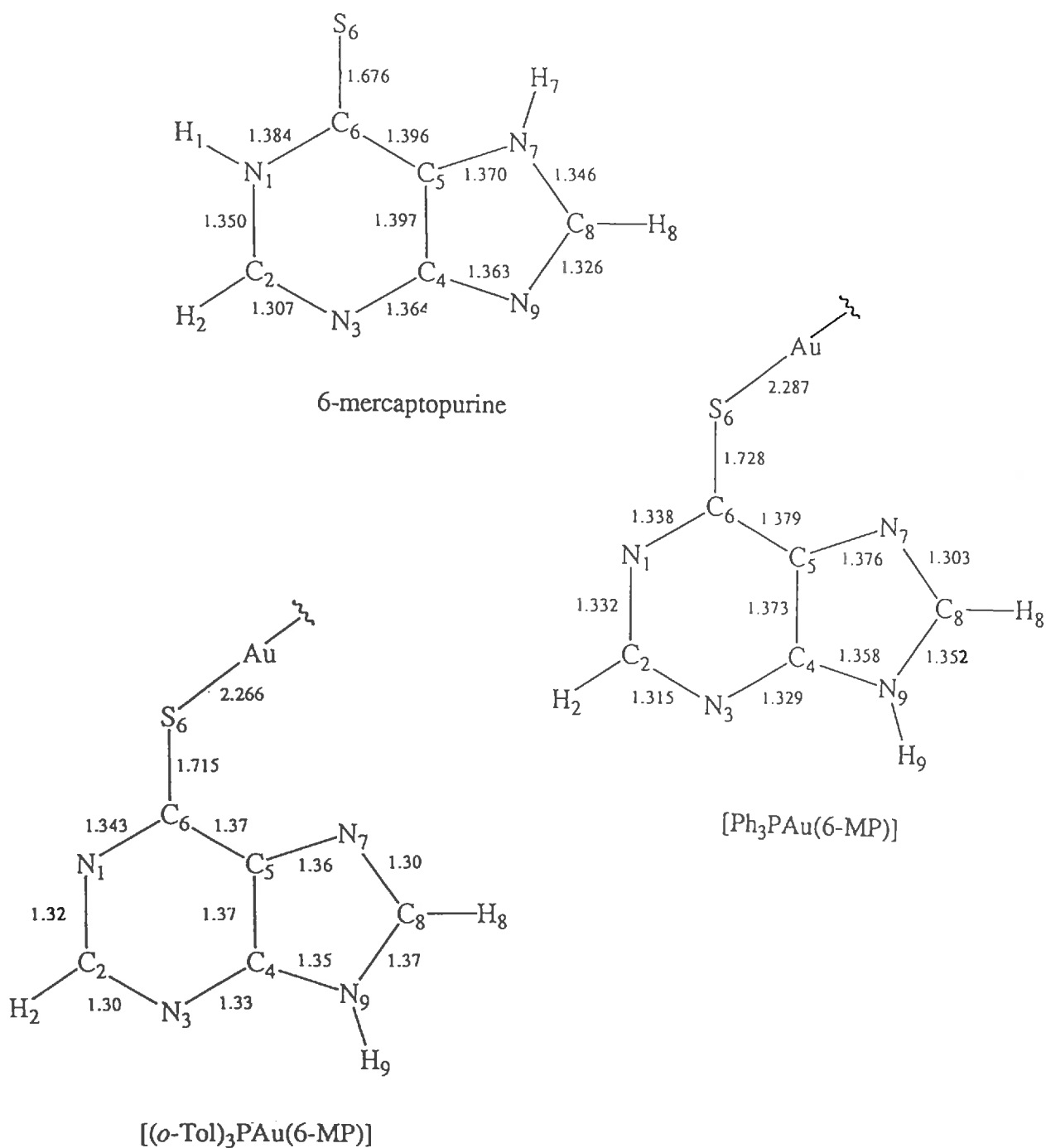


Figure 5.4.1: Schematic Representation Of The Bond Distances (Å) In The Purine Moiety Of 6-mercaptapurine Monohydrate, $[\text{Ph}_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$ and $[(o\text{-Tol})_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$.

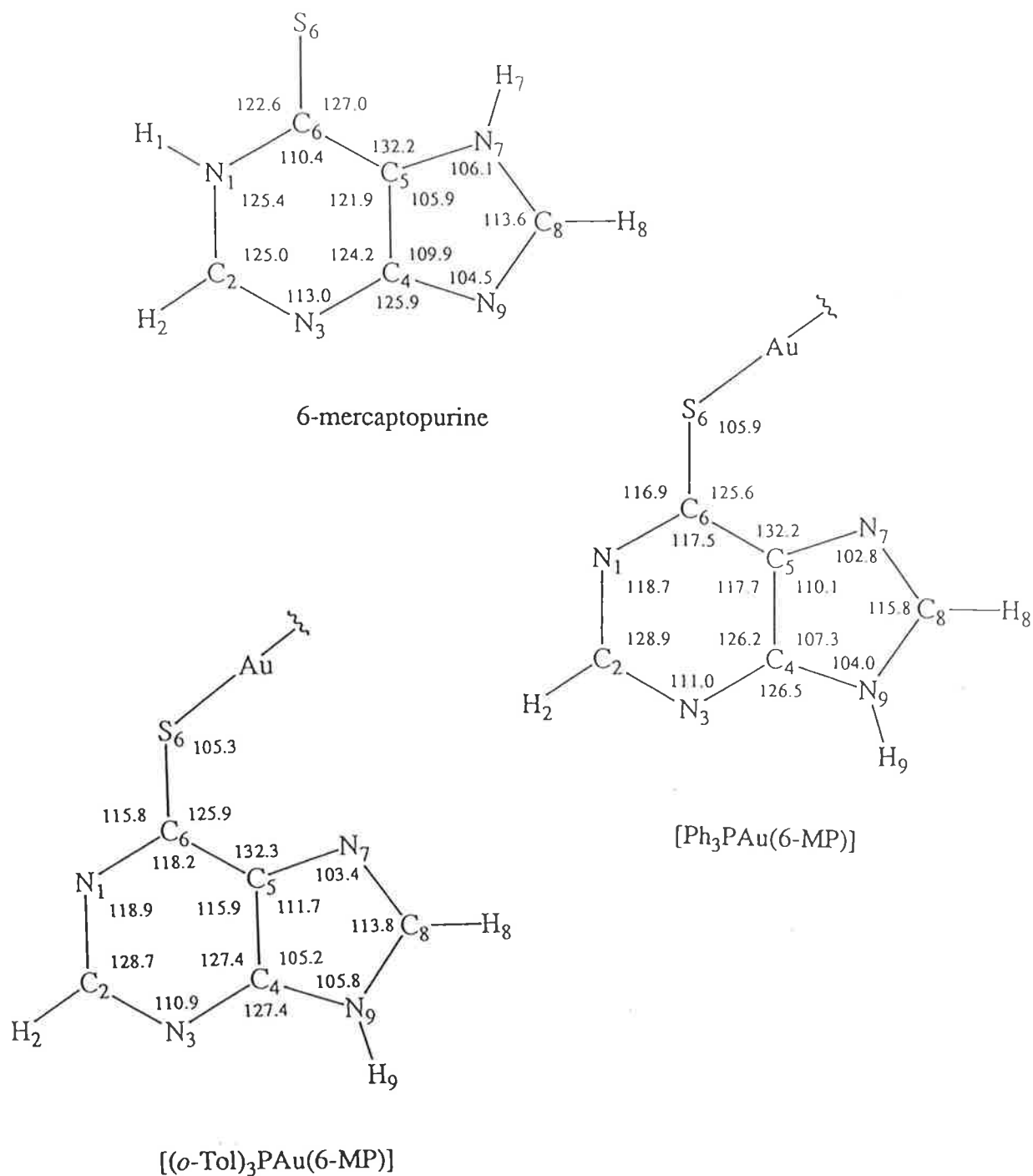


Figure 5.4.2: Schematic Representation Of The Bond Angles (°) In The Purine Moiety Of 6-mercaptopurine Monohydrate, [Ph₃PAu(6-MP)].C₂H₅OH and [(*o*-Tol)₃PAu(6-MP)].C₂H₅OH.

these diagrams are shown the 6-mercaptapurinate regions of the two complexes. The estimated standard deviations for the values of the complexes have not been given for reasons of clarity but they can be found in the tables on the previous pages.

Analysis of Figure 5.4.1 indicates how the internal molecular structure of the ring system is altered upon complexation of the 6-mercaptapurinate ligand to the gold atom via the sulphur atom. The most significant observation is that concerning the C(6)–S(6) bond. In 6-MPH, the C(6)–S(6) bond length is 1.676(2) Å, a value indicative of its significant double-bond character. Upon complexation, the length of this bond increases appreciably, to 1.728(5) Å for [Ph₃PAu(6-MP)] and to 1.715(8) Å for [(*o*-Tol)₃PAu(6-MP)]. Hence, as expected the C(6)–S(6) bond approaches single-bond character as the sulphur atom binds to the gold atom. The general increase in electron delocalization in the six-membered ring of 6-mercaptapurine upon complexation leads to the N(1)–C(6) and C(5)–C(6) bonds increasing in multiple-bond character, which in turn lessens the bond order of the C(6)–S(6) interaction. All the bonds in the six-membered ring system except for C(2)–N(3) demonstrate a decrease in length; N(1)–C(6) undergoes the greatest change. In the five-membered ring, the bonds of C(5)–N(7) and C(4)–N(9) vary only slightly. An increase in the bond length of C(8)–N(9) and a decrease in N(7)–C(8) is observed; the bond character in these two bonds has essentially been exchanged from the free ligand to the complex. The movement in electron density thus suggests that it is N(9) which is protonated in the complexes, whereas N(7) was protonated in the free ligand, and justifies the placement of H(9) at this position, with the N–H bond distance fixed at 0.95 Å.

The change in the internal angles, as illustrated in Figure 5.4.2, confirms what the previous paragraph suggested. The internal angles of the six-membered ring all appear to converge to 120°, consistent with an increase in aromaticity in the ring. For both complexes the angles C(4)–C(5)–N(7) and C(5)–C(4)–N(9) increase and decrease in size respectively upon complexation, which is consistent with the movement of the amino hydrogen from N(7) to N(9).

The arguments above are consistent with the data obtained via the spectral characterization of the complexes. Contraction of N(1)–C(6) upon complexation, which indicates an increase in double bond character, was manifested in the infrared spectra by a higher absorption frequency of the thioamide band I vibrational mode. Similarly, the decrease in double bond character of C(6)–S(6) is reflected by the decreasing absorption frequency of the thioamide band II chromophore in the complexes. The observed overall contraction in size of the ring system also supports the arguments made in Chapter 4 concerning the electronic effects of complexation via sulphur on the resonance frequencies of the carbon-13 nuclei. Hence, as these interpretations are now vindicated by the crystallographic evidence, then the identity of all the other complexes prepared can be confirmed by extrapolation.

Even with the inclusion of the sulphur atom as part of the ring system, the purine group in each of the molecules is planar; the mean deviation from the mean planes for [Ph₃PAu(6-MP)] and [(*o*-Tol)₃PAu(6-MP)] being 0.012 and 0.026 Å respectively. When excluded, calculations show that the system appears to be more planar e.g. in [Ph₃PAu(6-MP)] the mean deviation is 0.009 Å. Either way, the mean deviations are small.

All the effects noticed are expected for complexation to gold(I) via the sulphur atom, consistent with the structures of other triorganophosphinegold(I) thiolate complexes. However, a variety of coordination modes exist for crystal structures involving 6-mercaptapurine-type thionucleobases in the literature. One example is coordination via N(9) e.g. 6-mercaptapurine riboside⁷³, and another common type is the ligand acting in bidentate fashion, coordinating to the metal atom via the S(6) and N(7) atoms, e.g. dichloro-(6-mercapto-9-methylpurine)copper(II)⁷⁴, (6-mercaptapurine)diaquodcadmium(II)⁷⁵ and bis(6-mercapto-9-benzylpurine)palladium(II)-dimethylacetamide⁷⁶. In these cases the internal structure of the purine moiety is similar to those discussed above given the changes. An example of metal to sulphur coordination has been reported for dichloro(6-mercaptapurinium)copper(I)¹³, where the 6-MP is in a protonated form; the effects here on the internal structure of the

6-mercaptapurinate moiety are many ways the opposite to the effects observed for deprotonation.

The crystallographic investigations of $[\text{Ph}_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$ and $[(o\text{-Tol})_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$ has revealed a number of observations. As expected, the P–Au–S chromophore is a near linear group, consistent with previously observed data for analogous structures. The bond distances within the 6-mercaptapurinate moiety undergo a general contraction upon complexation, a result which is expected as the electron density increases within the ring structure. The ligand coordinates exclusively in the monodentate mode; the N(7) atom is oriented towards the gold centre in the solid state, although not at a intramolecular distance indicative of a significant interaction. The spectroscopic data concerning these complexes is consistent with the observed intramolecular parameters.

CHAPTER 6

Crystallographic Investigations of Triorganophosphinegold(I) Thiolates

6.1 Introduction

There are now several examples of crystal structures of triorganophosphinegold(I) thiolates with the general formula $[R_3PAu(SR')]$ in the literature. Common to all these complexes is the SR' moiety coordinating in a monodentate mode, as a thiolate ligand, i.e. with an $Au-S-R'$ linkage, and an invariably linear $P-Au-S$ chromophore. Such structural characteristics were also observed in Chapter 5 for the $[Ph_3PAu(6-MP)]$ and $[(o-Tol)_3PAu(6-MP)]$ complexes. Thus, in a similar fashion to the triorganophosphinegold(I) chlorides analyzed in Chapter 3, a study may be performed on these complexes to determine whether any simple relationship exists between the steric effects of the phosphine and thiolate groups, and the intramolecular parameters of $P-Au$, $Au-S$, $P-Au-S$ and possibly the $Au...N$ interaction observed in most of the complexes; such a study is performed in this chapter. In order to gain additional information on complexes possessing the $Cycl_3P$ phosphine group, the crystal structure determination of the complex 6-n-propyl-2-thiouracilato(tricyclohexylphosphine)gold(I), $[Cycl_3PAu(6p2-TU)]$, was performed.

6.2 Crystal structure of the $[Cycl_3PAu(6p2-TU)]$ complex

The complex $[Cycl_3PAu(6p2-TU)]$ was prepared by the literature method¹⁹; crystals of the complex were obtained from the slow evaporation of a saturated ethanolic solution of the compound, and the complex crystallizes in the monoclinic space group $P2_1/c$ (C_{2h}^5 , No. 14)⁴⁴.

Table 6.2.1: Crystallographic Parameters For The [Cycl₃PAu(6p2-TU)] Complex.

Data	[Cycl ₃ PAu(6p2-TU)]
Formula	C ₂₅ H ₄₂ AuSPON ₂
Formula weight	646.6
Crystal shape	rectangular prism
Crystal dimensions (mm)	0.16 x 0.16 x 0.29
Crystal system	monoclinic
Space group	P2 ₁ /c (C _{2h} ⁵ , No. 14)
<i>a</i> (Å)	9.539(2)
<i>b</i> (Å)	16.452(4)
<i>c</i> (Å)	16.880(2)
β (°)	95.37(2)
<i>V</i> (Å ³)	2637.4(8)
<i>Z</i>	4
$\rho_{\text{calc.}}$ (g cm ⁻³)	1.628
F(000)	324 1296
μ (cm ⁻¹)	57.56
θ limits, cell (°)	13.1 to 15.4
θ limits, data (°)	1.5 to 27.9
<i>hkl</i> range	0 to 12, 0 to 21, -20 to 20
Range of transmission factors	0.985 to 1.012
Scan technique	$\omega:2\theta$
No. of data measured	6899
No. of unique data	6524
<i>R</i> _{amal}	0.060
No. of unique data used	3695
Criterion of observability	$I \geq 3.0\sigma(I)$
No. of parameters	280
<i>R</i>	0.043
<i>R</i> _w	0.039
Residual electron density (e Å ⁻³)	-1.28 to 0.90

Crystal and refinement data are listed in Table 6.2.1, and the derived results are given in Tables 6.2.2 to 6.2.7. The observed and calculated structure factors can be found in the Appendix. The crystallographic numbering scheme is shown in the ORTEP⁴⁵ diagram in Figure 6.2.1, plotted with 30 % thermal ellipsoids.

The unit cell contents of [Cycl₃PAu(6p2-TU)] are displayed in Figure 6.2.2 and are clearly demonstrative of the P2₁/c space group. As observed in other triorganophosphinegold(I) derivatives containing the thiouracilate anion^{9,19,61}, centrosymmetrically related molecules are associated via hydrogen bonding contacts in the lattice. They involve the N(3)–H(3) and O(4') atoms such that the H(3)...O(4') separation is 1.90 Å and the N(3)–H(3)...O(4') angle is 161° (symmetry operation for O(4') is -x, 1-y, -z). The closest gold to gold interaction is 8.460(1) Å, which is too large to be considered as a 'significant' interaction⁴⁶.

The molecular structure of [Cycl₃PAu(6p2-TU)] is shown in Figure 6.2.2. As expected, the P–Au–S chromophore is linear, the angle being 177.6(1)°. The P(1)–Au and Au–S(2) bond lengths are 2.248(3) and 2.302(3) Å respectively. These values are equivalent to those observed for the related complex [Cycl₃PAu(6m2-TU)]⁷⁷; the implications of this will be discussed in the next section. The three P–C bond lengths are all equivalent to within standard deviation, suggesting the electronic environment about these bonds is identical. The cyclohexyl groups exist in the 'chair' conformation; relatively large thermal motion is associated with these atoms.

The 6-n-propyl-2-thiouracilate moiety is coordinated as expected to the gold centre via the sulphur atom, as a thiolate rather than a thione. The two bond distances of N(1)–C(2) and N(3)–C(2) are 1.28(1) and 1.37(1) Å respectively, indicating the former has greater double bond character. Hence deprotonation of the free ligand occurs at N(1), consistent with results observed for other triorganophosphinegold(I) thiouracilates^{9,19,62,78}. The distance between N(1) and Au is 3.061(9) Å, less than the sum of the van der Waal radii, but, like with the two 6-mercaptopurinate complexes in Chapter 5 and other analogous complexes, not indicative of a

Table 6.2.2: Fractional Atomic Coordinates For The [Cycl₃PAu(6p2-TU)] Complex.

Atom	x	y	z
Au	0.40117(4)	0.71525(3)	0.22652(3)
S(2)	0.2653(3)	0.6070(2)	0.1792(2)
P(1)	0.5417(3)	0.8183(2)	0.2714(2)
O(4)	-0.0151(9)	0.5635(5)	-0.0732(5)
N(1)	0.2145(9)	0.7192(6)	0.0689(6)
N(3)	0.1137(8)	0.5937(5)	0.0415(5)
C(2)	0.1957(9)	0.6458(7)	0.0896(6)
C(4)	0.0525(13)	0.6143(8)	-0.0323(8)
C(5)	0.0743(15)	0.6972(9)	-0.0507(8)
C(6)	0.1523(14)	0.7461(9)	-0.0009(9)
C(11)	0.7079(10)	0.8113(6)	0.2279(6)
C(12)	0.8076(11)	0.8823(7)	0.2425(7)
C(13)	0.9317(12)	0.8727(8)	0.1928(8)
C(14)	1.0081(12)	0.7950(10)	0.2090(8)
C(15)	0.9106(13)	0.7232(8)	0.2002(7)
C(16)	0.7821(11)	0.7306(7)	0.2450(6)
C(21)	0.5778(14)	0.8208(8)	0.3794(7)
C(22)	0.6144(14)	0.9027(8)	0.4149(8)
C(23)	0.6442(26)	0.8949(11)	0.5056(10)
C(24)	0.6002(24)	0.8381(15)	0.5457(10)
C(25)	0.5671(18)	0.7615(10)	0.5100(9)
C(26)	0.5342(25)	0.7634(11)	0.4218(9)
C(31)	0.4687(10)	0.9161(6)	0.2395(6)
C(32)	0.3266(12)	0.9291(7)	0.2674(8)
C(33)	0.2670(12)	1.0126(8)	0.2380(11)

Table 6.2.2 (continued)

C(34)	0.2619(17)	1.0240(9)	0.1533(10)
C(35)	0.4042(15)	1.0090(9)	0.1268(8)
C(36)	0.4580(12)	0.9267(7)	0.1505(7)
C(61)	0.1764(21)	0.8342(10)	-0.0176(11)
C(62)	0.1260(26)	0.8678(12)	-0.0779(16)
C(63)	0.1669(22)	0.9509(11)	-0.0911(14)

Table 6.2.3: Anisotropic Thermal Parameters For The [Cycl₃PAu(6p2-TU)] Complex.

Atom	U ₁₁	U ₂₂	U ₃₃	U ₁₂	U ₁₃	U ₂₃
Au	0.0382(2)	0.0390(2)	0.0547(2)	-0.0046(3)	-0.0015(2)	-0.0068(3)
S(2)	0.052(2)	0.042(2)	0.065(2)	-0.007(1)	-0.013(2)	0.000(1)
P(1)	0.042(2)	0.038(2)	0.046(2)	-0.005(1)	0.003(1)	-0.006(1)
O(4)	0.108(7)	0.074(7)	0.091(7)	-0.033(6)	-0.047(6)	0.013(6)
N(1)	0.060(6)	0.046(6)	0.075(7)	-0.013(6)	-0.011(5)	0.008(6)
N(3)	0.043(6)	0.041(6)	0.066(7)	-0.005(4)	-0.013(5)	0.005(5)
C(2)	0.024(6)	0.044(7)	0.067(8)	-0.000(5)	-0.003(5)	-0.012(6)
C(4)	0.060(9)	0.07(1)	0.07(1)	-0.015(7)	-0.023(7)	-0.002(8)
C(5)	0.11(1)	0.07(1)	0.10(1)	-0.02(1)	-0.04(1)	0.04(1)
C(6)	0.09(1)	0.07(1)	0.10(1)	-0.017(8)	-0.032(9)	0.011(9)
C(11)	0.037(6)	0.048(7)	0.049(7)	-0.001(5)	0.005(5)	0.005(5)
C(12)	0.036(7)	0.064(9)	0.09(1)	-0.007(6)	0.008(6)	-0.010(7)
C(13)	0.048(9)	0.07(1)	0.13(1)	-0.009(7)	0.029(8)	-0.00(1)
C(14)	0.052(8)	0.10(1)	0.13(1)	-0.012(9)	0.034(8)	-0.04(1)
C(15)	0.068(9)	0.08(1)	0.09(1)	0.017(9)	0.024(8)	-0.015(9)
C(16)	0.052(8)	0.06(1)	0.068(8)	0.013(6)	0.017(6)	-0.007(7)
C(21)	0.13(1)	0.062(9)	0.049(8)	-0.044(8)	0.021(8)	-0.016(7)
C(22)	0.12(1)	0.08(1)	0.06(1)	-0.013(9)	0.004(8)	-0.009(8)
C(23)	0.41(3)	0.10(2)	0.06(1)	-0.06(2)	0.00(2)	-0.04(1)
C(24)	0.23(2)	0.23(3)	0.06(1)	-0.09(2)	0.01(1)	0.06(2)
C(25)	0.16(2)	0.10(1)	0.08(1)	-0.06(1)	-0.00(1)	0.01(1)
C(26)	0.43(3)	0.16(2)	0.04(1)	-0.16(2)	0.01(1)	0.02(1)
C(31)	0.043(7)	0.046(7)	0.051(7)	-0.005(6)	0.007(6)	-0.013(6)
C(32)	0.050(8)	0.061(9)	0.11(1)	0.004(7)	0.001(8)	-0.023(8)
C(33)	0.043(9)	0.047(9)	0.22(2)	0.017(7)	-0.01(1)	-0.04(1)

Table 6.2.3 (*continued*)

C(34)	0.12(1)	0.07(1)	0.12(1)	0.02(1)	-0.06(1)	0.01(1)
C(35)	0.10(1)	0.08(1)	0.10(1)	0.02(1)	-0.02(1)	0.024(9)
C(36)	0.070(9)	0.066(9)	0.066(9)	0.021(7)	-0.009(7)	0.007(7)
C(61)	0.25(2)	0.07(1)	0.15(2)	-0.08(1)	-0.12(2)	0.07(1)
C(62)	0.30(3)	0.11(2)	0.30(3)	-0.07(2)	-0.12(2)	0.13(2)
C(63)	0.28(3)	0.09(2)	0.29(3)	-0.04(2)	-0.03(2)	0.11(2)

Table 6.2.4: *Hydrogen Atom Parameters For The [Cycl₃PAu(6p²-TU)] Complex.*

Atom	x	y	z	B(eq)
H(3)	0.0983	0.5403	0.0606	5.2
H(5)	0.0309	0.7187	-0.1007	9.1
H(11)	0.6835	0.8113	0.1710	4.5
H(12a)	0.7592	0.9325	0.2277	6.1
H(12b)	0.8429	0.8843	0.2984	6.1
H(13a)	0.8977	0.8745	0.1369	7.8
H(13b)	0.9975	0.9173	0.2047	7.8
H(14a)	1.0794	0.7895	0.1719	8.2
H(14b)	1.0532	0.7962	0.2629	8.2
H(15a)	0.8807	0.7166	0.1441	7.4
H(15b)	0.9637	0.6755	0.2193	7.4
H(16a)	0.8115	0.7267	0.3016	5.7
H(16b)	0.7187	0.6867	0.2291	19.8
H(21)	0.6743	0.8015	0.3805	7.2
H(22a)	0.6973	0.9239	0.3922	7.7
H(22b)	0.5365	0.9396	0.4033	7.7
H(23a)	0.7488	0.8927	0.5152	16.4
H(23b)	0.6130	0.9448	0.5278	4.2
H(24a)	0.5164	0.8585	0.5663	14.9
H(24b)	0.6726	0.8285	0.5896	14.9
H(25a)	0.4859	0.7404	0.5333	10.0
H(25b)	0.6471	0.7256	0.5225	10.0
H(26a)	0.4335	0.7627	0.4121	18.0
H(26b)	0.5735	0.7136	0.4020	18.0
H(31)	0.5304	0.9582	0.2629	4.5

Table 6.2.4 (continued)

H(32b)	0.2628	0.8864	0.2463	7.2
H(32a)	0.3329	0.9276	0.3251	7.2
H(33a)	0.1719	1.0180	0.2545	8.7
H(33b)	0.3260	1.0548	0.2635	8.7
H(34a)	0.1927	0.9860	0.1275	9.4
H(34b)	0.2300	1.0791	0.1405	9.4
H(35a)	0.3970	1.0137	0.0694	8.7
H(35b)	0.4671	1.0498	0.1506	8.7
H(36a)	0.5502	0.9201	0.1321	6.5
H(36b)	0.3943	0.8859	0.1260	7.2
H(61a)	0.1410	0.8639	0.0265	15.1
H(61b)	0.2761	0.8412	-0.0167	15.1
H(62a)	0.1496	0.8361	-0.1226	19.8
H(62b)	0.0226	0.8671	-0.0765	19.8
H(63a)	0.2661	0.9548	-0.0876	17.8
H(63b)	0.1255	0.9678	-0.1439	17.8
H(63c)	0.1281	0.9857	-0.0515	20.8

Table 6.2.5: Bond Distances (Å) For The [Cycl₃PAu(6p2-TU)] Complex.

Atom	Atom	Distance	Atom	Atom	Distance
Au	– S(2)	2.302(3)	C(13)	– C(14)	1.48(2)
Au	– P(1)	2.248(3)	C(14)	– C(15)	1.50(2)
S(2)	– C(2)	1.72(1)	C(32)	– C(33)	1.55(2)
P(1)	– C(11)	1.81(1)	C(33)	– C(34)	1.44(2)
P(1)	– C(21)	1.82(1)	C(34)	– C(35)	1.49(2)
P(1)	– C(31)	1.82(1)	C(35)	– C(36)	1.49(2)
O(4)	– C(4)	1.23(1)	C(61)	– C(62)	1.22(2)
N(1)	– C(2)	1.28(1)	C(62)	– C(63)	1.44(2)
N(1)	– C(6)	1.34(1)	C(15)	– C(16)	1.50(1)
N(3)	– C(2)	1.37(1)	C(21)	– C(22)	1.50(2)
N(3)	– C(4)	1.37(1)	C(21)	– C(26)	1.28(2)
C(4)	– C(5)	1.42(2)	C(22)	– C(23)	1.54(2)
C(5)	– C(6)	1.34(2)	C(23)	– C(24)	1.25(2)
C(6)	– C(61)	1.50(2)	C(24)	– C(25)	1.42(2)
C(11)	– C(12)	1.51(1)	C(25)	– C(26)	1.49(2)
C(11)	– C(16)	1.52(1)	C(31)	– C(32)	1.49(1)
C(12)	– C(13)	1.52(1)	C(31)	– C(36)	1.51(1)

Table 6.2.6: Bond Angles (°) For The [Cycl₃PAu(6p2-TU)] Complex.

Atom	Atom	Atom	Angle	Atom	Atom	Atom	Angle
S(2)	– Au	– P(1)	177.6(1)	C(11)	– C(12)	– C(13)	110(1)
Au	– S(2)	– C(2)	100.2(4)	C(12)	– C(13)	– C(14)	112(1)
Au	– P(1)	– C(11)	109.4(3)	C(13)	– C(14)	– C(15)	112(1)
Au	– P(1)	– C(21)	113.9(4)	C(14)	– C(15)	– C(16)	114(1)
Au	– P(1)	– C(31)	111.6(3)	C(11)	– C(16)	– C(15)	111(1)
C(11)	– P(1)	– C(21)	108.4(5)	P(1)	– C(21)	– C(22)	115.7(9)
C(11)	– P(1)	– C(31)	105.2(5)	P(1)	– C(21)	– C(26)	120(1)
C(21)	– P(1)	– C(31)	107.9(5)	C(22)	– C(21)	– C(26)	121(1)
C(2)	– N(1)	– C(6)	119(1)	C(21)	– C(22)	– C(23)	110(1)
C(2)	– N(3)	– C(4)	124(1)	C(22)	– C(23)	– C(24)	124(2)
S(2)	– C(2)	– N(1)	122.7(8)	C(23)	– C(24)	– C(25)	120(2)
S(2)	– C(2)	– N(3)	116.1(8)	C(24)	– C(25)	– C(26)	115(1)
N(1)	– C(2)	– N(3)	121(1)	C(21)	– C(26)	– C(25)	122(2)
O(4)	– C(4)	– N(3)	120(1)	P(1)	– C(31)	– C(32)	111.6(8)
O(4)	– C(4)	– C(5)	128(1)	P(1)	– C(31)	– C(36)	112.7(7)
N(3)	– C(4)	– C(5)	112(1)	C(32)	– C(31)	– C(36)	108.7(9)
C(4)	– C(5)	– C(6)	122(1)	C(31)	– C(32)	– C(33)	110(1)
N(1)	– C(6)	– C(5)	122(1)	C(32)	– C(33)	– C(34)	114(1)
N(1)	– C(6)	– C(61)	115(1)	C(33)	– C(34)	– C(35)	109(1)
C(5)	– C(6)	– C(61)	123(1)	C(34)	– C(35)	– C(36)	112(1)
P(1)	– C(11)	– C(12)	116.5(8)	C(31)	– C(36)	– C(35)	111(1)
P(1)	– C(11)	– C(16)	112.8(7)	C(6)	– C(61)	– C(62)	123(2)
C(12)	– C(11)	– C(16)	111.7(8)	C(61)	– C(62)	– C(63)	118(2)

Table 6.2.7: Mean Plane Data For The [Cycl₃PAu(6p2-TU)] Complex.

Least squares plane through the 6-n-propyl-2-thiouracilate moiety.

Atoms Defining Plane	Distance (Å)	esd (Å)
N(1)	0.0075	0.0092
N(3)	-0.0156	0.0085
C(2)	0.0059	0.0095
C(4)	0.0294	0.0130
C(5)	-0.0075	0.0155
C(6)	-0.0203	0.0149
Additional Atoms	Distance (Å)	
Au	0.1910	
S(2)	-0.0235	
P(1)	0.4893	
O(4)	0.0583	
C(61)	-0.0760	
C(62)	-0.1189	
C(63)	-0.0444	

Mean deviation from plane is 0.0144 (Å).

Chi-squared: 9.9.

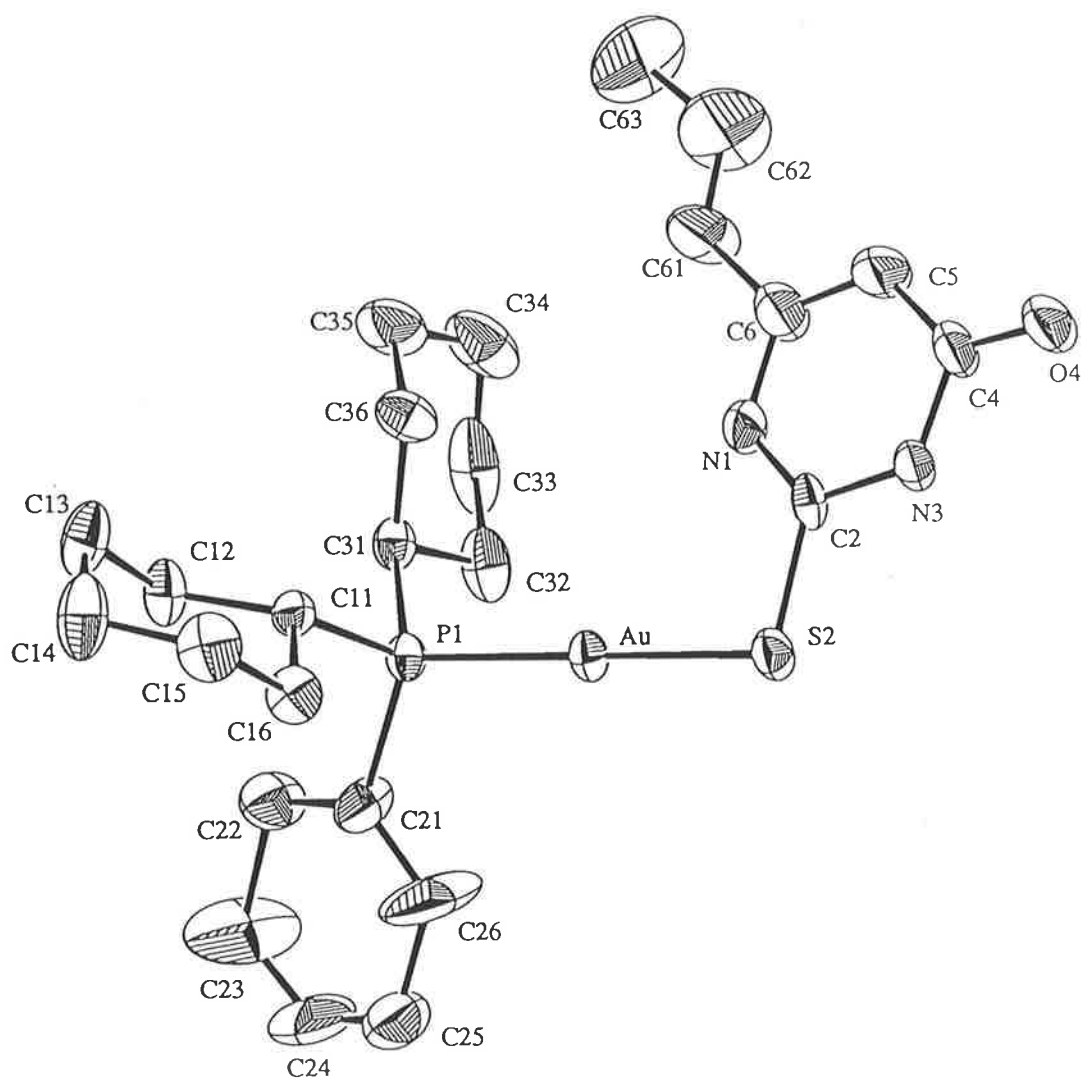


Figure 6.2.1: *Molecular Structure And Crystallographic Numbering Scheme For $[Cycl_3PAu(6p2-TU)]$.*

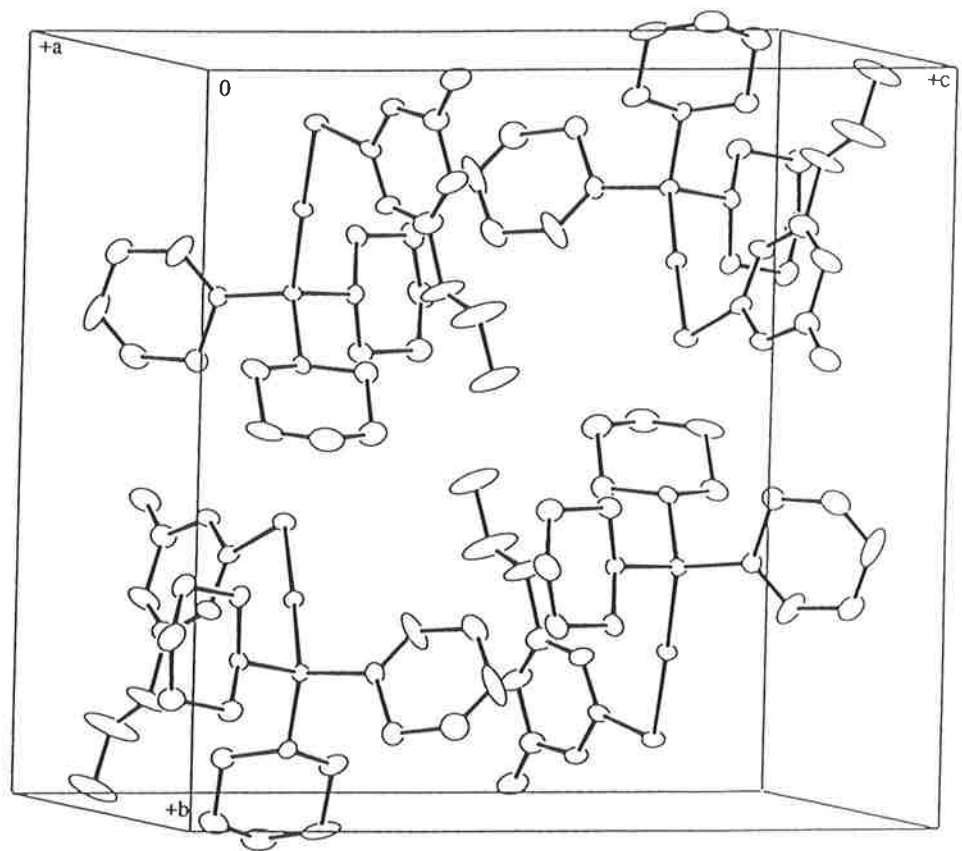


Figure 6.2.2: Unit Cell Diagram Of [Cycl₃PAu(6p₂-TU)].

bonding interaction. High thermal motion is noted for the carbon atoms of the propyl chain; this was also observed in the $[\text{Et}_3\text{PAu}(\text{6p2-TU})]^{61}$ complex.

The crystal structure of 6-n-propyl-2-thiouracil has been reported in the literature⁷⁸; however, the relatively high errors associated with the light atom positions in the complex and those in the structure of the thionucleobase itself preclude a meaningful discussion concerning the derived parameters. The structure of this complex is closely related to $[\text{Cycl}_3\text{PAu}(\text{6m2-TU})]^{77}$ (1), $[\text{Et}_3\text{PAu}(\text{6p2-TU})]^{61}$ (2), $[\text{Et}_3\text{PAu}(\text{2-TU})]^9$ (3) and $[\text{Ph}_3\text{PAu}(\text{2-TU})]^{19}$ (4). Comparisons can hence be made between these complexes for the thiolate group. Table 6.2.8 shows the intramolecular bond distances involved.

Table 6.2.8: Bond Distances (Å) For The Thiouracilate Moiety.

	$[\text{Cycl}_3\text{PAu}(\text{6p2-TU})]$	1	2	3	4
S(2)-C(2)	1.72(1)	1.73(1)	1.71(2), 1.74(2)	1.72(1)	1.722(8), 1.765(7)
N(1)-C(2)	1.28(1)	1.29(1)	1.33(2), 1.30(2)	1.31(1)	1.33(1), 1.29(1)
N(1)-C(6)	1.34(1)	1.37(1)	1.37(2), 1.38(2)	1.37(1)	1.37(1), 1.39(1)
N(3)-C(2)	1.37(1)	1.35(1)	1.37(2), 1.36(2)	1.35(1)	1.36(1), 1.34(1)
N(3)-C(4)	1.37(1)	1.40(1)	1.43(2), 1.36(2)	1.38(1)	1.37(1), 1.40(1)
C(4)-C(5)	1.42(2)	1.39(1)	1.43(2), 1.39(2)	1.40(1)	1.41(1), 1.41(5)
C(4)-O(4)	1.23(1)	1.23(1)	1.25(2), 1.23(2)	1.24(1)	1.25(1), 1.24(1)
C(5)-C(6)	1.34(2)	1.34(2)	1.33(2), 1.35(2)	1.35(2)	1.35(1), 1.34(1)

It is evident from Table 6.2.8 that most of the constituent distances are equivalent within standard deviation between the three complexes (both $[\text{Et}_3\text{PAu}(\text{6p2-TU})]$ and $[\text{Ph}_3\text{PAu}(\text{2-TU})]$ contained two molecules each in their asymmetric crystallographic units). The electron density around the ring is also identical, with C(4)-O(4) and C(5)-C(6) possessing the expected double bond character in all the structures. This suggests that the internal structure of ligands based on 2-thiouracilate is independent of the nature of the phosphine group and of the alkyl groups

external to the aromatic system. Mean plane calculations show that the six atoms of the thionucleobase in [Cycl₃PAu(6p2-TU)] form a planar system, with a mean deviation of 0.01(1) Å. The O(4), C(61) and S(2) atoms lie out of the plane by 0.0583, 0.0760 and 0.0235 Å, respectively.

6.3 Cone-angle correlation for triorganophosphinegold(I) thiolates

As discussed in Chapter 3, the cone-angle is a parameter that quantifies the steric effect of bulky phosphine groups. Triorganophosphinegold(I) thiolates of the general formula [R₃PAu(SR')], where the SR' group is a thionucleobase derivative, contain two bulky ligands, as the thiolate moiety can have significant size. The cone-angle correlation study for triorganophosphinegold(I) chlorides indicated that the P–Au bond distance in these complexes is invariant when the individual molecules crystallize as discrete molecules, i.e. with no significant intermolecular interactions such as close Au...Au contacts. However, the presence of another bulky ligand bound to the gold centre may cause a variation in the P–Au values in a series of complexes, or result in noticeable trends in the Au–S or P–Au–S chromophores.

Table 6.3.1 shows the important intramolecular parameters for a representative selection of triorganophosphinegold(I) thiolates reported in the literature. No significant Au...Au interactions were reported for any of these complexes. The cone-angles for Et₃P, Ph₃P, Cycl₃P and (*o*-Tol)₃P are 132, 145, 170 and 194 ° respectively⁴⁹, and the complexes are listed in Table 6.3.1 in order of increasing cone angle. The P–Au bond distance ranges in length from 2.248(2) to 2.255(5) Å for the Et₃P complexes, 2.237(2) to 2.260(3) Å for Ph₃P, 2.244(3) to 2.292(3) Å for Cycl₃P and 2.239(2) Å for [(*o*-Tol)₃PAu(6-MP)]. Not considering the value for [Ph₃PAu(6-MP)], the values for the other Ph₃P complexes are all equivalent within standard deviation, and this is also the case for the two Et₃P complexes. The P–Au range for the Cycl₃P complexes appears large; however, [Cycl₃PAu(2mba)] displays intermolecular hydrogen bonding, and 1-methyl-2-mercaptoimidazole is not a true thionucleobase derivative. Notable is that these two complexes possess no close Au...N interactions. Also notable is that the Au...N contacts for [Ph₃PAu(6-MP)] and [(*o*-Tol)₃PAu(6-MP)], which display the shortest P–Au

Table 6.3.1: *Intramolecular Parameters Of The P–Au–S Chromophore.*

Complex	Au–P (Å)	Au–S (Å)	P–Au–S (°)	Au...N (Å)	Ref.
[Et ₃ PAu(2-TU)]	2.248(2)	2.310(2)	176.9(1)	3.113(2)	[9]
[Et ₃ PAu(6p2-TU)]	2.249(5)	2.328(4)	175.0(2)	3.12(1)	[61]
	2.255(5)	2.314(5)	176.9(2)	3.11(1)	
[Ph ₃ PAu(2-TU)]	2.248(2)	2.296(2)	175.4(2)	3.23(1)	[19]
	2.248(2)	2.300(2)	177.0(2)	3.13(1)	
[Ph ₃ PAu(2-pymS)]	2.253(2)	2.310(3)	174.7(1)	2.951(8)	[62]
[Ph ₃ PAuL ¹]	2.256(2)	2.308(2)	178.6(2)	3.312(4)	[79]
[Ph ₃ PAu(2-pyS)]	2.258(1)	2.297(2)	177.9(1)	3.118(4)	[62]
[Ph ₃ PAuL ²]	2.258(2)	2.299(2)	176.43(8)	3.414(4)	[80]
[Ph ₃ PAu(S-C ₆ H ₅)]	2.259(2)	2.296(2)	179.12(7)	-	[81]
	2.258(2)	2.302(2)	175.79(6)	-	
[Ph ₃ PAu(S-2,4,6-C ₆ H ₂ Me ₃)]	2.255(2)	2.284(2)	175.24(7)	-	[81]
[Ph ₃ PAu(S-2,4,6-C ₆ H ₂ Et ₃)]	2.260(3)	2.288(4)	176.2(1)	-	[81]
[Ph ₃ PAu(S-2,4,6-C ₆ H ₂ iPr ₃)]	2.255(2)	2.284(2)	176.35(5)	-	[81]
[Ph ₃ PAu(6-MP)]	2.237(2)	2.287(1)	173.71(6)	2.884(5)	This work
[Cycl ₃ PAu(2mba)]	2.271(1)	2.313(1)	176.8(1)	-	[62]
[Cycl ₃ PAuL ³]	2.292(3)	2.330(3)	172.0(1)	3.641(5)	[82]
[Cycl ₃ PAu(6m2-TU)]	2.244(3)	2.299(3)	176.1(1)	3.095(8)	[77]
[Cycl ₃ PAu(6p2-TU)]	2.248(3)	2.302(3)	177.6(1)	3.061(9)	This work
[(<i>o</i> -Tol) ₃ PAu(6-MP)]	2.239(2)	2.266(2)	177.03(8)	2.860(7)	This work

Note: L¹H = 8-mercaptotheophylline, L²H = 2-mercaptobenzoxazole and L³H = 1-methyl-2-mercaptoimidazole.

distances, are significantly shorter than in the other complexes. Hence, the P–Au bond distance appears to be independent of the cone-angle for those complexes containing six-membered aromatic thionucleobase derivatives.

The Au–S bond distance shows a greater variation between the complexes, with a range extending from 2.266(2) to 2.330(3) Å. However, the variation is not related to the size of the phosphine cone-angle, nor does it appear to depend on the size of the thiolate e.g. in the series represented by the complexes of the general formula $[\text{Ph}_3\text{PAu}(\text{S}-2,4,6\text{-C}_6\text{H}_2\text{R}_3)]^{81}$ where R = H, Me, Et or *i*Pr, the 2,4,6- $\text{C}_6\text{H}_2\text{iPr}_3$ ligand would intuitively have the larger steric effects, but the C_6H_5 ligand gives rise to the longest Au–S bond in the resultant complexes. The P–Au–S bond angle displays the range of values 172.0(1) to 179.12(7)°. These values also appear to be independent of the bulkiness of the two ligands bound to the gold. The Au...N intramolecular contact varies largely in size, but maintains a relatively narrow range of 3.061(9) to 3.23(1) Å when the thiolate is a derivative of 2-thiouracil.

The conclusions from the analysis of monodentate triorganophosphinegold(I) thiolates based on thionucleobase related moieties are as follows: the linearity of the P–Au–S moiety is largely unaffected by the nature of the phosphine or thiolate bound to the gold centre; the P–Au and Au–S bond distances are independent of the cone-angle of the phosphine, but can be affected by the nature of the coordinated thiolate; and the Au...N interaction is of a secondary nature. The intramolecular parameters of the P–Au–S are thus not determined by steric factors but are more likely determined by the electronic effects of the coordinating thiolate.

CHAPTER 7

Assessment of Anti-Arthritic Activity

The effectiveness of the complexes described in this thesis against rheumatoid arthritis was determined via *in vivo* biological testing. The assessment was performed by Dr. M.W. Whitehouse and co-workers at the Department of Pathology, The University of Adelaide.

The complexes were tested on both male and female Dark Agouti rats. The choice of these rodents is based on the knowledge that Dark Agouti rats possess an immune response to rheumatoid arthritis which resembles closely that of humans, and this particular breed has been found to be sensitive to gold-based drugs. Hence, any positive results obtained may be extrapolated to give justification for later testing on humans.

Tests are continuing at the time of writing, but some interesting results have already been obtained for a selection of complexes. The protocol of testing involves injection of an arthritogen into the hind leg to induce swelling in the tail base and hind quarter joints. In a preliminary trial, the complex $[\text{Ph}_3\text{PAu}(6\text{-MP})]$, as 100 mg in saline solution, was co-administered with the arthritogen as a single dose to a female specimen. 6-mercaptopurine (6MPH), Auranofin and $[\text{Au}(\text{CN})_2]^-$ were similarly injected and assayed against an untreated specimen. The results are summarized in Table 7.1.

The untreated specimen, injected only with arthritogen, suffered weight loss and a significant increase in swelling in the hind quarters, an obviously adverse reaction. By comparison, the other complexes produced greater than 50% less swelling, indicating the complexes were

Table 7.1: Biological Data For Female Dark Agouti Rats After Anti-arthritic Treatment.

Complex	milligram dose	Δ weight (g)	Δ paw swelling (mm)
[Ph ₃ PAu(6-MP)]	100	+09	0.37
6MPH	25	+09	0.51
Auranofin	100	+03	0.71
[Au(CN) ₂] ⁻	20	+14	0.10
Untreated	-	-11	1.83

negating the effects of the arthritogen to a large extent. [Au(CN)₂]⁻ was most successful, being the species mentioned in Chapter 1 as being perhaps the main active species in the inflammation reduction process. [Ph₃PAu(6-MP)] leads to only half the swelling than for Auranofin in these trials, and to an increase in weight, so that the rat is generally healthier. Auranofin has also been observed to be toxic to specimens when given as multiple doses⁸. Therefore [Ph₃PAu(6-MP)], on the basis of these results, appears to be more successful than Auranofin in treating rheumatoid arthritis, at least in this model. 6-mercaptopurine itself also appears to be beneficial.

Three other complexes have been tested via eight regular doses after initial arthritogen injection. The results are given in Table 7.2.

Table 7.2: Biological Data For Female Dark Agouti Rats After Anti-arthritic Treatment.

Complex	milligram dose	Δ weight (g)	Δ paw swelling (mm)	incidence of disease	Rating
[Et ₃ PAu(6-MP)]	10	-09	0.45	0/3	+
[Cycl ₃ PAu(6-MP)]	10	-05	0.08	0/3	+
[(<i>m</i> -Tol) ₃ PAu(6-MP)]	10	-17	1.07	2/3	+3
Untreated	-	-19	1.69	-	+4

Incidence of disease indicates how many specimens of the three rats tested for each complex acquired the symptoms of arthritis. Thus it is evident from qualitative analysis of Table 7.2 that both [Et₃PAu(6-MP)] and [Cycl₃PAu(6-MP)] are very efficient at combatting the disease. However, the associated weight loss indicates that some side effects are present which are clearly detrimental to the health of the rats. The effectiveness of each complex was determined via a qualitative rating system⁸, indicating severity of disease, compared simultaneously with the results obtained from the standard, aurothiomalate. A number of factors were assessed, including weight, overall health of the rat and the effect on the swelling. From the same testing protocol, Auranofin receives a rating of +, so [Et₃PAu(6-MP)] and [Cycl₃PAu(6-MP)] are comparable with this complex. [(*m*-Tol)₃PAu(6-MP)] evidently has little effect on swelling and produces a significant weight loss, hence a rating of +3, only slightly more effective than no treatment at all. On the same scale, [Ph₃PAu(6-MP)] would receive a rating of + or even 0.

Some interesting observations can hence be made in terms of a tentative structure/activity relationship. 6-mercaptopurine and its complexes appear to have a significant effect on arthritic swelling; possibly the solubility characteristics of this ligand provide for a comparatively large quantity of gold to be transferred to the site of action⁸. Comparisons with triorganophosphinegold(I) thiouracilates with the same phosphine show an improvement in activity with the change in thiolate e.g. [Ph₃PAu(2-TU)] and [Et₃PAu(2-TU)] have ratings of +2 and +3 respectively^{8,9}. The triphenylphosphine ligand also appears to be the best phosphine in terms of both activity and toxicity (roughly associated with weight loss); the addition of a methyl group, as in [(*m*-Tol)₃PAu(6-MP)], gives a significant decrease in the efficiency of treatment. This difference might be related to the comparative solubilities of the phosphine ligands⁸. Generally, the effectiveness of the complexes is in the order of Ph₃P > Cycl₃P > Et₃P > (*m*-Tol)₃P, in accordance with earlier results^{9,10}.

The combination of Ph₃P and 6-MPH thus appears to possess suitable characteristics for an effective gold(I) complex which could be utilized in the treatment of rheumatoid arthritis. This is at variance with the structure of the Et₃P-containing drug Auranofin, but the effectiveness of

[Ph₃PAu(6-MP)] is promising. Further work is continuing on the other complexes, to possibly result in an improved understanding of any structure/activity relationship that exists for triorganophosphinegold(I) thiolates.

CHAPTER 8

Conclusion

The aim of the work presented in this thesis was to synthesize a range of novel triorganophosphinegold(I) complexes containing the 6-mercaptapurinate ligand and a variety of phosphines, with possible applications in the treatment of rheumatoid arthritis. Complexes of the type $[R_3PAu(6-MP)]$ (where $R_3P = Et_3P, Cycl_3P, PhMe_2P, Ph_3P, (o-Tol)_3P, (m-Tol)_3P$ or $(p-Tol)_3P$), $[(Ph_2P(CH_2)_nPPh_2)(AuCl)(Au(6-MP))]$ (where $n = 2$ or 3), and $[(Ph_2P(CH_2)_nPPh_2)(Au(6-MP))_2]$ (where $n = 1, 2$ or 3) were prepared from the corresponding triorganophosphinegold(I) chloride precursors via the metathetical reaction with base and 6-mercaptapurine in the appropriate molar quantities, as described in Chapter 2. The composition and purity of the samples was determined from microanalytical and spectroscopic evidence.

The triorganophosphinegold(I) chloride and triorganophosphinegold(I) 6-mercaptapurinate complexes were characterized using the techniques of infrared and multinuclear nuclear magnetic resonance (nmr) spectroscopy (1H and ^{13}C , and ^{31}P for the 6-MP complexes), and for the 6-mercaptapurinate complexes by Fast Atom Bombardment - mass spectroscopy (FAB-MS). Infrared data obtained for these complexes indicated complex formation via the appearance of both 6-mercaptapurine and phosphine absorption peaks in the spectra, and by alteration of the frequencies of the thioamide chromophore. Absorptions of the phosphine moieties were found to be independent of the presence of the purine moiety.

The resonances and integration found in the ^1H nmr spectra of the complexes are consistent with the proposed stoichiometries. The chemical shifts of the H^2 and H^8 proton resonances showed little variation with complexation. The ^{13}C nmr spectra revealed that the C^5 and C^8 nuclei were the most sensitive to, and thus the most indicative of, complexation. ^{31}P nmr spectra showed single signals in all the spectra, which suggested that, for the complexes with the general formula $[(\text{Ph}_2\text{P}(\text{CH}_2)_n\text{PPh}_2)(\text{AuCl})(\text{Au}(6\text{-MP}))]$ (where $n = 2$ or 3), the two independent phosphorus atoms possess the same resonance frequency. Low temperature nmr studies on one of these complexes revealed two signals, thus suggesting the ligands are fluxional at ambient temperature.

The FAB-MS studies of the complexes revealed the occurrence of high nuclearity phosphorus-gold(I)-sulphur clusters in most of the complexes, consistent with the observations for other triorganophosphinegold(I) thiolates. The molecular ion was more commonly observed for the complexes of the type $[\text{R}_3\text{PAu}(6\text{-MP})]$, but the aggregate containing two gold atoms to one 6-mercaptopurinate fragment was observed in high abundance for the complexes based on the dppm, dppe and dppp phosphines also. No otherwise significant trends were observed for the latter series. The combined spectroscopic evidence suggested product formation in all cases, with no significant differences between the complexes.

Single crystal X-ray structure determinations were performed on the complexes $[\text{PhMe}_2\text{PAuCl}]$, $[\text{Ph}_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$, $[(o\text{-Tol})_3\text{PAu}(6\text{-MP})]\cdot\text{C}_2\text{H}_5\text{OH}$ and $[\text{Cycl}_3\text{PAu}(6\text{p}2\text{-TU})]$. The studies all revealed the near linear P–Au–Cl or S chromophore expected from analogous examples found in the literature. The results from the analysis of $[\text{PhMe}_2\text{PAuCl}]$ were used in a study to compare the phosphine cone angle (θ) and a new quantity found in the literature, the ligand repulsive energy (E_R), with the P–Au and Au–Cl bond lengths for a variety of structurally characterized triorganophosphinegold(I) chloride complexes. The study revealed that the P–Au bond distance was independent of θ for complexes containing no significant gold to gold interactions in the lattice. As expected, the Au–Cl bond length was found to be unrelated to θ and probably dependent on electronic factors. The E_R quantity appeared to be no better than θ in determining any steric effects of the

phosphine on the rest of the molecule, despite its more sophisticated calculation. The crystal structure determinations on $[\text{Ph}_3\text{PAu}(6\text{-MP})].\text{C}_2\text{H}_5\text{OH}$ and $[(o\text{-Tol})_3\text{PAu}(6\text{-MP})].\text{C}_2\text{H}_5\text{OH}$ confirm the spectroscopic observations and reveal that the deprotonated 6-mercaptapurine molecule is coordinated in a monodentate mode to the gold(I) atom via the sulphur atom exclusively. The 6-mercaptapurine moiety is planar, and is oriented with the N(7) atom directed to the gold(I) atom; similar observations have been made for other triorganophosphinegold(I) thiolates. The six-membered ring system displays a general shortening of all the constituent bonds upon complexation when compared to the free ligand, an effect that is expected due to the changes in aromaticity. The observed alterations in intramolecular parameters are consistent with the spectroscopic data. The complex $[\text{Cycl}_3\text{PAu}(6\text{p}2\text{-TU})]$ possessed intramolecular characteristics consistent with analogous complexes. The phosphine and thiolate groups in triorganophosphinegold(I) thiolates have no discernable steric effects on the P–Au–S chromophore, the variations in the intramolecular parameters likely being due to electronic or packing factors.

The complexes that were assessed for their anti-arthritic activity were found to be generally effective, with activity ratings rivalling that of the clinically available drug Auranofin. The complex $[\text{Ph}_3\text{PAu}(6\text{-MP})]$ showed improved activity and less toxicity over Auranofin on a rating scheme based on the severity of rheumatoid arthritis experienced by Dark Agouti rats. The trend of Ph_3P being more active than Cycl_3P which was more active than Et_3P or $(m\text{-Tol})_3\text{P}$ for the same thiolate utilized was observed, consistent with earlier results. The 6-mercaptapurine ligand was found to be effective as the thiolate ligand, which is possibly related to its solubility in the bodily fluids.

This project also has further scope in the study of the anti-arthritic activity of these complexes. For instance, specially modified phosphines or 6-mercaptapurinates could be utilized, and the resultant gold(I) complexes investigated to determine what effects structural alterations have on the activity of these types of complexes, and possibly lead to a better understanding of the structure / activity relationships.

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APPENDIX

Structure Factor Tables

Table A1: Observed And Calculated Structure Factors For The [PhMe₂PAuCl] Complex.

10|F|o vs 10|F|c

h	k	l	Fo	Fc	sigF	h	k	l	Fo	Fc	sigF	h	k	l	Fo	Fc	sigF	h	k	l	Fo	Fc	sigF						
0	0	2	1198	1186	9	0	6	9	309	353	36	0	14	6	804	845	24	1	4	1	1274	1239	14	1	10	2	1909	1958	19
0	0	4	3571	3498	22	0	7	1	864	835	17	0	15	1	1283	1277	21	1	4	2	707	703	16	1	10	3	525	543	21
0	0	8	536	464	18	0	7	2	562	566	15	0	15	4	622	620	17	1	4	3	977	998	19	1	10	4	472	467	18
0	1	1	1824	1802	10	0	7	3	1690	1717	16	0	15	5	1063	1096	19	1	4	4	1978	1966	15	1	10	5	1442	1434	30
0	1	2	1600	1537	10	0	7	4	1344	1344	22	0	15	7	441	478	42	1	4	5	1194	1142	15	1	10	6	637	605	24
0	1	3	3772	3643	21	0	7	5	1608	1636	23	0	16	0	283	334	23	1	4	6	382	367	17	1	10	7	689	683	18
0	1	4	564	502	12	0	7	6	1302	1346	25	0	16	1	1228	1242	30	1	4	7	939	962	22	1	10	9	675	682	35
0	1	5	1328	1234	14	0	7	7	1150	1142	21	0	16	2	966	946	23	1	4	9	840	918	24	1	11	0	1083	1051	19
0	1	6	1697	1518	21	0	7	8	562	615	19	0	16	3	408	453	31	1	5	1	925	919	22	1	11	1	1411	1395	25
0	1	7	1987	1896	24	0	7	9	866	884	24	0	16	5	663	693	18	1	5	2	1204	1180	20	1	11	2	642	629	18
0	1	9	930	900	22	0	7	10	497	527	30	0	16	7	368	335	35	1	5	3	549	553	14	1	11	3	576	642	21
0	1	10	839	796	20	0	8	0	1790	1777	16	0	17	4	910	918	19	1	5	4	823	799	16	1	11	4	528	526	24
0	2	0	2059	2085	13	0	8	1	1079	1099	23	0	18	0	500	561	14	1	5	5	1343	1349	16	1	11	5	1466	1536	30
0	2	1	1465	1406	12	0	8	2	2887	2985	24	0	18	1	705	718	16	1	5	6	1959	2017	18	1	11	6	599	605	18
0	2	2	665	658	10	0	8	3	1896	1886	22	0	18	2	699	709	21	1	5	7	521	558	23	1	11	7	413	490	40
0	2	3	3659	3550	20	0	8	4	429	456	15	0	18	3	414	452	32	1	5	8	344	306	27	1	11	9	609	614	28
0	2	4	1952	1872	14	0	8	5	652	693	14	0	18	5	513	556	31	1	5	9	722	804	26	1	12	0	453	439	16
0	2	5	1009	956	17	0	8	6	1176	1172	20	0	19	1	752	750	16	1	5	10	977	996	21	1	12	1	637	627	14
0	2	6	1346	1350	19	0	8	7	798	841	19	0	20	2	796	804	20	1	6	0	640	626	13	1	12	2	2058	2025	20
0	2	9	1234	1235	26	0	8	9	459	525	40	1	0	1	1942	2226	11	1	6	1	2200	2185	21	1	12	3	784	810	16
0	3	1	779	758	15	0	9	1	1408	1376	17	1	0	2	1578	1588	10	1	6	2	2137	2191	21	1	12	4	579	533	16
0	3	2	741	756	14	0	9	2	899	931	20	1	0	3	2852	2878	16	1	6	3	1979	1978	21	1	12	5	484	498	20
0	3	3	1136	1135	14	0	9	3	1281	1321	25	1	0	4	734	738	14	1	6	4	953	966	17	1	12	6	768	787	17
0	3	4	262	249	13	0	9	4	682	661	17	1	0	6	295	307	21	1	6	6	859	856	22	1	13	0	304	267	23
0	3	5	351	342	13	0	9	5	1642	1676	19	1	0	7	1624	1545	19	1	6	7	1110	1125	22	1	13	1	1076	1041	21
0	3	6	2348	2366	17	0	9	6	1390	1339	29	1	0	8	850	788	20	1	6	8	725	682	18	1	13	2	269	247	28
0	3	9	397	450	29	0	9	7	399	386	25	1	0	9	1136	1097	29	1	6	9	471	498	27	1	13	3	414	405	21
0	3	10	1340	1422	29	0	9	8	581	623	21	1	1	1	2278	2156	16	1	7	0	607	555	14	1	13	4	1008	1034	24
0	4	0	3751	3809	21	0	9	9	733	748	28	1	1	2	2341	2240	13	1	7	1	2371	2338	18	1	13	5	750	743	18
0	4	1	1790	1772	13	0	9	10	575	600	41	1	1	3	1600	1537	12	1	7	2	961	1000	24	1	13	6	369	332	29
0	4	2	1413	1420	12	0	10	0	1400	1392	23	1	1	4	746	691	14	1	7	3	1828	1857	23	1	13	8	988	1021	22
0	4	3	1559	1561	12	0	10	1	1455	1422	24	1	1	5	851	791	14	1	7	4	1466	1487	23	1	14	1	1591	1507	20
0	4	4	1846	1873	18	0	10	2	1504	1599	24	1	1	6	2321	2204	18	1	7	5	1221	1224	19	1	14	2	257	204	30
0	4	5	355	379	18	0	10	3	1197	1264	26	1	1	7	1181	1102	20	1	7	6	1645	1639	19	1	14	3	593	591	16
0	4	6	700	689	15	0	10	4	796	820	19	1	1	10	1396	1347	30	1	7	7	607	565	23	1	14	4	1063	1096	27
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10	7	2	938	916	23	11	2	5	629	662	26	12	0	0	713	648	18	13	1	1	951	947	23	15	0	1	637	672	26
10	7	3	546	514	26	11	2	6	516	519	35	12	0	1	618	589	20	13	1	2	664	648	21	15	1	0	416	466	35
10	7	4	788	777	23	11	2	7	474	469	45	12	0	2	771	788	19	13	1	3	511	459	27	15	1	1	393	336	36
10	7	5	504	543	36	11	3	0	613	677	16	12	0	4	879	777	20	13	1	5	697	735	26	15	2	0	368	412	37
10	7	6	536	610	39	11	3	1	739	803	18	12	0	5	1007	966	23	13	2	1	755	768	18	15	2	1	671	697	24
10	7	7	525	545	45	11	3	2	767	743	18	12	0	7	536	536	37	13	2	2	854	887	21	15	3	0	351	352	38
10	8	2	1343	1312	30	11	3	3	985	1056	22	12	1	0	906	889	21	13	2	4	320	429	47	15	3	0	447	447	33
10	8	3	837	798	20	11	3	4	523	537	26	12	1	1	878	880	22	13	2	5	633	661	27	15	3	1	583	558	21
10	8	4	309	316	49	11	3	7	482	500	48	12	1	2	1132	1150	29	13	3	1	956	956	21	15	3	2	956	951	23
10	8	5	547	576	33	11	4	0	725	786	18	12	1	3	627	598	21	13	3	2	956	951	23	15	3	3	273	301	42
10	8	6	640	620	33	11	4	3	498	491	26	12	1	4	568	598	25	13	4	1	273	301	42	15	3	4	548	541	25
10	8	7	544	564	42	11	4	4	433	458	34	12	1	5	380	451	40	13	4	2	548	541	25	15	3	5	589	621	32
10	9	0	333	351	35	11	4	5	741	749																			

Table A2: Observed And Calculated Structure Factors For The $[Ph_3PAu(6-MP)] \cdot C_2H_5OH$ Complex.

10|F|o vs 10|F|c

h	k	l	Fo	Fc	sigF	h	k	l	Fo	Fc	sigF	h	k	l	Fo	Fc	sigF	h	k	l	Fo	Fc	sigF	h	k	l	Fo	Fc	sigF
0	0	1	1554	1614	8	0	3	5	295	300	7	0	6	1	830	839	7	0	9	4	440	440	9	0	16	-1	191	153	14
0	0	2	1202	1182	7	0	3	6	140	142	14	0	6	2	969	987	8	0	9	5	427	420	10	0	16	0	131	108	20
0	0	3	762	798	7	0	3	7	166	145	14	0	6	3	894	853	9	0	9	6	459	488	11	0	17	-2	138	155	20
0	0	4	1342	1331	9	0	3	8	155	101	17	0	6	4	398	406	8	0	9	7	308	302	11	0	17	2	207	202	15
0	0	5	1310	1313	10	0	3	9	146	119	21	0	6	5	268	266	8	0	9	8	262	250	13	0	17	3	144	144	20
0	0	6	627	636	10	0	4	-10	167	191	20	0	6	6	687	680	11	0	9	9	160	182	22	0	-17	-3	168	178	18
0	0	7	503	511	12	0	4	-9	239	207	13	0	6	7	379	376	10	0	10	-4	158	173	14	0	-17	-2	164	160	18
0	0	8	396	406	9	0	4	-8	255	256	11	0	6	8	322	324	10	0	10	-2	135	138	13	0	-17	-1	145	115	19
0	0	9	382	393	10	0	4	-7	494	479	13	0	7	-9	155	148	20	0	10	-1	194	202	9	0	-17	1	142	165	22
0	0	10	272	262	13	0	4	-6	595	611	11	0	7	-8	304	285	10	0	10	0	191	184	9	0	-17	3	186	151	20
0	1	-6	252	227	8	0	4	-5	677	670	9	0	7	-7	320	293	9	0	10	1	255	263	8	0	-15	-5	164	184	19
0	1	-5	98	105	15	0	4	-4	848	839	8	0	7	-6	218	222	10	0	10	2	153	144	12	0	-15	-4	285	294	12
0	1	-4	116	102	9	0	4	-3	612	626	7	0	7	-5	237	218	8	0	10	4	224	232	10	0	-15	-3	263	274	12
0	1	-3	193	189	5	0	4	-2	1409	1429	9	0	7	-4	555	559	10	0	10	5	249	254	11	0	-15	-2	175	153	15
0	1	-2	626	632	5	0	4	-1	1097	1082	7	0	7	-3	887	864	9	0	10	6	188	166	15	0	-15	0	223	229	12
0	1	-1	316	319	4	0	4	0	1023	1021	7	0	7	-2	724	722	8	0	11	-7	165	170	18	0	-15	1	309	290	10
0	1	0	329	310	4	0	4	1	659	683	6	0	7	-1	607	601	8	0	11	-6	306	317	10	0	-15	2	206	191	15
0	1	1	574	573	5	0	4	2	1029	1018	7	0	7	0	456	457	8	0	11	-5	449	447	11	0	-15	4	140	104	23
0	1	2	197	202	5	0	4	3	1104	1084	8	0	7	1	693	696	8	0	11	-4	494	492	10	0	-15	5	150	124	22
0	1	3	241	226	6	0	4	4	708	710	8	0	7	2	618	613	8	0	11	-3	436	420	9	0	-15	6	138	119	22
0	1	4	205	192	9	0	4	5	411	412	8	0	7	3	575	579	9	0	11	-2	366	357	9	0	-14	-2	130	117	18
0	1	5	189	200	12	0	4	6	356	363	8	0	7	4	520	527	10	0	11	-1	638	631	10	0	-14	1	198	182	13
0	2	-10	302	323	12	0	4	7	459	452	10	0	7	5	366	350	8	0	11	0	579	572	10	0	-14	2	147	144	20
0	2	-9	274	282	12	0	4	8	448	460	11	0	7	6	606	598	12	0	11	1	568	570	11	0	-14	3	139	81	23
0	2	-8	256	266	11	0	4	9	266	304	13	0	7	7	348	356	9	0	11	2	481	488	9	0	-13	-7	172	160	18
0	2	-7	449	450	9	0	5	-8	247	254	12	0	7	8	231	262	14	0	11	3	393	397	8	0	-13	-6	220	201	14
0	2	-6	752	773	10	0	5	-7	492	499	10	0	7	9	170	169	20	0	11	4	470	456	10	0	-13	-5	228	241	14
0	2	-5	835	839	9	0	5	-6	336	340	8	0	8	-7	173	205	16	0	11	5	409	426	10	0	-13	-4	313	324	10
0	2	-4	720	725	8	0	5	-5	195	187	9	0	8	-6	326	335	10	0	11	6	370	368	10	0	-13	-3	270	268	10
0	2	-3	799	802	7	0	5	-4	321	311	8	0	8	-5	204	205	10	0	11	7	192	205	17	0	-13	-2	212	223	11
0	2	-2	955	1000	7	0	5	-3	644	646	8	0	8	-4	339	346	8	0	11	8	208	224	17	0	-13	-1	281	291	9
0	2	-1	1399	1364	8	0	5	-2	769	753	7	0	8	-3	341	341	9	0	13	-6	201	192	16	0	-13	0	367	370	9
0	2	0	992	971	6	0	5	-1	594	591	7	0	8	-2	498	498	10	0	13	-5	254	245	12	0	-13	1	365	363	9
0	2	1	772	773	5	0	5	0	689	684	6	0	8	-1	503	494	9	0	13	-4	250	265	12	0	-13	2	191	173	15
0	2	2	1149	1140	7	0	5	1	438	432	7	0	8	0	480	460	9	0	13	-3	220	213	15	0	-13	3	233	187	13
0	2	3	1199	1224	8	0	5	2	1122	1104	8	0	8	1	607	621	9	0	13	-2	361	370	9	0	-13	4	256	229	12
0	2	4	1478	1494	10	0	5	3	566	564	8	0	8	2	418	407	10	0	13	-1	468	469	10	0	-13	5	253	217	13
0	2	5	819	811	9	0	5	4	560	550	9	0	8	3	233	234	8	0	13	0	469	469	9	0	-13	6	146	137	20
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0	2	9	366	393	10	0	5	8	227	227	14	0	8	7	209	224	14	0	13	4	427	432	10	0	-12	-3	211	197	11
0	2	10	275	275	13	0	5	9	161	165	20	0	9	-9	189	210	18	0	13	5	295	306	11	0	-12	-2	197	203	11
0	3	-7	244	226	10	0	5	10	155	167	23	0	9	-8	266	266	13	0	13	6	186	169	17	0	-12	-1	297	312	8
0	3	-6	386	392	9	0	6	-10	166	153	19	0	9	-7	181	177	15	0	13	7	199	207	17	0	-12	0	241	238	9
0	3	-5	349	342	8	0	6	-9	193	162	16	0	9	-6	207	202	12	0	15	-5	155	106	19	0	-12	1	212	212	11
0	3	-4	542	514	8	0	6	-8	268	280	11	0	9	-5	465	473	9	0	15	-3	228	222	12	0	-12	2	167	167	16
0	3	-3	403	399	7	0	6	-7	260	263	10	0	9	-4	602	594	11	0	15	-2	287	292	10	0	-12	3	219	217	13
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0	3	0	626	622	5	0	6	-4	441	435	10	0	9	-1	441	433	10	0	15	1	165	162	16	0	-11	-8	142	136	23
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0	3	4	442	445	9	0	6	0	283	311	6	0	9	3	427	415	8	0	16	-2	177	156	16	0	-11	-4	228	227	10
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0	-11	-1	493																										

1 5 0 965 976 7 1 9 0 905 890 9 1 13 0 282 296 9 2 -12 -1 472 486 10 2 -7 5 129 123 16
1 5 1 1230 1255 8 1 9 1 838 831 9 1 13 1 203 209 11 2 -12 0 472 479 10 2 -7 6 153 170 16
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1 5 7 568 592 14 1 9 7 255 267 13 1 14 -6 179 165 16 2 -11 -2 130 127 16 2 -6 -4 712 697 9
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1 7 4 603 590 10 1 11 5 364 362 9 2 -15 0 191 186 15 2 -8 -3 772 761 9 2 -5 5 321 315 7
1 7 5 792 800 11 1 11 6 209 210 15 2 -15 2 153 109 21 2 -8 -2 856 847 9 2 -5 6 361 372 9
1 7 6 592 579 12 1 11 7 202 198 17 2 -14 -5 210 234 15 2 -8 -1 628 614 9 2 -5 7 228 214 12
1 7 7 345 339 9 1 11 8 226 240 16 2 -14 -4 238 226 12 2 -8 0 499 517 10 2 -4 -10 218 219 15
1 7 8 261 270 13 1 12 -6 220 202 13 2 -14 -3 157 196 17 2 -8 1 787 780 11 2 -4 -9 267 249 11
1 7 9 212 234 17 1 12 -5 189 165 13 2 -14 -2 226 227 12 2 -8 2 801 802 10 2 -4 -8 309 318 9
1 8 -5 275 261 8 1 12 -4 155 158 15 2 -14 -1 286 286 10 2 -8 3 568 578 11 2 -4 -7 363 368 8
1 8 -4 262 245 7 1 12 -3 302 292 8 2 -14 0 335 339 9 2 -8 4 259 246 9 2 -4 -6 670 682 10
1 8 0 163 163 8 1 12 -1 399 398 10 2 -14 1 223 213 13 2 -8 5 268 260 9 2 -4 -5 707 688 9
1 8 1 413 405 10 1 12 0 213 224 11 2 -14 2 138 174 23 2 -8 6 344 348 9 2 -4 -4 941 926 8
1 8 2 275 267 7 1 12 1 210 207 11 2 -14 3 196 157 16 2 -8 7 320 321 10 2 -4 -3 617 622 7
1 8 3 207 191 9 1 12 2 165 183 13 2 -14 4 205 211 16 2 -7 -8 233 230 13 2 -4 -2 941 933 7
1 9 -9 218 238 15 1 12 3 289 309 10 2 -14 6 144 147 23 2 -7 -4 330 328 7 2 -4 -1 613 616 7
1 9 -8 179 192 17 1 12 4 278 286 10 2 -13 -1 130 146 19 2 -7 -3 479 466 9 2 -4 0 326 322 6
1 9 -7 254 240 11 1 12 5 240 252 13 2 -13 3 163 135 18 2 -7 -2 125 87 10 2 -4 1 1069 1035 8
1 9 -6 401 398 10 1 13 -7 133 129 22 2 -12 -7 225 227 14 2 -7 -1 350 344 7 2 -4 2 643 647 7
1 9 -5 567 570 11 1 13 -5 172 156 14 2 -12 -6 334 342 10 2 -7 0 158 144 9 2 -4 3 516 513 8
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2 -4 7 155 136 16 2 -1 4 923 921 9 2 3 -1 483 479 7 2 6 4 445 412 11 2 10 -7 215 201 13
2 -4 9 173 200 19 2 -1 5 809 804 10 2 3 0 967 966 7 2 6 5 288 275 8 2 10 -6 334 331 8
2 -3 -10 265 286 13 2 -1 6 467 471 9 2 3 1 1298 1313 8 2 6 6 318 321 9 2 10 -5 376 371 8
2 -3 -9 328 324 10 2 -1 7 452 460 10 2 3 2 1907 1868 11 2 6 7 144 155 20 2 10 -4 344 330 9
2 -3 -8 312 330 9 2 -1 8 443 454 10 2 3 3 1105 1096 9 2 6 8 221 224 15 2 10 -3 386 360 8
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2 -2 -7 253 264 9 2 1 -5 884 873 8 2 5 -10 203 196 16 2 8 -8 203 211 14 2 11 3 370 367 9
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2 -2 -2 654 650 6 2 1 0 1008 979 7 2 5 -5 479 497 10 2 8 -3 467 447 12 2 12 -4 253 236 11
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2 -2 1 236 248 5 2 1 3 1243 1230 9 2 5 -2 1110 1095 9 2 8 0 778 770 9 2 12 -1 400 398 10
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8	-3	3	215	226	12	8	1	0	489	498	12	8	5	-7	129	139	17	8	10	2	216	249	14	8	-6	-2	164	137	13
8	-3	4	294	309	11	8	1	1	436	435	9	8	5	-6	196	186	11	8	10	3	262	259	12	8	-6	-1	132	129	17
8	-3	5	329	316	11	8	1	2	420	424	10	8	5	-5	296	305	8	8	11	-6	211	206	13	8	-6	0	152	136	15
8	-3	6	240	262	14	8	1	3	313	300	9	8	5	-4	256	267	9	8	11	-5	193	177	13	8	-6	1	136	125	18
8	-2	-6	130	142	15	8	1	4	303	298	10	8	5	-3	188	185	11	8	11	-1	213	206	13	8	-6	2	125	118	20
8	-2	-5	242	243	9	8	1	5	186	208	17	8	5	-2	251	243	9	8	11	0	144	133	19	8	-5	-9	211	183	14
8	-2	-4	233	235	8	8	2	-10	152	137	18	8	5	-1	145	152	14	8	11	3	147	131	20	8	-5	-8	194	188	14
8	-2	-1	171	165	11	8	2	-8	201	199	12	8	5	0	149	152	14	8	12	-7	166	171	17	8	-5	-7	211	214	12
8	-2	0	241	245	9	8	2	-7	304	294	8	8	5	1	232	221	10	8	12	-6	195	219	15	8	-5	-6	226	247	11
8	-2	1	203	190	11	8	2	-6	359	374	8	8	5	2	221	214	12	8	12	-5	178	171	15	8	-5	-5	312	310	8
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8	-2	3	134	102	18	8	2	-4	265	259	8	8	6	-10	156	139	17	8	12	-3	220	226	13	8	-5	-3	329	332	8
8	-1	-11	139	161	22	8	2	-3	425	415	9	8	6	-9	238	234	11	8	12	-2	312	303	9	8	-5	-2	229	215	10
8	-1	-10	194	227	16	8	2	-2	615	620	11	8	6	-8	399	375	9	8	12	-1	313	310	9	8	-5	-1	291	278	9
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8	-1	-8	270	259	9	8	2	0	325	329	8	8	6	-6	337	342	7	8	12	1	183	182	16	8	-5	1	322	333	9
8	-1	-7	211	210	11	8	2	1	241	236	9	8	6	-5	484	489	9	8	12	2	205	209	15	8	-5	2	225	248	12
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8	-1	-4	495	508	11	8	2	4	249	255	13	8	6	-2	470	487	9	8	13	-3	161	152	16	8	-4	-10	182	121	15
8	-1	-3	355	362	9	8	2	5	142	134	22	8	6	-1	325	334	8	8	13	-2	274	264	11	8	-4	-6	159	167	14
8	-1	-2	463	472	12	8	2	6	190	169	18	8	6	0	483	480	9	8	13	-1	215	222	13	8	-4	-5	351	354	9
9	-4	-4	300	292	8	9	0	-7	422	421	10	9	4	2	425	418	9	9	9	-2	222	241	12	10	-6	2	140	139	19
9	-4	-3	202	207	11	9	0	-6	500	501	12	9	4	3	288	309	11	9	9	-1	312	302	9	9	-6	3	142	115	19
9	-4	-2	118	113	18	9	0	-5	463	467	9	9	4	4	238	223	13	9	9	0	343	356	9	10	-5	-7	133	120	18
9	-4	-1	216	203	11	9	0	-4	319	328	8	9	4	5	214	245	17	9	9	1	150	184	19	10	-4	-10	161	164	18
9	-4	0	262	254	9	9	0	-3	431	433	9	9	5	-5	150	150	14	9	9	2	149	164	19	10	-4	-9	160	150	18
9	-4	1	184	178	13	9	0																						

10	2	2	364	365	9	10	7	-5	407	413	10	11	-4	-2	240	241	11	11	1	1	292	286	10	11	9	-1	231	211	12
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10	3	-9	166	166	16	10	7	-2	246	249	10	11	-4	1	178	137	14	11	2	-7	118	87	19	11	10	-6	142	119	16
10	3	-8	239	246	11	10	7	-1	361	361	8	11	-4	2	166	146	17	11	2	-6	132	112	16	11	10	-3	218	203	12
10	3	-7	186	166	12	10	7	0	420	415	9	11	-4	3	165	164	18	11	2	-5	139	122	15	11	10	-2	167	203	16
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10	3	0	271	265	9	10	9	-7	242	242	11	11	-2	-9	136	128	19	11	3	-5	357	358	9	12	-6	-2	174	120	14
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10	4	-8	216	206	11	10	9	-3	261	264	10	11	-2	-5	204	229	12	11	3	-1	322	312	8	12	-5	-2	172	164	15
10	4	-7	256	264	10	10	9	-2	373	375	9	11	-2	-4	166	169	13	11	3	0	375	364	8	12	-5	-1	165	162	16
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10	4	-4	279	264	9	10	9	2	259	250	13	11	-2	-1	215	204	11	11	3	3	141	151	21	12	-3	-8	176	154	15
10	4	-3	295	296	9	10	11	-7	212	205	12	11	-2	0	182	179	13	11	4	-3	124	98	17	12	-3	-7	142	174	11
10	4	-2	236	234	10	10	11	-6	184	197	14	11	-2	2	184	190	16	11	5	-9	203	192	13	12	-3	-6	215	194	11
10	4	-1	244	244	10	10	11	-5	160	182	15	11	-1	-8	206	179	12	11	5	-8	183	191	13	12	-3	-5	162	148	14
10	4	0	288	285	9	10	11	-4	201	204	13	11	-1	-7	204	199	12	11	5	-7	150	175	16	12	-3	-4	192	177	13
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10	4	3	165	148	19	10	11	-1	277	276	11	11	-1	-4	189	176	11	11	5	-4	297	299	9	12	-3	-1	192	179	13
10	5	-10	144	148	18	10	11	0	207	210	14	11	-1	-3	297	283	8	11	5	-3	245	254	10	12	-3	0	149	129	17
10	5	-9	225	223	12	10	12	-3	134	96	18	11	-1	-2	273	273	9	11	5	-2	326	333	8	12	-1	-9	162	156	16
10	5	-8	226	210	11	10	13	-3	203	211	13	11	-1	-1	181	200	13	11	5	-1	368	376	9	12	-1	-8	194	184	13
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10	5	-6	242	250	10	11	-10	-3	152	126	17	11	-1	1	164	148	15	11	5	1	297	289	10	12	-1	-6	184	166	12
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10	5	-2	265	281	9	11	-8	0	170	156	16	11	0	-6	137	130	16	11	7	-7	217	220	12	12	-1	-2	239	234	10
10	5	-1	356	368	8	11	-8	1	137	145	20	11	0	-5	195	193	12	11	7	-6	276	270	10	12	-1	-1	174	169	14
10	5	0	452	475	11	11	-6	-8	142	116	19	11	0	-4	217	193	10	11	7	-5	282	277	9	12	-1	0	168	165	15
10	5	1	316	319	10	11	-6	-7	197	176	14	11	0	-3	195	200	12	11	7	-4	228	232	11	12	-1	1	215	212	14
10	5	2	243	214	12	11	-6	-6	214	197	12	11	0	-2	172	195	13	11	7	-3	278	266	9	12	-1	2	210	213	14
12	1	-9	139	162	19	12	8	-1	137	160	19	13	6	-3	247	222	10	13	6	-2	202	230	13	13	6	-2	202	230	13
12	1	-8	190	173	13	12	9	-2	122	61	19	13	6	-2	202	230	13	13	6	-1	196	193	13	13	6	-1	196	193	13
12	1	-7	156	176	16	12	10	-4	177	215	14	13	6	-1	196	193	13	13	6	-1	196	193	13	13	6	-1	196	193	13
12	1	-6	242	245	11	12	10	-3	240	252	11	13	8	-4	189	190	13	13	8	-4	189	190	13	13	8	-4	189	190	13
12	1	-5	254	239	10	13	7	-6	130	126	20	14	-3	-6	122	55	18	14	-3	-6	122	55	18	14	-3	-6	122	55	18
12	1	-4	331	337	8	13	-5	-7	151	134	16	14	-2	-3	135	109	17	14	-2	-3	135	109	17	14	-2	-3	135	109	17
12	1	-3	280	264	9	13	-5	-6	127	129	19	14	0	-6	159	151	15	14	0	-6	159	151	15	14	0	-6	159	151	15
12	1	-2	173	187	14	13	-5	-5	132	106	18	14	0	-5	171	164	14	14	0	-5	171	164	14	14	0	-5	171	164	14
12	1	-1	198	196	12	13	-5	-2	171	150	15	14	0	-2	158	140	15	14	0	-2	158	140	15	14	0	-2	158	140	15
12	1	0	248	241	11	13	-3	-8	142	119	18	14	2	-7	123	111	18	14	2	-7	123	111	18	14	2	-7	123	111	18
12	1	1	249	259	12	13	-3	-6	165	110	14	14	2	-6	155	144	15	14	2	-6	155	144	15	14	2	-6	155	144	15
12	1	2	169	155	16	13	-3	-4	134	157	19	14	2	-5	189	163	12	14	2	-5	189	163	12	14	2	-5	189	163	12
12	2	-4	132	132	17	13	-3	-3	195	195	13	14	2	-4	150	130	15	14	2	-4	150	130	15	14	2	-4	150	130	15
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12	3	-9	152	144	16	13	-3	-1	131	124	18	14	4	-6	156	139	14	14	4	-6	156	139	14	14	4	-6	156	139	14
12	3	-8	148	140	17	13	-3	0	132	115	19	14	4	-5	140	139	16	14	4	-5	140	139	16	14	4	-5	140	139	16
12	3	-7	184	181	13	13	-2	-8	146	102	17	14	4	-4	155	145	14	14	4	-4	155	145	14	14	4</				

Table A3: Observed And Calculated Structure Factors For The [(o-Tol)₃PAu(6-MP)].C₂H₅OH Complex.

10|F_o vs 10|F_c

h k l Fo Fc sigF					h k l Fo Fc sigF					h k l Fo Fc sigF					h k l Fo Fc sigF														
0	0	4	2676	2460	10	0	2	15	408	384	10	0	4	21	266	245	11	0	6	24	271	291	14	0	9	8	727	715	8
0	0	6	2661	2520	10	0	2	16	902	909	8	0	4	24	613	601	14	0	6	25	490	467	13	0	9	11	282	286	8
0	0	8	358	331	6	0	2	17	367	368	9	0	4	25	308	269	13	0	6	28	228	234	22	0	9	12	663	633	9
0	0	10	1786	1822	8	0	2	18	606	616	10	0	4	28	342	353	15	0	6	29	200	221	27	0	9	13	181	190	12
0	0	12	640	650	6	0	2	20	1133	1135	10	0	4	29	212	166	25	0	7	1	617	662	6	0	9	14	259	233	9
0	0	14	1078	1119	7	0	2	22	212	203	14	0	4	30	279	259	21	0	7	2	1195	1206	7	0	9	15	163	184	15
0	0	16	804	843	8	0	2	24	758	761	13	0	5	1	548	604	5	0	7	3	826	841	6	0	9	16	373	383	11
0	0	18	855	889	8	0	2	26	256	267	15	0	5	2	502	558	5	0	7	4	872	875	6	0	9	17	398	392	10
0	0	20	941	1019	9	0	2	28	283	306	16	0	5	3	986	1019	6	0	7	5	171	163	8	0	9	18	493	500	14
0	0	24	975	1030	10	0	2	30	459	439	14	0	5	4	348	317	6	0	7	6	500	501	7	0	9	21	418	397	12
0	0	26	523	561	15	0	3	1	1236	1283	5	0	5	5	350	325	7	0	7	7	864	874	7	0	9	22	431	417	13
0	0	28	440	440	13	0	3	2	271	282	5	0	5	6	698	665	6	0	7	8	1001	994	7	0	10	0	163	169	10
0	0	30	392	433	12	0	3	3	1813	1881	7	0	5	7	1747	1721	8	0	7	11	584	545	8	0	10	1	559	546	8
0	1	1	1403	1560	6	0	3	4	242	232	6	0	5	8	1139	1085	6	0	7	12	625	577	8	0	10	2	152	140	11
0	1	2	428	449	3	0	3	5	285	269	6	0	5	9	387	365	9	0	7	13	607	573	9	0	10	3	298	304	9
0	1	3	3333	3158	13	0	3	6	705	683	5	0	5	10	360	329	9	0	7	14	315	297	10	0	10	4	350	347	11
0	1	4	1681	1719	7	0	3	7	2043	2021	8	0	5	11	1348	1310	7	0	7	15	324	293	10	0	10	5	767	731	8
0	1	5	1645	1635	7	0	3	8	952	989	6	0	5	12	897	886	7	0	7	16	334	330	10	0	10	6	354	352	12
0	1	6	443	408	5	0	3	9	1721	1716	8	0	5	13	1448	1389	8	0	7	17	466	443	12	0	10	7	159	118	11
0	1	7	2019	2007	8	0	3	10	864	862	6	0	5	14	604	594	8	0	7	18	482	473	13	0	10	8	138	134	14
0	1	8	399	373	6	0	3	11	1641	1680	8	0	5	15	162	168	12	0	7	20	265	227	11	0	10	9	625	609	9
0	1	9	740	742	6	0	3	12	725	698	7	0	5	16	862	825	8	0	7	21	392	376	12	0	10	10	199	178	10
0	1	10	795	754	6	0	3	13	1177	1169	7	0	5	17	899	853	9	0	7	22	516	510	15	0	10	11	484	485	11
0	1	11	922	944	7	0	3	14	387	340	10	0	5	18	250	211	10	0	7	23	434	422	13	0	10	12	115	13	18
0	1	12	483	453	8	0	3	15	370	344	10	0	5	19	329	331	9	0	7	25	284	249	14	0	10	13	181	172	13
0	1	13	1395	1437	8	0	3	16	296	276	10	0	5	21	512	496	13	0	7	26	265	243	16	0	10	14	239	240	10
0	1	14	270	326	8	0	3	17	1615	1570	9	0	5	22	461	417	11	0	7	27	287	292	17	0	10	15	577	552	11
0	1	15	400	405	9	0	3	18	570	588	11	0	5	23	328	317	11	0	7	28	343	326	15	0	10	17	149	134	18
0	1	17	1320	1344	8	0	3	19	521	521	12	0	5	24	196	181	18	0	8	0	856	879	7	0	10	18	326	311	9
0	1	18	379	376	9	0	3	21	745	722	12	0	5	25	228	194	17	0	8	1	1015	1011	7	0	10	19	403	407	12
0	1	19	415	433	12	0	3	22	389	435	13	0	5	26	276	263	15	0	8	2	130	112	10	0	10	20	230	186	14
0	1	21	891	944	10	0	3	23	518	515	15	0	5	27	314	302	15	0	8	3	266	280	8	0	10	21	393	347	13
0	1	23	857	920	11	0	3	27	567	547	19	0	6	0	1118	1156	6	0	8	4	819	931	7	0	11	1	198	182	9
0	1	24	213	215	14	0	3	31	341	355	18	0	6	1	1017	1050	6	0	8	5	1049	1043	7	0	11	2	500	498	9
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9 3 14 210 196 13 9 6 -12 414 442 13 10 0 8 576 587 12 10 3 -1 239 219 9 10 6 -5 333 348 10
9 4 -21 492 493 13 9 6 -11 399 425 13 10 0 10 588 546 12 10 3 1 572 617 10 10 6 -4 244 257 9
9 4 -19 182 201 19 9 6 -10 293 293 9 10 0 12 184 175 17 10 3 2 240 231 9 10 6 -2 212 224 10
9 4 -18 199 167 16 9 6 -7 431 420 11 10 1 -19 464 439 12 10 3 3 450 465 11 10 6 -1 332 355 10
9 4 -17 344 356 11 9 6 -6 490 496 10 10 1 -17 735 661 14 10 3 4 207 202 10 10 6 0 403 420 13
9 4 -16 232 240 13 9 6 -5 288 280 9 10 1 -13 926 833 11 10 3 5 177 190 12 10 6 3 288 307 8
9 4 -15 458 482 14 9 6 -4 160 171 13 10 1 -11 335 283 10 10 3 7 346 355 9 10 6 4 374 357 11
9 4 -12 306 292 9 9 6 -3 323 332 10 10 1 -9 521 518 12 10 3 11 374 359 11 10 6 5 286 296 10
9 4 -11 564 558 11 9 6 -2 380 403 12 10 1 -7 654 650 10 10 4 -16 252 239 13 10 7 -7 248 253 9
9 4 -10 224 225 10 9 6 -1 418 436 11 10 1 -5 136 98 15 10 4 -15 211 187 14 10 7 -4 227 232 9
9 4 -7 656 638 9 9 6 0 224 247 9 10 1 -3 672 682 9 10 4 -14 310 303 10 10 7 -3 254 270 9
9 4 -6 386 418 13 9 6 2 215 209 10 10 1 -1 350 372 9 10 4 -12 216 212 13 10 7 -2 230 248 9
9 4 -5 388 364 12 9 6 3 414 427 13 10 1 0 154 139 13 10 4 -11 201 209 13 11 0 -15 506 453 12
9 4 -3 443 451 11 9 6 4 294 295 9 10 1 1 478 498 11 10 4 -10 559 566 12 11 0 -13 438 418 12
9 4 -2 351 357 9 9 6 5 327 307 10 10 1 3 535 548 10 10 4 -9 184 203 14 11 0 -9 554 508 13
9 4 -1 888 915 8 9 6 6 7 241 246 11 10 1 5 337 385 10 10 4 -7 219 183 11 11 0 -5 489 472 12
9 4 0 280 293 9 9 6 8 324 307 10 10 1 6 139 131 15 10 4 -6 776 746 9 11 0 -3 415 416 14
9 4 2 218 219 9 9 6 9 438 411 15 10 1 7 762 779 10 10 4 -5 295 303 8 11 0 -1 485 464 12
9 4 3 860 894 9 9 7 -14 388 409 10 10 1 8 139 91 18 10 4 -4 334 335 10 11 0 1 623 596 10
9 4 4 315 313 9 9 7 -13 281 298 9 10 1 11 601 564 12 10 4 -3 127 101 16 11 0 5 700 730 11
9 4 5 326 307 11 9 7 -12 212 239 11 10 1 13 358 331 10 10 4 -2 421 448 12 11 0 7 340 345 10
9 4 7 381 378 10 9 7 -10 278 273 8 10 2 -20 565 530 13 10 4 -1 314 346 10 11 -16 472 434 13
9 4 8 291 286 9 9 7 -9 455 495 11 10 2 -18 214 186 18 10 4 0 574 582 10 11 -12 494 485 15
9 4 9 370 355 11 9 7 -8 301 322 9 10 2 -16 572 535 15 10 4 1 147 114 13 11 -10 263 245 11
9 4 10 169 153 15 9 7 -7 255 282 9 10 2 -14 472 402 16 10 4 3 297 316 9 11 -8 323 287 10
9 4 13 357 347 12 9 7 -6 168 185 12 10 2 -13 143 123 21 10 4 4 405 431 12 11 -7 176 155 14
9 5 -19 258 250 13 9 7 -5 277 300 8 10 2 -12 402 388 11 10 4 5 201 214 11 11 -6 403 396 10
9 5 -18 390 387 12 9 7 -4 323 347 10 10 2 -11 146 131 19 10 4 6 176 187 14 11 -5 183 173 13
9 5 -15 250 264 11 9 7 -3 343 390 9 10 2 -10 796 752 10 10 4 8 224 222 12 11 -4 243 249 10
9 5 -14 491 514 13 9 7 0 230 226 9 10 2 -8 159 174 15 10 4 9 158 152 16 11 -2 601 602 10
9 5 -13 270 270 10 9 7 1 256 277 9 10 2 -7 215 229 11 10 4 10 287 266 10 11 1 2 621 631 10
11 1 4 463 459 11 11 3 -2 547 554 11 12 1 -6 178 157 16
11 1 6 296 296 9 11 3 -1 234 253 10 12 1 -5 432 434 10
11 1 7 134 106 17 11 3 0 135 142 15 12 1 -3 237 231 11
11 1 8 556 533 12 11 3 2 356 364 9 12 1 -1 592 555 11
11 2 -15 480 425 12 11 3 3 197 203 11
11 2 -13 429 398 12 11 3 4 267 271 8
11 2 -11 251 250 13 11 4 -11 327 299 11
11 2 -9 515 497 13 11 4 -10 355 351 10
11 2 -8 127 132 21 11 4 -9 585 579 12
11 2 -5 615 585 10 11 4 -6 215 195 12
11 2 -4 299 313 10 11 4 -5 602 602 11
11 2 -3 293 311 8 11 4 -4 223 244 11
11 2 -2 162 163 14 11 4 -3 280 267 9
11 2 -1 384 384 12 11 4 -1 326 343 11
11 2 0 150 137 14 11 4 0 204 226 11
11 2 1 530 521 10 11 4 1 350 371 9
11 2 2 180 190 12 11 5 -8 484 439 12
11 2 5 561 601 11 11 5 -7 275 289 10
11 2 7 223 229 10 11 5 -6 474 470 12
11 3 -13 262 242 13 11 5 -4 153 129 15
11 3 -12 636 595 13 11 5 -3 265 266 9
11 3 -10 267 233 11 11 5 -2 423 430 12
11 3 -9 194 155 14 11 5 0 130 115 16
11 3 -8 531 513 13 12 0 -8 382 369 10
11 3 -7 240 247 11 12 0 -4 617 601 12
11 3 -6 486 471 12 12 0 -2 232 201 12
11 3 -4 225 222 10 12 1 -7 481 459 14

Table A4: Observed And Calculated Structure Factors For The [Cycl3PAu(6p2-TU)] Complex.

10 Fo vs 10 F c																			page 1					
h	k	l	Fo	Fc	sigF	h	k	l	Fo	Fc	sigF	h	k	l	Fo	Fc	sigF	h	k	l	Fo	Fc	sigF	
0	0	2	3218	3272	8	0	2	22	236	235	28	0	5	12	561	589	13	0	8	3	1958	1986	11	
0	0	4	3475	3505	9	0	3	1	1415	1401	7	0	5	13	1013	1010	17	0	8	4	377	379	11	
0	0	6	2181	2248	10	0	3	2	652	643	10	0	5	14	375	358	15	0	8	5	1142	1176	13	
0	0	8	870	832	13	0	3	3	1118	1097	9	0	5	15	568	560	14	0	8	6	687	702	15	
0	0	10	286	297	15	0	3	4	606	576	12	0	5	16	189	231	29	0	8	7	520	516	11	
0	0	12	136	40	35*	0	3	5	820	822	11	0	5	17	322	333	20	0	8	8	287	310	17	
0	0	14	414	446	15	0	3	6	1353	1344	11	0	5	18	114	104	65*	0	8	9	254	221	19	
0	0	16	465	462	14	0	3	7	1225	1233	11	0	5	19	86	137	78*	0	8	10	282	287	18	
0	0	18	481	491	16	0	3	8	1690	1654	12	0	5	20	0	3	77*	0	8	11	132	86	71*	
0	0	20	466	451	16	0	3	9	960	1036	14	0	5	21	146	21	76*	0	8	12	139	109	36*	
0	0	22	277	328	27	0	3	10	1277	1248	14	0	6	0	1219	1230	10	0	8	13	228	263	24	
0	1	1	136	729	6	0	3	11	1002	1023	16	0	6	1	1990	2045	10	0	8	14	97	93	55*	
0	1	2	1009	991	7	0	3	12	970	991	16	0	6	2	234	235	12	0	8	15	280	274	20	
0	1	3	494	491	10	0	3	13	682	682	14	0	6	3	2034	2030	10	0	8	16	128	139	75*	
0	1	4	2361	2307	9	0	3	14	662	626	15	0	6	4	628	599	13	0	8	17	289	362	22	
0	1	5	642	596	12	0	3	15	509	499	13	0	6	5	1468	1451	11	0	8	18	91	126	77*	
0	1	6	2242	2251	10	0	3	16	225	234	25	0	6	6	0	4	458*	0	8	19	318	371	21	
0	1	7	977	934	12	0	3	17	342	325	18	0	6	7	989	979	13	0	8	20	115	90	75*	
0	1	8	1436	1456	12	0	3	18	15	118	78*	0	6	8	165	152	24	0	9	1	682	613	15	
0	1	9	215	215	19	0	3	19	0	151	79*	0	6	9	560	525	14	0	9	2	197	217	19	
0	1	10	1296	1311	14	0	3	20	0	9	79*	0	6	10	82	106	67*	0	9	3	636	656	16	
0	1	11	199	126	23	0	3	21	138	33	80*	0	6	11	0	47	70*	0	9	4	487	472	11	
0	1	12	1022	989	16	0	4	0	1147	1211	8	0	6	12	35	145	71*	0	9	5	1297	1310	13	
0	1	13	93	192	74*	0	4	1	1727	1698	8	0	6	13	261	321	22	0	9	6	339	362	14	
0	1	14	736	721	14	0	4	2	1624	1597	9	0	6	14	102	117	61*	0	9	7	966	1002	15	
0	1	15	113	218	77*	0	4	3	1333	1356	9	0	6	15	366	407	17	0	9	8	218	229	22	
0	1	16	481	470	14	0	4	4	1275	1238	10	0	6	16	0	61	74*	0	9	9	819	853	17	
0	1	17	0	158	77*	0	4	5	1518	1498	10	0	6	17	271	336	22	0	9	10	331	315	15	
0	1	18	199	207	33	0	4	6	1068	1028	11	0	6	18	0	4	477*	0	9	11	603	623	15	
0	1	19	0	82	77*	0	4	7	923	883	13	0	6	19	235	249	26	0	9	12	323	332	17	
0	1	20	199	61	33	0	4	8	667	677	15	0	6	20	106	39	63*	0	9	13	582	584	13	
0	1	21	132	12	81*	0	4	9	246	242	17	0	6	21	26	204	80*	0	9	14	217	228	26	
0	1	22	87	12	82*	0	4	10	125	42	65*	0	7	1	1001	958	11	0	9	15	388	396	16	
0	2	0	3450	3449	8	0	4	11	61	118	69*	0	7	2	805	778	12	0	9	16	238	162	23	
0	2	1	381	368	9	0	4	12	207	227	24	0	7	3	687	740	13	0	9	17	234	211	24	
0	2	2	2255	2229	8	0	4	13	180	159	30	0	7	4	553	573	11	0	9	18	0	90	75*	
0	2	3	168	134	14	0	4	14	494	511	13	0	7	5	1288	1301	12	0	9	19	110	81	58*	
0	2	4	2870	2925	9	0	4	15	394	377	15	0	7	6	356	378	12	0	9	20	143	80	74*	
0	2	5	277	293	11	0	4	16	325	329	18	0	7	7	1651	1633	13	0	10	0	1545	1535	13	
0	2	6	1549	1498	10	0	4	17	211	237	28	0	7	8	94	85	48*	0	10	1	2139	2159	12	
0	2	7	404	364	9	0	4	18	117	273	80*	0	7	9	1371	1357	14	0	10	2	1336	1347	13	
0	2	8	720	683	14	0	4	19	0	166	78*	0	7	10	102	191	52*	0	10	3	1387	1398	13	
0	2	9	0	89	62*	0	4	20	0	205	80*	0	7	11	1119	1116	16	0	10	4	1014	967	14	
0	2	10	93	68	44*	0	4	21	0	138	79*	0	7	12	124	72	71*	0	10	5	1025	1056	15	
0	2	11	0	83	67*	0	5	1	98	121	26*	0	7	13	914	862	17	0	10	6	498	493	11	
0	2	12	284	262	18	0	5	2	92	79	51*	0	7	14	0	67	72*	0	10	7	732	724	18	
0	2	13	64	24	73*	0	5	3	866	861	11	0	7	15	520	428	40*	0	10	8	242	224	21	
0	2	14	515	510	12	0	5	4	796	800	12	0	7	16	138	180	10	0	10	9	183	174	28	
0	2	15	0	54	76*	0	5	5	1017	1004	11	0	7	17	346	325	18	0	10	10	261	232	20	
0	2	16	454	467	14	0	5	6	847	874	12	0	7	18	0	109	77*	0	10	11	0	64	69*	
0	2	17	182	154	32*	0	5	7	1660	1658	12	0	7	19	145	129	76*	0	10	12	0	96	73*	
0	2	18	443	452	15	0	5	8	655	667	11	0	7	20	179	25	32*	0	10	13	196	260	28	
0	2	19	132	128	48*	0	5	9	1653	1635	13	0	8	0	434	418	10	0	10	14	83	28	72*	
0	2	20	311	368	23	0	5	10	1006	992	15	0	8	1	1955	1952	11	0	10	15	327	338	17	
0	2	21	0	98	80*	0	5	11	1525	1485	14	0	8	2	696	686	14	0	10	16	133	145	76*	
0	2	22	0	12	74*	0	5	12	548	566	15	1	0	-2	4346	4266	8	1	2	-22	291	224	30	
0	2	23	16	274	251	21	0	17	7	105	228	83*	1	0	2	1621	1622	5	1	2	-21	218	44	35
0	2	24	17	43	55	76*	0	17	8	474	493	16	1	0	4	576	587	9	1	2	-20	323	314	25
0	2	25	0	908	862	18	0	17	9	244	200	28	1	0	6	665	665	10	1	2	-19	160	18	85*
0	2	26	1	322	333	18	0	17	10	381	387	19	1	0	8	332	312	10	1	2	-18	237	239	30
0	2	27	2	952	920	18	0	17	11	245	220	28	1	0	10	212	267	18	1	2	-17	0	34	80*
0	2	28	3	142	135	75*	0	17	12	272	351	25	1	0	12	702	704	17	1	2	-16	181	186	34*
0	2	29	4	733	702	15	0	17	13	118	149	79*	1	0	14	848	899	18	1	2	-15	221	155	27
0	2	30	5	172	154	32*	0	18	0	562	553	14	1	0	16	759	759	14	1	2	-14	0	56	75*
0	2	31	6	801	742	20	0	18	1	712	659	17	1	0	18	675	626	16	1	2	-13	126	73	72*
0	2	32	7	53	79	76*	0	18	2	481	467	16	1	0	20	468	430	16	1	2	-12	533	502	13
0	2	33	8	315	328	20	0	18	3	531	491	15	1	1	-22	205	141	42*	1	2	-11	230	145	19
0	2	34	9	0	27	76*	0	18	4	330	338	22	1	1	-21	216	90	37*	1	2	-10	900	838	14
0	2	35	10	102	82	77*	0	18	5	455	418	17	1	1	-20	344	222	23	1	2	-9	141	119	24*
0	2	36	11	189	110	29	0	18	6	133	166	70*	1	1	-19	160	106	84*	1	2	-8	1659	1653	10

1	4	5	746	746	11	1	5	17	167	27	94*	1	7	-14	133	41	69*	1	8	0	740	719	13	1	9	14	151	207	77*
1	4	6	161	127	18	1	5	18	0	58	76*	1	7	-13	943	931	16	1	8	1	1085	1123	11	1	9	15	163	140	35*
1	4	7	482	498	12	1	5	19	71	77	78*	1	7	-12	196	222	23	1	8	2	1096	1040	11	1	9	16	0	89	77*
1	4	8	370	377	11	1	5	20	149	87	78*	1	7	-11	930	882	14	1	8	3	1022	1034	12	1	9	17	130	58	76*
1	4	9	370	341	12	1	5	21	0	99	79*	1	7	-10	338	368	11	1	8	4	608	637	11	1	9	18	92	52	77*
1	4	10	705	714	16	1	6	-21	114	182	82*	1	7	-9	1006	1080	10	1	8	5	65	161	50*	1	9	19	0	130	79*
1	4	11	517	495	12	1	6	-20	56	19	81*	1	7	-8	335	360	9	1	8	6	199	238	19	1	10	19	180	194	32*
1	4	12	774	795	18	1	6	-19	214	183	30	1	7	-7	653	658	13	1	8	7	228	219	18	1	10	-18	33	30	77*
1	4	13	686	679	14	1	6	-18	111	21	78*	1	7	-6	376	386	9	1	8	8	281	271	16	1	10	-17	220	160	27*
1	4	14	481	490	13	1	6	-17	124	147	48*	1	7	-5	203	203	14	1	8	9	314	305	16	1	10	-16	42	64	78*
1	4	15	423	416	14	1	6	-16	0	27	63*	1	7	-4	430	425	11	1	8	10	0	57	71*	1	10	-15	0	64	76*
1	4	16	379	380	16	1	6	-15	144	70	37*	1	7	-3	0	103	51*	1	8	11	764	711	19	1	10	-14	0	47	52*
1	4	17	375	384	17	1	6	-14	69	93	50*	1	7	-2	428	384	10	1	8	12	0	152	73*	1	10	-13	109	135	63*
1	4	18	276	244	22	1	6	-13	432	398	13	1	7	-1	1202	1185	10	1	8	13	549	564	14	1	10	-12	259	275	16
1	4	19	124	176	51*	1	6	-12	247	190	18	1	7	0	551	550	13	1	8	14	0	58	75*	1	10	-11	236	275	19
1	4	20	162	162	37*	1	6	-11	817	773	14	1	7	1	1213	1202	10	1	8	15	631	650	13	1	10	-10	229	243	20
1	4	21	175	136	35*	1	6	-10	185	146	19	1	7	2	524	514	10	1	8	16	206	129	28	1	10	-9	874	885	16
1	5	-21	117	140	71*	1	6	-9	1079	1080	11	1	7	3	1548	1613	10	1	8	17	553	541	13	1	10	-8	684	675	17
1	5	-20	0	107	84*	1	6	-8	132	130	18	1	7	4	412	382	10	1	8	18	202	124	29	1	10	-7	1078	1079	14
1	5	-19	420	386	17	1	6	-7	1491	1532	9	1	7	5	1847	1847	11	1	8	19	369	391	18	1	10	-6	980	984	14
1	5	-18	233	266	29	1	6	-6	536	512	12	1	7	6	210	157	17	1	8	20	0	77	78*	1	10	-5	1324	1329	13
1	5	-17	554	539	23	1	6	-5	1735	1752	9	1	7	7	1472	1533	13	1	9	-20	0	102	77*	1	10	-4	1003	992	14
1	5	-16	265	325	23	1	6	-4	268	302	9	1	7	8	0	32	65*	1	9	-19	187	215	32*	1	10	-3	1746	1790	12
1	5	-15	876	923	19	1	6	-3	1344	1370	9	1	7	9	1482	1490	14	1	9	-18	84	113	64*	1	10	-2	1305	1312	13
1	5	-14	501	524	12	1	6	-2	651	621	10	1	7	10	381	367	14	1	9	-17	274	340	22	1	10	-1	1885	1892	12
1	5	-13	1083	1063	15	1	6	-1	1685	1688	9	1	7	11	842	822	18	1	9	-16	116	118	74*	1	10	0	999	968	13
1	5	-12	624	646	14	1	6	0	300	265	9	1	7	12	0	42	70*	1	9	-15	430	379	14	1	10	1	1175	1195	13
1	5	-11	1211	1258	13	1	6	1	1285	1280	9	1	7	13	492	422	12	1	9	-14	56	153	62*	1	10	2	450	460	10
1	5	-10	506	572	12	1	6	2	89	46	49*	1	7	14	163	155	33*	1	9	-13	709	718	13	1	10	3	779	770	16
1	5	-9	878	890	12	1	6	3	1019	1043	10	1	7	15	232	217	25	1	9	-12	262	273	15	1	10	4	99	181	65*
1	5	-8	535	540	13	1	6	4	161	100	18	1	7	16	0	27	75*	1	9	-11	467	471	10	1	10	5	284	276	16
1	5	-7	470	479	8	1	6	5	557	555	10	1	7	17	67	16	74*	1	9	-10	39	70	47*	1	10	6	193	193	23
1	5	-6	139	139	14	1	6	6	497	468	12	1	7	18	93	25	65*	1	9	-9	757	795	15	1	10	7	110	24	57*
1	5	-5	161	179	12	1	6	7	84	148	62*	1	7	19	0	130	79*	1	9	-8	192	156	21	1	10	8	138	122	71*
1	5	-4	82	65	23*	1	6	8	0	33	62*	1	7	20	161	100	37*	1	9	-7	849	833	14	1	10	9	80	96	72*
1	5	-3	325	347	8	1	6	9	661	635	12	1	7	21	151	69	40*	1	9	-6	0	76	63*	1	10	10	211	221	24
1	5	-2	282	263	8	1	6	10	131	115	67*	1	7	22	187	215	32*	1	9	-5	208	214	18	1	10	11	512	541	12
1	5	-1	273	266	8	1	6	11	784	782	17	1	7	23	0	40	76*	1	9	-4	169	123	20	1	10	12	117	106	72*
1	5	0	510	562	11	1	6	12	86	59	70*	1	7	24	0	101	76*	1	9	-3	280	260	13	1	10	13	488	491	13
1	5	1	982	987	8	1	6	13	865	869	18	1	7	25	0	57	74*	1	9	-2	399	396	10	1	10	14	292	317	21
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3	18																													

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11	0	-8	411	400	34	11	3	-7	94	46	99*	11	6	-9	257	221	57*	11	9	3	314	285	39	12	4	-2	0	217	99*
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11	1	-9	0	51	99*	11	3	5	310	312	37	11	6	3	253	257	46*	12	0	2	174	91	98*	12	6	-3	326	333	48
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* = reflection not used in the data refinement